Metal-Metal Redox Exchange to Produce Heterometallic Manganese-Cobalt Oxo Cubanes *via* a "Dangler" Intermediate

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Abstract: Pendent metals bound to heterocubanes are key components of well-known active sites in enzymes that mediate difficult chemical transformations. Investigations into the specific role of these metal ions, sometimes referred to as "danglers," have been hindered by a paucity of rational synthetic routes to appropriate model structures. To generate pendent metal ions bonded to an oxo cubane through a carboxylate bridge, the cubane $Co_4(\mu_3\text{-}O)_4(OAc)_4(t\text{-}Bupy)_4$ (OAc = acetate, *t*-Bupy = 4-*tert*-butylpyridine) was exposed to various metal acetate complexes. Reaction with $Cu(OAc)_2$ gave the structurally characterized (by X-ray diffraction) dicopper dangler $Cu_2Co_4(\mu_4\text{-}O)_2(\mu_3\text{-}O)_2(OAc)_6(Cl)_2(t\text{-}Bupy)_4$. In contrast, the analogous reaction with $Mn(OAc)_2$ produced the Mn^{IV}-containing cubane cation $[MnCo_3(\mu_3\text{-}O)_4(OAc)_4(t\text{-}Bupy)_4]^+$ by way of a metal-metal exchange that gives $Co(OAc)_2$ and $[Co^{III}(\mu\text{-}OH)(OAc)]_n$ oligomers as byproducts. Additionally, reaction of the formally Co^{IV} cubane complex $[Co_4(\mu_3\text{-}O)_4(OAc)_4(t\text{-}Bupy)_4][PF_6]$ with $Mn(OAc)_2$ gave the corresponding Mn-containing cubane in 80% yield. A kinetic and mechanistic examination of the related metal-metal exchange reaction between $Co_4(\mu_3\text{-}O)_4(OBz)_4(py)_4$ (OBz = benzoate) and $[Mn(acac)_2(py)_2][PF_6]$ by UV-vis spectroscopy provided support for a process involving rate-determining association of the reactants and electron transfer through a μ-oxo bridge in the adduct intermediate. The rates of exchange correlate with the donor strength of the cubane pyridine and benzoate ligand substituents; more electron-donating pyridine ligands accelerate metal-metal exchange, while both electron donating and withdrawing benzoate ligands can accelerate exchange. These experiments suggest that the basicity of the cubane oxo ligands promotes metal-metal exchange reactivity. The redox potentials of the Mn and cubane starting materials, and isotopic labeling studies, suggest an inner-sphere electron transfer mechanism in a dangler intermediate.

Introduction

Tetrametallic tetrachalcogenide heterocubane clusters (M_4E_4, E = O, S) are important reaction centers that effectively mediate proton-coupled, multi-electron chemical processes in biological systems.[1-4] Prominent examples include $[Fe_4S_4]$ iron sulfur clusters that are essential electron-transfer cofactors for a range of enzymes.[5-7] The $CaMn_4O_4$ cluster known as the oxygen evolving complex (OEC, Figure 1a) of Photosystem II, and the $NiFe_4S_4$ active site of carbon monoxide dehydrogenase (CODH) are structurally similar heterocubane complexes that play an important role in enzymatic catalysis.[8] The importance of this structural motif in biology has motivated significant research into the role of the tetrametallic assembly, and much of this effort has benefited from the use of synthetic models that mimic structural and spectroscopic aspects of the appropriate cubane.[5, 9-11] Indeed, synthetic models were instrumental in establishing the basic geometry of the OEC before high resolution X-ray diffraction enabled determination of the active site arrangement.[12-14] Current OEC model systems effectively replicate EPR spectra observed for the S2 and S3 state of the OEC, and structurally reproduce nearly all of the key OEC bonding motifs (Figure 1b).[15-24] Despite the progress made from these studies, the synthesis of cubane models that accurately mimic biological function remains a significant challenge, in part due to a limited understanding of cubane assembly mechanisms. Investigations into the assembly and metal-metal exchange reactivity of iron sulfur clusters has demonstrated that these clusters can readily undergo Fe-Fe exchange, as shown by isotopic labeling studies.[25-29] However, general synthetic methods that allow construction of heterometallic cubanes analogous to those observed in biology are quite limited, and the mechanisms by which heterocubanes assemble remain poorly understood.[30-34]

An interesting feature of certain cubane active sites is their close association with an additional redox-active metal ion that is directly involved in the function of the enzyme.[8] In the OEC, a $CaMn_3O_4$ oxo cubane is associated with a pendent Mn ion referred to as the "dangler" (Figure 1a), whereas an Fe_3NiS_4 cubane of the CODH active site is bound to an Fe ion.[35, 36] In proposed catalytic mechanisms for both the OEC and CODH, these exogenous metal ions directly assist in utilizing accumulated redox equivalents of the cubane to perform difficult reactions at low overpotentials.[37-43] Despite clear indication that these dangler ions are essential, a detailed mechanistic understanding of their specific role is lacking.

Therefore, an important research objective is the preparation and study of cubane model systems that possess redox-active, dangler metal centers. Only a few dangler-cubane assemblies have been synthesized (e.g., Figure 1b,c), including a manganese-cobalt oxo cubane synthesized by complexation of "$Co^{II}(NO_3)(OAc)$" to the oxo cubane $MnCo_3(\mu_3\text{-}O)_4(OAc)_5(py)_3$ (OAc = acetate, py = pyridine).[19, 20, 44-46] In this synthesis, an acetate migration assists in complexation of the incoming metal ion to produce a new, κ^1,κ^1-(OAc)MOM' binding mode that supports the dangler structure (Figure 1c). Thus, it appeared that a reasonable starting point for the synthesis of other dangler complexes could involve established oxo cubanes such as $Co_4(\mu_3\text{-}O)_4(OAc)_4(py)_4$ in reactions of acetate-containing sources of a dangler ion.[47]

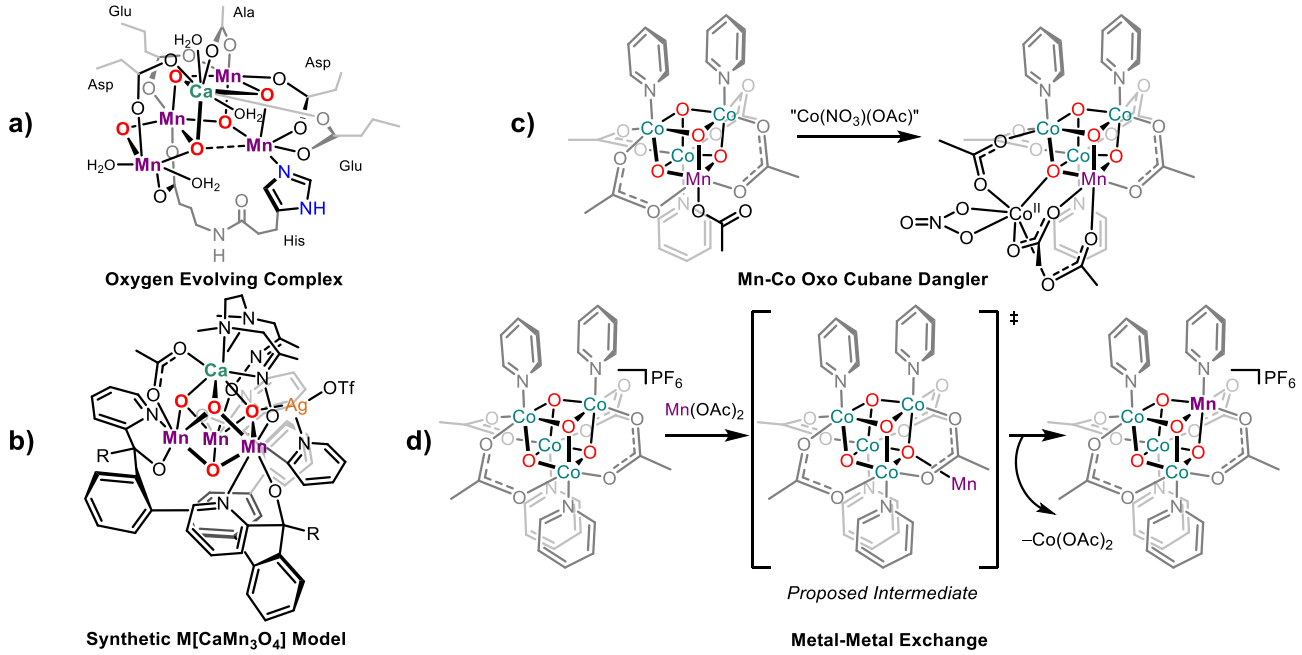


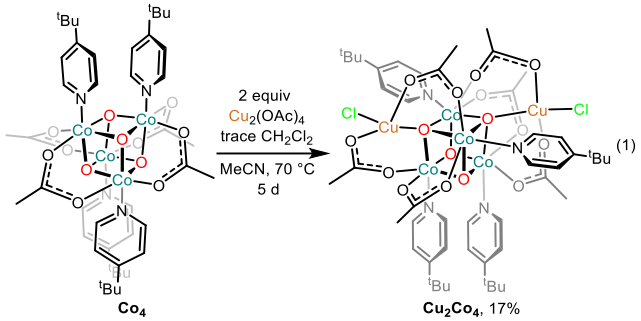
Figure 1. Established metal oxo cubane "dangler" complexes. (a) The oxygen evolving complex (OEC) of Photosystem II. (b) A pentametallic structural analogue of the OEC. (c) A pentametallic model of the OEC complex. (d) Generalized exchange reaction described in this contribution.

This contribution describes the application of this strategy to investigate incorporation of divalent metal ions as danglers onto a cobalt oxo cubane ($[Co_4O_4]$) core (Scheme 1d). A dangler complex containing two Cu^{II} ions has been discovered, as well as an unexpected metal-metal (Co-Mn) exchange reaction that provides a convenient route to heterometallic oxo cubanes with a manganese-cobalt oxo cubane ($[MnCo_3O_4]$) core.[44] As shown herein, this exchange likely occurs through a dangler intermediate possessing acetate-bridged Co and Mn centers. This redox-based exchange process incorporating a Mn^{IV} ion is possibly relevant to the mechanism of biosynthesis of the OEC, which is believed to require multiple Mn^{II} ion oxidation events and a conformational rearrangement in order to generate a Mn_2Ca OEC precursor.[27, 30, 48, 49] This process may generally point the way to synthesis of new heterometallic cubane complexes, and may also serve as a molecular model for "Galvanic exchange" processes used for doping metal oxide nanoparticles and bulk materials.[50-54]

Results and Discussion

Synthesis of an Oxo Cubane with Pendent Copper Centers. The possible binding of a monometallic "dangler" metal ion to the oxo cubane cluster, $Co_4O_4(OAc)_4(t\text{-Bupy})_4$ (**Co_4**, *t*-Bupy = 4-*tert*butylpyridine), was investigated by screening its reactions with divalent transition metal acetates. This approach was motivated by the prior observation that "Co(OAc)(NO_3)" adds to a $[MnCo_3O_4]$ cubane to produce the $MnCo_4O_4(OAc)_6(NO_3)(py)_3$ dangler complex (Figure 1c).[44] In the latter process, a bridging acetate ligand detaches from the cubane core to capture the additional metal ion, and this hemilability serves as a pathway to the dangler complex.

Treatment of cubane **Co₄** with $Ca(OAc)_2$, $Fe(OAc)_2$, $Ni(OAc)_2$, or $Zn(OAc)_2$ in MeCN solution at 23 °C or 55 °C led to complex reaction mixtures that did not appear to contain a dangler complex, as determined by ^{1}H NMR spectroscopy and high-resolution electrospray ionization mass spectrometry (HR-ESI-MS). However, when a solution of **Co₄** and $Cu_2(OAc)_4(H_2O)_2$ in MeCN solution (with trace CH_2Cl_2 as solvent of crystallization for **Co₄**) was heated at 70 °C for 16 h, followed by cooling to 23 °C, dark green crystals of paramagnetic $Cu_2Co_4(\mu_4\text{-}O)_2(\mu_3\text{-}O)_2(OAc)_6Cl_2(t\text{-}Bupy)_4$ (**Cu_2Co_4**) were obtained in 17% yield (eq 1). Single-crystal X-ray diffraction (SC-XRD) analysis enabled determination of the solid-state molecular structure, which may be described as a "double dangler" complex, with each copper atom bound to a bridging oxo ligand of the cubane (to create two μ_4-oxos) and linked to the cubane core by two bridging acetates. Presumably the chloride ligands on copper originate from the CH_2Cl_2 present in samples of **Co₄**. Thus, the reaction results in retention of the original cubane ligands, and formal addition of two "CuCl(OAc)" units to adjacent corners of the cubane. The resulting cluster displays approximate C_2 symmetry that relates the two four-coordinate Cu centers.



The ^{1}H NMR spectrum of complex **Cu_2Co_4** in acetonitrile-d_3 is consistent with the desymmetrization of **Co₄** (Figure S6); however, the broad paramagnetic resonances cannot definitively confirm the structure. This bonding arrangement is accompanied by exchange of acetate and pyridine coordination sites for two of the Co centers, to produce the observed heterobimetallic CoOCu(1κ^2,2κ^2-OAc) chelate rings. Binding of the Cu ions results in slight elongation of the Co–O(oxo) bonds (1.89(2) Å) with respect to those for $Co_4(\mu_3\text{-}O)_4(OAc)_4(py)_4$ (1.86(2) Å), and all Co–(μ_4-O) distances are elongated to an average of 1.906(6) Å. The solution magnetic moment of 3.3 μ_B (Evans method[55]) is consistent with two isolated Cu^{II} centers.

Cyclic voltammetry of compound **Cu_2Co_4** in MeCN reveals a reversible oxidation at 1.05 V vs. ferrocene/ferrocenium (Figure S8). Interestingly, the binding of the copper fragments increases the oxidation potential of the parent cubane **Co₄** by 0.77 V. Sweeping cathodically reveals a host of overlapping, irreversible reduction events that could not be assigned.

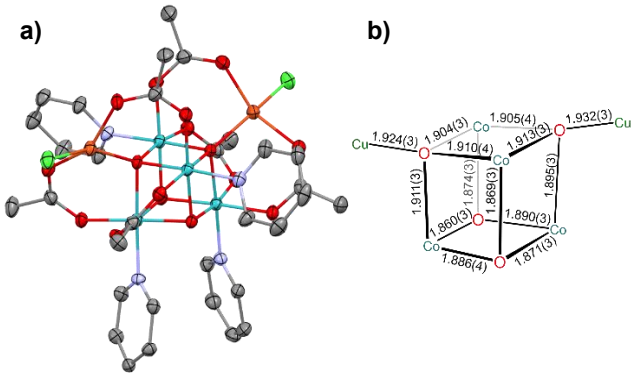
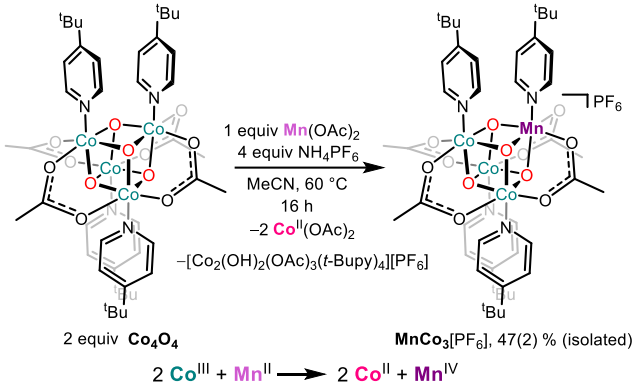


Figure 2. (a) Solid-state molecular structure of complex **Cu_2Co_4**. Solvent of recrystallization, other equiv of **Cu_2Co_4** in the asymmetric unit, and *tert*-butyl groups removed for clarity (b) Cluster bond distances.

Mn-Co Exchange in Cobalt Oxo Cubane Co₄. The isolation of complex **Cu_2Co_4** suggested that other metal acetates could generate dangler complexes, and a Mn dangler would be most relevant as a model for the OEC. Addition of 1 equiv of $Mn(OAc)_2$ to a solution of **Co₄** in MeCN at 60 °C resulted in a color change from green to brown over 16 h, and in the presence of two equiv of NH_4PF_6 (to enhance solubility), a red homogeneous solution was obtained (Scheme 1). Analysis of this crude reaction mixture by HR-ESI-MS indicated the presence of a tetra(*t*-Bupy)-substituted cubane cation analogous to the previously reported $[MnCo_3(\mu_3\text{-}O)_4(OAc)_4(py)_4]^+$, as well as several Co^{III} acetate oligomers, most notably the dicobalt complex $[Co_2(\mu\text{-}OH)_2(OAc)_3(t\text{-}Bupy)_4]^+$ which was also identified by ^{1}H NMR spectroscopy (*vide infra*; Scheme 1).[44,56] Generation of the formally $Mn^{IV}Co^{III}_3$ cubane corresponds to a redox reaction whereby Mn^{II} is oxidized by two equiv of Co^{III}, with one equiv of Co^{III} acting as a sacrificial oxidant (Scheme 1).

Workup of the reaction mixture involved filtration to separate the precipitated byproduct, pink $Co(OAc)_2$ powder (*vide infra*), followed by crystallization from CH_2Cl_2/pentane to give red crystals of $[MnCo_3(\mu_3\text{-}O)_4(OAc)_4(t\text{-}Bupy)_4][PF_6]$ (**$MnCo_3$**$[PF_6]$, 47(2)% isolated yield) that were suitable for single-crystal X-ray diffraction. The solid-state molecular structure indicates retention of the basic oxo cubane framework, and the refinement improved when the metal vertex was modeled as partially occupied by Mn. Similar to what was previously observed for an analogous tetrapyridine $[MnCo_3]$ cubane system (with py ligands), **$MnCo_3$**$[PF_6]$ crystallizes as an inversion-twin in the $P\bar{4}n2$ space group, such that ¼ of the cubane resides in the asymmetric unit, making determination of the precise location of the Mn atom impossible.[44] The ^{1}H

Scheme 1. Metal-metal Exchange between Co_4 and $Mn(OAc)_2$ with the Corresponding Formal Redox Equation.



NMR spectrum of **$MnCo_3$**[PF_6] in acetonitrile-d_3 contained broad resonances consistent with a paramagnetic species, as confirmed by an Evans method measurement (3.31 μ_B, S = 3/2). The identity of the [$Co_2(\mu\text{-}OH)_2(OAc)_3$(*t*-Bupy)$_4$][$PF_6$] byproduct (10% yield by ^{1}H NMR spectroscopy) was confirmed by isolation and characterization by ^{1}H NMR spectroscopy and HR-ESI-MS.[56, 57] Other Co^{III} acetate *t*-Bupy species present (by ^{1}H NMR spectroscopic and mass spectrometric analysis) appear to correspond to oligomeric species.[58, 59]

The limited solubility of $Mn(OAc)_2$ in organic solvents made determination of the stoichiometry by NMR spectroscopy difficult, and motivated use of the more soluble Mn^{II} starting material $Mn(OTf)_2$. Acetonitrile-d_3 solutions with various **Co_4**:$Mn(OTf)_2$ ratios were heated to 50 °C for 24 h and the reactions were monitored by ^{1}H NMR spectroscopy, with 1,3,5-trimethoxybenzene (TMB) as an internal standard (Figure S94). When the **Co_4**:$Mn(OTf)_2$ molar ratio was lower than 2:1, **Co_4** was completely converted to several diamagnetic species, presumably Co^{III} oligomers (as substantiated by HR-ESI-MS), and broad paramagnetic resonances for **$MnCo_3^+$** were observed. Consistent with the expected stoichiometry for this redox reaction, the isolated yield of **$MnCo_3$**[PF_6] from the reaction of **Co_4** with 1 equiv of $Mn(OTf)_2$ and NH_4PF_6 was 40(5)%.

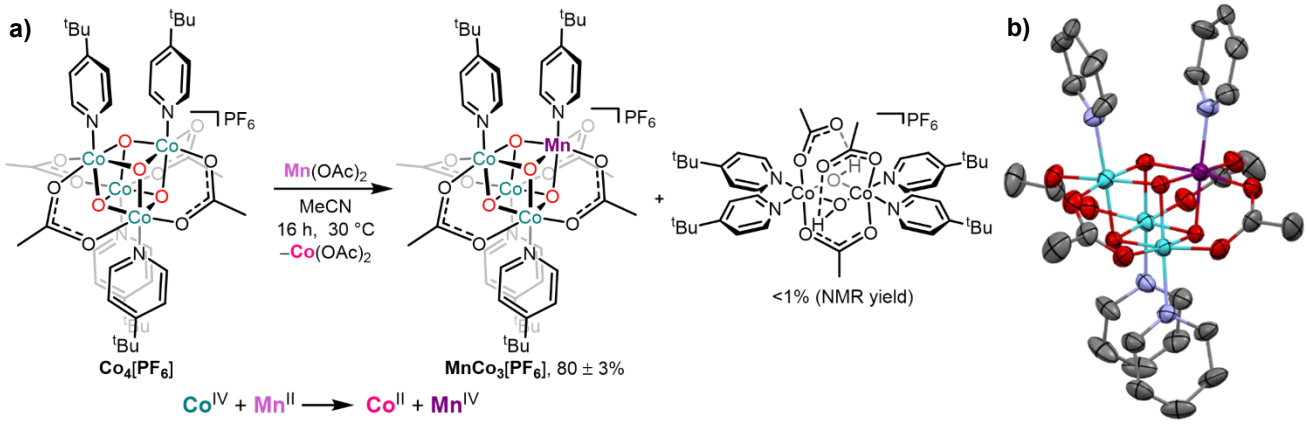


Figure 2. (a) Synthesis of **$MnCo_3$**[PF_6] from **Co_4^+**[PF_6]. (b) Solid-state molecular structure of **$MnCo_3$**[PF_6] with 50% probability thermal ellipsoids shown and one metal ion arbitrarily assigned as Mn (see main text). Hydrogen atoms, *tert*-butyl groups, and counterion removed for clarity.

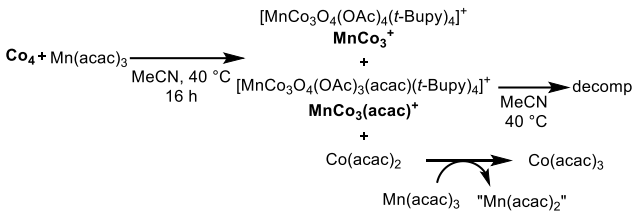
Since the product of this metal-metal exchange reaction contains a Mn^{IV} center, the reaction of $Mn(OAc)_2$ with the oxidized and formally Co^{IV}-containing [$Co^{IV}Co_3(\mu_3\text{-}O)_4(OAc)_4$(^{t}BuPy)$_4$][$PF_6$] (**$Co_4^+$**[$PF_6$]) was also examined. A vigorously stirred solution of $Mn(OAc)_2$ and **Co_4^+**[PF_6] (1:1) in MeCN at 30 °C resulted in a slow color change from brown/olive green to red over 16 h (Figure 3). Removal of the solvent *in vacuo* and extraction with CH_2Cl_2 resulted in a dark brown-red solution and a pink solid (presumably $Co(OAc)_2(L)_n$,) that was separated by filtration. Layer diffusion of pentane into the dichloromethane solution at –5 °C over 48 h produced red blocks of **$MnCo_3$**[PF_6] (80(3)% yield) and accordingly, HR-ESI-MS analysis of the crystalline material indicated the presence of a single cation corresponding to **$MnCo_3^+$**. The pink solid byproduct was analyzed by HR-ESI-MS, ^{1}H NMR, and FTIR spectroscopy, all of which suggest the presence of $Co(OAc)_2(L)_n$ (L = *t*-Bupy, H_2O, Figure S110). The stoichiometry of this exchange, initiated from either $Mn(OAc)_2$ or $Mn(OTf)_2$ in acetonitrile-d_3, was examined by ^{1}H NMR spectroscopy, and in both cases one equivalent of Mn^{II} per equivalent of Co^{IV} was required for the complete consumption of **Co_4^+**[PF_6] (Figure S96). A trace amount (<1% yield vs internal standard) of [$Co_2(\mu\text{-}OH)_2(OAc)_3$(^{t}BuPy)$_4$][$PF_6$] in the product mixture may be attributed to reaction of residual unoxidized **Co_4** in the sample of **Co_4^+**[PF_6] (Figure S89). Monitoring the metal-metal exchange of **Co_4** or **Co_4^+**[PF_6] with $Mn(OAc)_2$ by ^{1}H NMR spectroscopy in dry, degassed acetonitrile-d_3 in a sealed J-Young NMR tube under an N_2 atmosphere indicated that the reactions are not influenced by water and air, since the product distributions and qualitative reaction rates were unaffected. A balanced

redox equation for this metal-metal exchange reaction requires one equivalent of Co^{IV} and one equivalent of Mn^{II} to give Co^{II} and Mn^{IV} in a two-electron redox process (Figure 3)

Reactions of Co_4 with Mn^{III} sources. To probe the possible participation of intermediate Mn^{III} species in the above processes, redox reactions of **Co_4** with $Mn^{III}(acac)_3$ were investigated by ^{1}H NMR spectroscopy, with yields determined by comparison to TMB as an internal standard. An acetonitrile-d_3 solution of equimolar $Mn(acac)_3$ and **Co_4**, heated at 40 °C for 16 h, gave complete conversion of $Mn(acac)_3$, 95% conversion of **Co_4**, formation of both **$MnCo_3^+$** and $[MnCo_3(\mu_3\text{-}O)_4(OAc)_3(acac)(t\text{-}Bupy)_4]^+$ (**$MnCo_3(acac)^+$**; by HR-ESI-MS and ^{1}H NMR spectroscopy), and a 27% yield of $Co(acac)_3$ in addition to other unidentified paramagnetic side products. Use of two equiv of $Mn(acac)_3$ led to complete conversion of **Co_4** and a 92% yield of $Co(acac)_3$ by ^{1}H NMR spectroscopy, along with the same $[MnCo_3O_4]$ cubane products according to the idealized stoichiometry outlined in Scheme 2. Note that an accurate determination of the yield and product ratios associated with **$MnCo_3^+$** and **$MnCo_3(acac)^+$** is difficult to establish due to very broad ^{1}H NMR resonances for these paramagnetic complexes.

Presumably, one equiv of $Mn(acac)_3$ is also consumed in the production of $Co(acac)_3$ (Scheme 2). The $[MnCo_3O_4]$ cubane products (**$MnCo_3^+$** and **$MnCo_3(acac)^+$**) produced in this transformation were isolated as PF_6^- salts by chromatographic separation following the reaction of **Co_4** with $Mn(acac)_3$ and NH_4PF_6 which gave the coeluting products **$MnCo_3$**$[PF_6]$ and **$MnCo_3(acac)$**$[PF_6]$ in a combined yield of 33%.

Scheme 2. Idealized reaction of Co_4 with $Mn(acac)_3$.



Another possible secondary reaction involves radical processes of the acac ligands, as this has been reported for other $Mn(acac)_3$ systems.[60-63] Indeed, a dichloromethane-d_2 solution of equimolar $MnCo_3(\mu_3\text{-}O)_4(OAc)_5(py)_3$ and acetylacetone heated to 35 °C for 16 h resulted in partial (ca. 33%) decomposition of the $[MnCo_3O_4]$ cubane and formation of Co^{III}(acac) products by ^{1}H NMR spectroscopy and HR-ESI-MS. Likewise, heating a solution containing a mixture of purified **$MnCo_3$**$[PF_6]$ and **$MnCo_3(acac)$**$[PF_6]$ to 40 °C resulted in depletion of **$MnCo_3(acac)^+$** by 25% relative to **$MnCo_3^+$** after 16 h (HR-ESI-MS, Figure S185). Thus, radical-based reactions of acac ligands appear to contribute to a lower stability of the **$MnCo_3(acac)^+$** cubane and may contribute to the greater than stoichiometric consumption of **Co_4** by $Mn(acac)_3$ (Table 1, entries 1-3). Overall, the data in Table 1 are consistent with the stoichiometry outlined in Scheme 2.

Table 1. Investigation of Mn^{III} Reagents.

[Mn] Source	Equiv	Conversion of Co_4 (%)[a]	Yield Co^{III} (%)[a]
$Mn(acac)_3$	0.5	80	7
	0.75	91	11
	1	95	14
	1.5	100	29
	2	100	45
Mn^{III}(*t*-Bupy)$[PF_6]$	0.5	44	13
	0.75	58	17
	1	69	25
	1.5	83	38
	2	100	63

[a]Yields determined by ^{1}H NMR spectroscopy vs 1,3,5-trimethoxybenzene as an internal standard as the sum of potential Co^{III} products.

To limit possible side products involving acac-induced decomposition, reactions with $[Mn(acac)_2(t\text{-}Bupy)_2][PF_6]$ (**Mn^{III}(*t*-Bupy)**$[PF_6]$), a Mn^{III} source with one less acac ligand, were also investigated. An acetonitrile-d_3 solution of equimolar **Mn^{III}(*t*-Bupy)**$[PF_6]$ and **Co_4**, heated at 40 °C for 16 h, resulted in full conversion of **Mn^{III}(*t*-Bupy)$^+$**, 69% conversion of **Co_4**, a 50% yield of $[Co(acac)_2(t\text{-}Bupy)_2][PF_6]$ (**Co^{III}(*t*-Bupy)**$[PF_6]$, Table 1) and an undetermined amount of **$MnCo_3^+$/$MnCo_3(acac)^+$** (observed by ^{1}H NMR and HR-ESI-MS). The reaction of **Co_4** and **Mn^{III}(*t*-Bupy)$^+$** to give **$MnCo_3^+$/$MnCo_3(acac)^+$** and **Co^{III}(*t*-Bupy)$^+$** obeyed a 1:2 (**Co_4**:Mn^{III}) stoichiometry (Table 1), following a pathway analogous to that depicted in Scheme 2. Note that, unlike the reaction of **Co_4** with $Mn(acac)_3$, this reaction does not lead to overconsumption of **Co_4** at lower **Co_4**:Mn^{III} ratios. The conversion of **Co_4** to **Co^{III}(*t*-Bupy)$^+$** suggests that one equiv of Mn^{III} undergoes metal-metal exchange with **Co_4**, and another equiv of Mn^{III} oxidizes the putative $Co(acac)_2$ byproduct to form one equiv of Co^{III}. In agreement with this hypothesis, the reaction of equimolar $Co(acac)_2$ and **Mn^{III}(*t*-Bupy)** in acetonitrile-d_3 resulted in quantitative conversion to **Co^{III}(*t*-Bupy)** by ^{1}H NMR spectroscopy.[64] On a synthetic scale, the reaction of two

equiv of **Mn^{III}(*t*-Bupy)**[PF_6] with **Co_4** gave an unoptimized yield of 11% for the mixture of **$MnCo_3^+$/$MnCo_3$(acac)$^+$**, while red crystals of **Co^{III}(*t*-Bupy)**[PF_6] were isolated in 42% yield.

By ^{1}H NMR spectroscopy, no reaction occurred for an acetonitrile-d_3 solution of $MnCo_3(\mu_3\text{-}O)_4(OAc)_5(py)_3$ and $[Mn(acac)_2(py)_2][PF_6]$ (**Mn^{III}(py)**[PF_6]) heated to 40 °C for 16 h. This indicates that a secondary reaction of the $[MnCo_3O_4]$ cubane product with **Mn^{III}(py)$^+$** does not influence the observed product distributions under these conditions. Additionally, monitoring the reaction of **Co_4** with **Mn^{III}(*t*-Bupy)$^+$** by ^{1}H NMR spectroscopy indicates that a non-oxidative ligand exchange does not compete with metal-metal exchange for the consumption of **Co_4** in the reaction of **Co_4** and **Mn^{III}(*t*-Bupy)$^+$**. Thus, the Co-Mn exchange involving **Co_4** and **Mn^{III}(*t*-Bupy)$^+$** more closely follows the idealized stoichiometry of Scheme 2 and early in this reaction there are no significant, competing processes that consume **Co_4**. The reduced 63% yield of Co^{III} in the 1:2 reaction of **Co_4** with **Mn^{III}(*t*-Bupy)$^+$** is likely due to radical-induced processes.

Metal-Metal Exchange Kinetics Involving Mn^{III}. To probe the mechanism of cobalt-manganese exchange, kinetic studies involving $[Co_4O_4]$ cubanes and Mn^{III} complexes were conducted, using initial rate data collected from reactions monitored by UV-vis spectroscopy. The cobalt oxo cubanes chosen for study are benzoate derivatives that allow tuning of the cubane electronic properties with *para* substituents in the pyridine and benzoate ligands. These substituted pyridine and benzoate complexes, e.g., $Co_4(\mu_3\text{-}O)_4(OBz)_4(py)_4$ (**Co_4(py-H)**) and $Co_4(\mu_3\text{-}O)_4(OBz)_4(\underline{t\text{-Bupy}})_4$ (**$Co_4(O_2CC_6H_4$-H)**), were readily prepared by the known procedure involving $Co(OAc)_2(H_2O)_4$ and the corresponding pyridine with H_2O_2, or by ligand-exchange reactions of a benzoic acid derivative with **Co_4** (see SI).[65-67] The Mn^{III}-based reactant $[Mn(acac)_2(py)_2][PF_6]$ **Mn^{III}(py)**[PF_6] was chosen due to its favorable solubility properties and slow rates of ligand-exchange side reactions. As determined by ^{1}H NMR spectroscopy, no reaction was observed within 0.5 h of combining equimolar solutions of **Mn^{III}(py)$^+$** and **Co_4(py-H)**. The Mn-Co exchange proceeds in a manner similar to that involving **Co_4** and **Mn^{III}(*t*-Bupy)$^+$**, to give $[MnCo_3(\mu_3\text{-}O)_4(OBz)_4(py)_4][PF_6]$ (**$MnCo_3$(py)**[PF_6]) and $[Co(acac)_2(py)_2][PF_6]$ (**Co^{III}(py)**[PF_6]). Critically, there are no side reactions that occur before metal-metal exchange in the reaction with Mn^{III} sources and **Co_4(py-H)**, and it is assumed that the measured initial rates are minimally affected by competing secondary processes. Note that ligand exchange was observed in other cases; for example, $Mn(OTf)_2$ reacts rapidly with **Co_4** by initial ligand exchange, to give triflate-containing cobalt cubane complexes (by ^{1}H NMR and ESI-MS; Figure S99) before metal-metal exchange occurred. Ligand exchange was also observed in the transfer of an acac ligand in the reaction of **Co_4** with $Mn(acac)_3$ (eq 2), which suggests operation of an inner-sphere electron-transfer process.

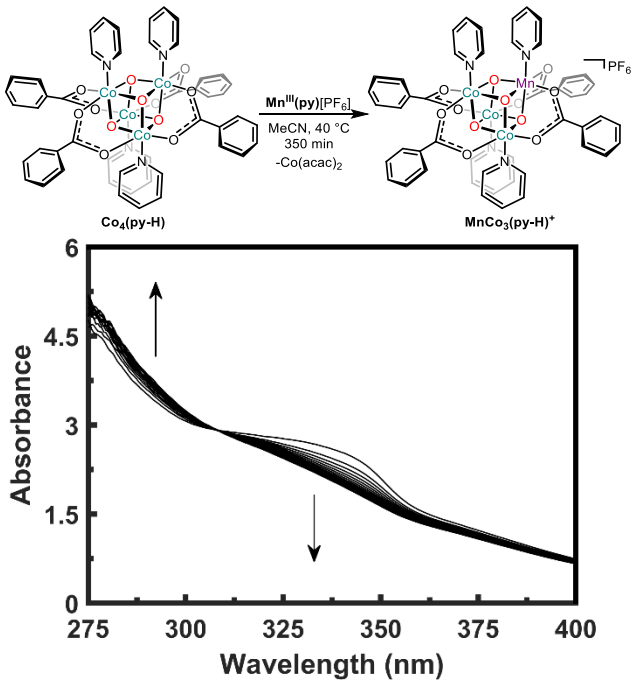


Figure 3. Representative UV-vis spectra tracking the metal-metal exchange reaction between equimolar $Mn^{III}(py)^+$ and Co_4(py-H) over 350 min.

Equimolar MeCN solutions of **Co_4(py-H)** and **Mn^{III}(py)$^+$** were combined in a sealed cuvette and the resulting reaction mixture was heated to 40 °C for 350 minutes. Monitoring by UV-vis spectroscopy revealed an approximate isosbestic point at 308 nm (Figure 3), indicating a linear relationship between the concentrations of products and reactants. Generally, in further kinetics experiments, the entire UV-vis spectrum was collected at each timepoint, and rate data was determined simultaneously in triplicate to estimate the experimental error. Attempts to fit the data to an integrated rate law did not cleanly delineate between various mechanistic possibilities, likely due to the aforementioned background oxidation reaction of $Co(acac)_2$ to give **Co^{III}(py)**.

Kinetics for the metal-metal exchange reaction were followed by UV-vis spectroscopy, with an excess of **Co_4(py-H)** or **Mn^{III}(py)$^+$** to establish pseudo first-order conditions. The slope of the log/log plot of $[\mathbf{Mn^{III}(py)^+}]_0$ vs. k_{init} (Figure 5) indicated a reaction order of 0.8(2), while the log/log plot of $[\mathbf{Co_4(py\text{-}H)}]_0$ vs. k_{init} indicated a reaction order of 0.55(3) (Figure S143). Although the fractional reaction orders are difficult to interpret, the dependence of the rate on both **Mn^{III}(py)$^+$** and **Co_4(py-H)** concentrations indicates that both species are involved in a rate-determining step. This behavior is consistent with both outer- and inner-sphere electron-

transfer mechanisms, but an inner-sphere redox process in an intermediate adduct formed by coordination of **$Mn^{III}(py)^{+}$** to a μ_3-oxo ligand of **Co_4(py-H)** seems likely. Note that in the observations described above, ligand transfer often, but not always, accompanies metal-metal exchange (*vide supra*).

Given the apparent lability of the complexes in this process, the role of added ligands on the reaction kinetics was investigated by monitoring the reaction of **Co_4(py-H)** and **$Mn^{III}(py)^{+}$** in the presence of added pyridine or benzoate (as $^{n}Bu_4NOBz$). Additional pyridine had no discernible effect on k_{init} (Figure S145), indicating that pyridine ligand dissociation does not significantly influence the rate of metal-metal exchange. Additional benzoate inhibited the rate of metal-metal exchange, as indicated by the –1.0(3) order dependence of k_{init} on $[^{n}Bu_4NOBz]_0$; this inhibitory effect could result from binding of benzoate to **$Mn^{III}(py)^{+}$** to render it less reactive. Alternatively, benzoate might add to the cubane via coordination to a cobalt center, to form an anionic complex. While a solution of $^{n}Bu_4NOBz$ and **Co_4** (1:1) remained unchanged after 16 h (by ^{1}H NMR spectroscopy), addition of $^{n}Bu_4NOBz$ (1 equiv) to a solution of **$Mn^{III}(py)^{+}$** in acetonitrile-d_3 immediately resulted in several new products, as indicated by broad resonances in the ^{1}H NMR spectrum. It is presumed that benzoate adds to **$Mn^{III}(py)^{+}$** to generate one or more less-reactive (likely less electrophilic) Mn complexes. The rate of metal-metal exchange was also examined in CH_2Cl_2, *ortho*-difluorobenzene, DMSO, 1,1,2,2-tetrachloroethane, and THF; however, no relationship between the solvent dielectric constant and k_{init} was discerned across this series.

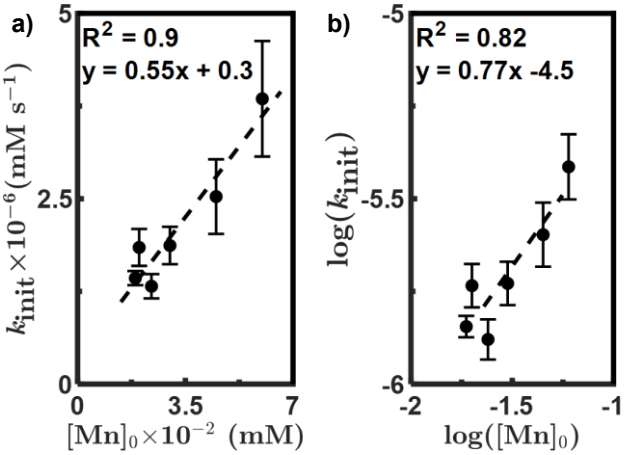
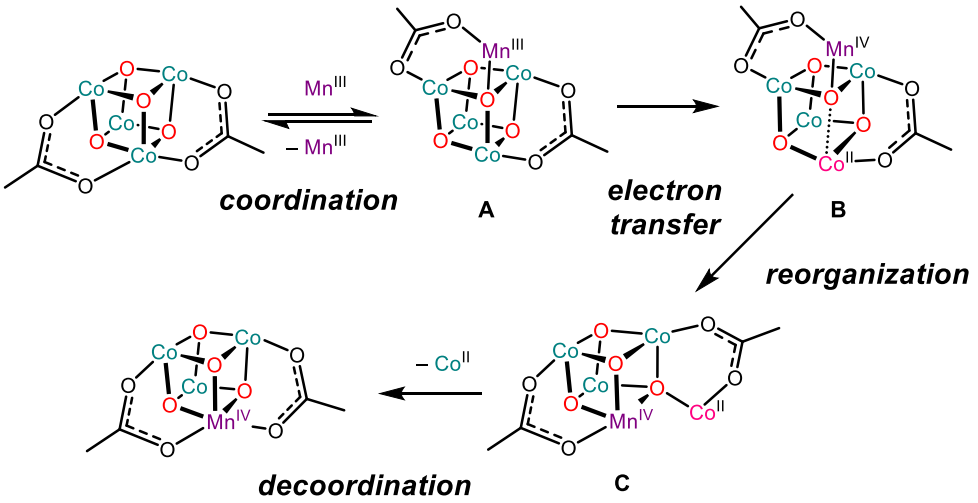


Figure 4. Linear Relationship between $[\mathbf{Mn^{III}(py)^{+}}]_0$ and the initial rate constant for metal-metal exchange between **Co_4(py-H)** and **$Mn^{III}(py)^{+}$**. (a) $[\mathbf{Mn^{III}(py)^{+}}]_0$ vs the rate of metal-metal exchange; k_{init} varies linearly with $[\mathbf{Mn^{III}(py)^{+}}]_0$. (b) Double logarithm plot of $[\mathbf{Mn^{III}(py)^{+}}]_0$ vs. k_{init} indicating a reaction order of 0.8(2) order in $[\mathbf{Mn^{III}(py)^{+}}]_0$.

The above results are consistent with the simplified mechanistic model for metal-metal exchange shown in Scheme 3. Initially, the $[Co_4O_4]$ cubane binds the incoming Mn center by donation from a μ_3-oxo ligand, which is likely assisted by an acetate migration (**A**). An inner-sphere electron transfer in this intermediate produces a Co^{II} center, which should significantly weaken the resulting μ_4-O–Co^{II} bond (**B**). Note that Co–O bond elongations occur upon binding a Lewis acid, even without a reduction process as observed in dicopper adduct **Cu_2Co_4** (*vide supra*), but Co–O bond elongations are expected to be more pronounced with a reduced Co^{II} center and the resulting population of Co–O antibonding orbitals. The more labile Co^{II} center should then readily participate in a reorganization that exchanges it for the Mn^{IV} ion that is more stable in the octahedral site of the cubane (**C**). Note that this rearrangement may occur without dissociation or change in the binding mode for the acetate ligands, to give a $MnCo_3$ oxo cubane core with a labile Co^{II} species in a "dangler" position, analogous to that of the previously isolated complex $MnCo_4(\mu_4\text{-}O)(\mu_3\text{-}O)_3(OAc)_6(NO_3)(py)_3$.[46] This labile Co^{II} ion is readily lost by ligand exchange reactions in the reaction mixture; however, this step may be relatively slow given isolation of the aforementioned nitrate analogue. Evidence for an analogous dangler intermediate during metal-metal exchange comes from ^{1}H NMR monitoring of the formation of **$MnCo_3$** from the reaction of **Co_4** and $Mn(OAc)_2$, which reveals the build-up of an intermediate, proposed to be $MnCo_4(\mu_4\text{-}O)(\mu_3\text{-}O)_3(OAc)_7(t\text{-}Bupy)_3$, with resonances corresponding to those for $MnCo_4(\mu_4\text{-}O)(\mu_3\text{-}O)_3(OAc)_6(NO_3)(py)_3$ (Figure S98). In general, the binding of a Lewis acidic metal center to a bridging oxo ligand is expected to raise the reduction potential of the cubane. As described above, this effect is observed in complex **Cu_2Co_4**, and is also observed for the protonated cobalt oxo cubane, which exhibits an irreversible reduction at approximately –0.6 V vs. Fc/Fc^+ in MeCN.[67] It is noteworthy that Mn^{III} reduces Co^{II} in this scenario because Mn^{III} complexes such as $Mn(acac)_3$ ($Mn^{III/IV}$ $E_{1/2}$ in MeCN = 0.58 V vs. Fc/Fc^+), **$Mn^{III}(t\text{-}Bupy)^{+}$**, and basic manganese acetate $Mn_3(OAc)_6(H_2O)_3(\mu_3\text{-}O)$, are generally mild oxidants that would not be expected to reduce **Co_4** (E_{pc} in MeCN = ~–1.3 V vs. Fc/Fc^+) in an outer sphere manner.[68-70]

In general, the binding of a Lewis acidic metal center to a bridging oxo ligand is expected to raise the reduction potential of the cubane. As described above, this effect is observed in complex **Cu_2Co_4**, and is also observed for the protonated cobalt oxo cubane, which exhibits an irreversible reduction event at approximately –0.6 V vs. Fc/Fc^+ in MeCN.[67] It is noteworthy that Mn^{III} reduces Co^{II} in this scenario because Mn^{III} complexes such as $Mn(acac)_3$ ($Mn^{III/IV}$ $E_{1/2}$ in MeCN = 0.58

Scheme 3. Proposed Mechanism for M-M Exchange Between a [Co_4O_4] Cubane and Mn^{III}.



V vs. Fc/Fc^+), **Mn^{III}(*t*-Bupy)$^+$**, and basic manganese acetate $Mn_3(OAc)_6(H_2O)_3(\mu_3\text{-}O)$, are generally mild oxidants that would not be expected to reduce **Co_4** (E_{pc} in MeCN = ~–1.3 V vs. Fc/Fc^+) in an outer sphere manner.[68-70] This mismatch in redox potentials between Mn^{III} sources and Co_4O_4 cubanes supports the hypothesis that an outer-sphere electron transfer pathway for metal-metal exchange is unlikely. Thus, the structural data of **Cu_2Co_4** and electrochemical data of **Co_4**, **Cu_2Co_4**, and **Mn^{III}(py)**, support the hypothesis that the reduction from Co^{III} to Co^{II} occurs in an intermediate pentametallic species such as **A**.

The reduction of a substitutionally inert, low spin, d^6 Co^{III} ion by a d^4 Mn^{III} ion is reminiscent of the classic studies on inner-sphere electron transfer between $Co^{III}Cl$ and a d^4 Cr^{II} ion, to give Co^{II} and $Cr^{III}Cl$.[71, 72] However, in the Co-Mn exchanges discussed here, ligand exchange is likely a consequence of the binding event (**A**) and the dynamic behavior of the acetate ligands, rather than electron transfer through an acetate bridge.[73] In this inner-sphere process, the bridging ligand critical to electron-transfer is probably the μ_4–oxo ligand.[74-76] Nonetheless, the carboxylate ligands are assumed to play an important, secondary role in influencing the electron-transfer driving force by modulating redox potentials, and by contributing to the rate of initial adduct formation.

To investigate these electronic effects on the reaction rates, Hammett plots were constructed with data from reactions of cubanes possessing substituted pyridine and benzoate ligands. Due to solubility constraints, the data were collected on reactions in CH_2Cl_2, with 10 equiv of **Mn^{III}(py)** or **Mn^{III}(*t*-Bupy)** to give pseudo first-order conditions. Interestingly, a plot of k_{init} vs. σ_{para} parameters shows that for benzoate ligands, the rate may be accelerated by both electron-donating and electron-withdrawing substituents (Figure 5), suggesting the operation of opposing electronic effects. More donating benzoate ligands (X = ^{t}Bu, Me) are expected to promote the basicity of the μ_3-oxo ligands, thereby favoring formation of adduct species **A**. On the other hand, electron-withdrawing substituents (X = NO_2, CN, Br, Cl) decrease the basicity of the oxo ligands, but increase the reduction potential of the cubane, rendering Co^{III} more susceptible to the reduction step to produce **C**. An inflection point associated with the balance between these two effects corresponds to X = H/Ph.

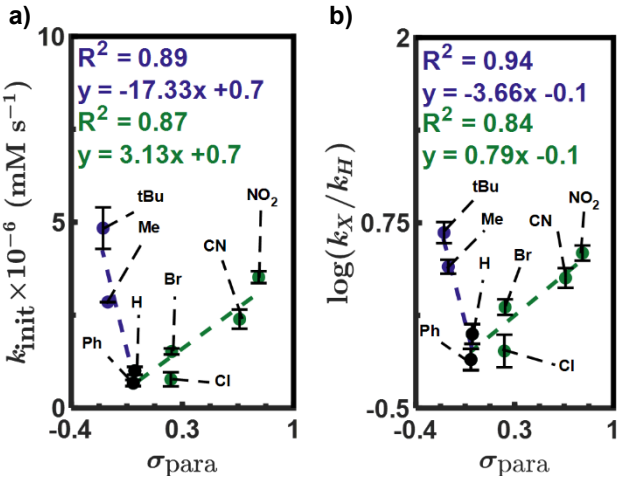


Figure 5. Influence of the *para*-benzoate ligand substituent of **$Co_4(O_2CC_6H_4\text{-}X)$** on the initial rate constant for metal-metal exchange. (a) *para*-Hammett parameter vs. initial rate constant of metal-metal exchange for **$Co_4(O_2CC_6H_4\text{-}^tBu$**, **-Me**, **-Ph**, **-H**, **-Cl**, **-Br**, **-CN**, and **-NO_2)**. Rates increase with both more donating (^{t}Bu, Me, H) and withdrawing (NO_2, CN, Cl, Br, H) substituents, with an inflection point at **$Co_4(O_2CC_6H_4\text{-}H)$**. (b) Hammett plot relating the *para*-benzoate substituent and $\log(k_X/k_H)$ with ρ values determined from the slope of the linear fits. The ρ value for electron donating substituents (–3.7(6)) compared to ρ for the electron withdrawing substituents (0.8(2)) indicates that benzoate plays a more substantial role in the electron donating regime than it does in the electron withdrawing regime.

The influence of the more basic benzoate ligands, determined by the ratio of the rate constants k_X/k_H (ρ = –3.7(6)), is greater than that of the more acidic benzoate ligands (ρ = 0.8(2)), which indicates that the ability of the cubane to coordinate Mn^{III} is critical to metal-metal exchange reactivity. This would suggest that for the Co_4O_4 cubane system, the Lewis acidity of the incoming metal ion is more important than its reducing potential in influencing the rate of metal-metal exchange. Interestingly, the most electron-rich oxo cubane in this series, **$Co_4(O_2CC_6H_4$-OMe)**, reacts with **Mn^{III}(*t*-Bupy)$^+$** but does not undergo metal-metal exchange (by ^{1}H NMR spectroscopy). This is perhaps due to stabilization of the oxidized cubane, **$Co_4(O_2CC_6H_4$-OMe)$^+$** in this case.

In the case of the substituted pyridine ligands, stronger donors result in faster metal-metal exchange, and the small, negative slope (ρ = –0.4(1) from the Hammett plot indicates that pyridine substituents are less effective than benzoates in modulating k_{init} (Figure 6). This is consistent with previous observations that the pyridine ligands impact the redox properties less than the chelating, anionic ligands.[66]

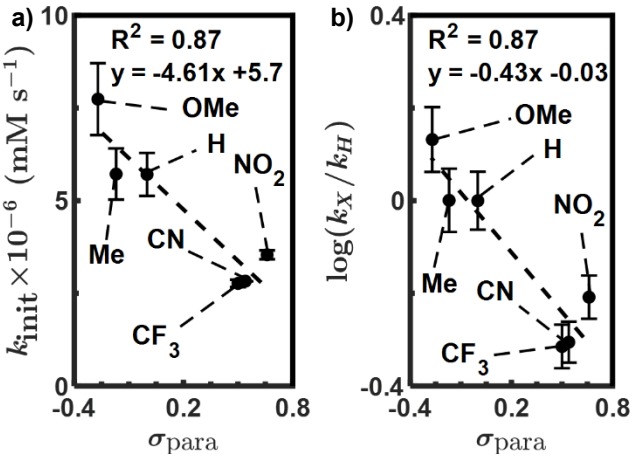


Figure 6. Influence of *para*-pyridyl substituents of Co_4(py-X) on the rate constant for metal-metal exchange. (a) *para*-Hammett parameter vs. initial rate of metal-metal exchange for **Co_4(py-OMe**, **-Me**, **-H**, **-CF_3**, **-CN**, **-NO_2**). The rate inversely correlates with the Hammett parameter (σ_{para}) of the substituent, with stronger donors leading to more rapid exchange. (b) Hammett plot relating the *para*-pyridyl substituent and $\log(k_X/k_H)$ with ρ determined from the slope of the linear fit. The low ρ value (–0.4(1)) determined from the Hammett plot is consistent with prior observations that the pyridine ligands influence the cubane redox potentials less than the chelating ligands (S140), with the pyridine ligands mostly tuning the basicity of the oxo ligands.

Role of Acetate and Pyridine Ligand Transfers in Co-Mn Exchange. To further investigate the role of ligand transfer during metal-metal exchange, deuterium-labeling studies were conducted on acetate complexes. When an MeCN solution of **Co_4** was treated with the Mn^{II} reagent Mn(OAc-d_3)$_2$ (corresponding to a 4-fold excess of OAc-d_3) for 48 h at only 23 °C, HR-ESI-MS analysis did not detect **$MnCo_3^+$** as a product, while complete scrambling of the OAc-d_3 label into **Co_4** gave a binomial distribution of isotopologues (Figure S197). This indicates that the initial adduct formation step in the metal-metal exchange process to give **A** may be reversible. Heating an identical reaction mixture at 70 °C for 2 h, and analysis by HR-ESI-MS, revealed that acetate ligand scrambling was incomplete as judged by the degree of OAc-d_3 incorporation into **Co_4** (33%) and **$MnCo_3^+$** (51%). Thus, acetate ligand exchange appears to be somewhat faster than metal-metal exchange such that the latter is accompanied by some degree of acetate ligand transfer. In another set of experiments, an acetonitrile solution of oxidized cubane **Co_4^+**[PF_6] was treated with 2 equiv of Mn(OAc-d_3)$_2$ at 40 °C for 2 h. In this case, HR-ESI-MS analysis revealed 39% incorporation of the acetate label compared to 30% incorporation of the label in "unreacted" **Co_4^+**. These experiments indicate that the cubane acetate ligands can exchange with the Mn acetate ligands before and during the metal-metal exchange reaction and suggests that a pentametallic complex may form during the metal-metal exchange process. In both cases, a binomial distribution of labeled acetates was also observed for the byproducts $Co(OAc)_n(OAc\text{-}d_3)_{2-n}$ (n = 0-2) and $[Co_2(\mu_2\text{-}OH)_2(OAc)_n(OAc\text{-}d_3)_{3-n}(^tBupy)_4]^+$ (n = 0-3).

For comparison, acetate ligand incorporation into cubane **Co_4** from acetic acid and acetate was examined. A MeCN solution of **Co_4** was stirred with four equiv of acetic acid-d_4, or triethylammonium acetate-d_4 (generated in situ by treatment of triethylamine with acetic acid-d_4) at 23 °C, and the reaction mixtures were subjected to HR-ESI-MS analysis. In both experiments, the deuterium label was incorporated rapidly (ca. 10 min), resulting in a binomial distribution of all five possible cubane isotopologues $Co_4(\mu_3\text{-}O)_4(OAc)_n(OAc\text{-}d_3)_{4-n}$ (n = 0-4). In a control experiment, equimolar MeCN solutions of $Co_4(\mu_3\text{-}O)_4(OAc)_4(py)_4$ and $Co_4(\mu_3\text{-}O)_4(OAc\text{-}d_3)_4(py\text{-}d_5)_4$ (the fully deuterated isotopologue) were combined and after 24 h at 40 °C, no acetate or pyridine isotope crossover was observed by HR-ESI-MS, which is consistent with slow ligand dissociation from a low spin Co^{III} center (Figure S192). In the presence of 2 equiv of $Mn(OAc)_2$, and heating at 70 °C for 0.5 h, these isotopologues undergo exchange of pyridine ligands. Analysis by HR-ESI-MS revealed a binomial distribution of all possible product isotopologues of the starting cubane $Co_4(\mu_3\text{-}O)_4(OAc)_4(py)_n(py\text{-}d_5)_{4-n}$ (n = 1-4) before any of the corresponding Mn cubane was detected. This experiment indicates that the pyridine ligands redistribute before the metal-metal exchange reaction. It should be noted that a redistribution of pyridine ligands was also observed when HOAc was

added in place of $Mn(OAc)_2$ under otherwise identical conditions. A pyridine-exchange process involving related dinuclear cobalt species has been reported.[56]

In summary, the acetate ligands of a $[Co_4O_4]$ cubane exchange with the acetate ligands of $Mn(OAc\text{-}d_3)_2$, and this process is competitive with metal-metal exchange. The $[Co_4O_4]$ cubane pyridine ligands exchange under the experimental conditions at a rate significantly faster than metal-metal exchange, which indicates a pyridine-dissociative mechanism, but the measured order in pyridine indicates that this does not influence the metal-metal exchange reactivity. The observed isotopologues of $\mathbf{MnCo_3^+}$ produced in these labeling experiments suggest the formation of an intermediate pendent metal complex in which ligand exchange is facile.

Concluding Remarks

The study described here significantly enhances the small but growing knowledge base associated with the synthesis and properties of transition metal oxo cubanes that possess an associated "dangler" metal ion. Such clusters may serve as models for the $Mn[CaMn_3O_4]$ oxygen evolving complex, especially if the dangler ion is redox active and bound to the cubane through carboxylate bridges.[19,20,45] Formation of the dicopper dangler $\mathbf{Cu_2Co_4}$ from well-defined precursors, by use of additional acetate ligands to tether the copper ions to the cubane, provides a unique example whereby dangler ions are bound to a $[Co_4O_4]$ cubane. Notably, the acetate ligands in this complex, like the carboxylate groups of the OEC, support a dangler ion-μ_4-oxo interaction that may play an important role in chemical transformations of the assembly. In the OEC, the comparable dangler interaction is thought to enable the structural changes required for water activation.[37] Indeed, attempts to extend the synthesis of $\mathbf{Cu_2Co_4}$ to manganese provides evidence for formation of a carboxylate-supported $Mn[Co_4O_4]$ intermediate complex possessing a $Mn(\mu_4\text{-}O)Co_3$ arrangement. The latter transitory structure appears to be key to electron-transfer processes that ultimately lead to a redox-driven exchange of metals that introduce manganese *via* a structural rearrangement of the oxo cubane core.

These studies on the chemical behavior of metal clusters in solution provide design principles for the development of new metal-metal exchange reactions.[25] The observed exchange reaction is a rare example of a well-defined metal-metal redox exchange for a tetrametallic cubane cluster. Importantly, the efficiency of this process suggests its broader utility for the controlled synthesis of predetermined oxo metal clusters with intricate structures.[77] Furthermore, the conveyance of a manganese ion to a cubane cluster is perhaps relevant to the biochemical mechanism for assembly of the OEC, which is believed to occur by photochemically driven redox processes whereby Mn^{II} is incorporated by a stepwise process into the cluster.[31,78] Additionally, the details of metal additions and exchanges of this sort may help in the development of synthetic schemes for assembly of mixed-metal nanoclusters and bulk materials. Indeed, related galvanic exchange reactions have been employed to modify an oxide or sulfide structure to incorporate a heterometal.[51-53,77,79]

Further investigations into the role of redox active dangler ions should help shed light on the possible role of these species in controlling chemical transformations of the oxo cluster core. The concepts developed in this study are now being used in syntheses of heterometallic oxo cubanes.

Associated Content

Experimental details, characterization data, kinetics data and crystallographic details. Accession Codes 2349139-2349140 ($\mathbf{Cu_2Co_4}$) and ($\mathbf{[MnCo_3][PF_6]}$) contain the supplementary crystallographic data for this paper. These data can be obtained free of charge through www.ccdc.ac.uk/data_request/cif or by emailing data_request@ccdc.cam.ac.uk, or by contacting The Cambridge Crystallographic Data Centre,12 Union Road, Cambridge CB2 1EZ, UK; fax: +44 1223 336033.

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Graphical Abstract