Lead-Free Semiconductors: Phase-Evolution and Superior Stability of Multinary Tin Chalcohalides

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ABSTRACT: Tin-based semiconductors are highly desirable materials for energy applications due to their low toxicity and biocompatibility relative to analogous lead-based semiconductors. In particular, tin-based chalcohalides possess optoelectronic properties that are ideal for photovoltaic and photocatalytic applications. In addition, they are believed to benefit from increased stability compared to halide perovskites. However, to fully realize their potential, it is first necessary to better understand and predict the synthesis and phase evolution of these complex materials. Here, we describe a versatile solution-phase method for the preparation of the multinary tin chalcohalide semiconductors $Sn_2SbS_2I_3$, $Sn_2BiS_2I_3$, Sn_2BiSI_5, and Sn_2SI_2. We demonstrate how certain thiocyanate precursors are selective toward the synthesis of chalcohalides, thus preventing the formation of binary and other lower order impurities rather than the preferred multinary compositions. Critically, we utilize ^{119}Sn ssNMR spectroscopy to further assess the phase purity of these materials. Further, we validate that the tin chalcohalides exhibit excellent water stability under ambient conditions, as well as remarkable resistance to heat over time compared to halide perovskites. Together, this work enables the isolation of lead-free, stable, direct band gap chalcohalide compositions that will help engineer more stable and biocompatible semiconductors and devices.

Semiconductors are ubiquitous optoelectronic materials that are critical to the functionality and efficiency of many existing technologies and devices. Their widespread use and our reliance on semiconductors for many new emerging technologies make it especially important to focus on materials that demonstrate enhanced stability and biocompatibility. For instance, while lead halide perovskites are increasingly used in photovoltaic, catalytic, and light-emitting devices (LEDs), they are often based on a toxic element (Pb), and suffer from water-, thermal- and photo-induced phase segregation and degradation.[1,2] Combined, these problems result in significant environmental and technological challenges that can hinder device performance and significantly slow down the deployment of new technologies.[3,4,5]

In the search for alternatives, chalcohalides are quickly gaining attention as semiconductors for multiple applications.[6,7,8,9,10,11,12,13] For example, quaternary lead-free $Sn_2PnS_2I_3$ (where Pn = Sb or Bi) chalcohalides display some of the most desirable features of chalcogenide and halide materials, including direct, visible band gaps (< 1.6 eV).[14,15,16,17,18,19] Following their original discovery and crystallographic determination, these materials have earned renewed interest in photovoltaics, thermoelectrics, and catalysis applications.[20,21,22,23,24] Furthermore, based on preliminary solar devices, tin chalcohalides are believed to exhibit inherently high stability and power conversion efficiency (PCEs), highlighting their potential as valuable materials for energy conversion.[25,26]

Typically, chalcohalide single crystals are grown after direct solid-state reaction between the elements (Figure 1).[27,28,29] However, aside from the preparation of $Sn_2SbS_2I_3$ thin-films employing a high temperature annealing approach,[25] the solution-phase synthesis of tin-based chalcohalides remains relatively unexplored.[30] Moreover, while the colloidal synthesis of bismuth-based, ternary chalcohalides such as BiSI and $Bi_{13}S_{18}I_2$ has undergone a renaissance in the past few years,[31,3233,34,35,36,37] surprisingly little synthetic attention has been paid to tin-based, ternary chalcohalides such as Sn_2SI_2.[38,39,40]

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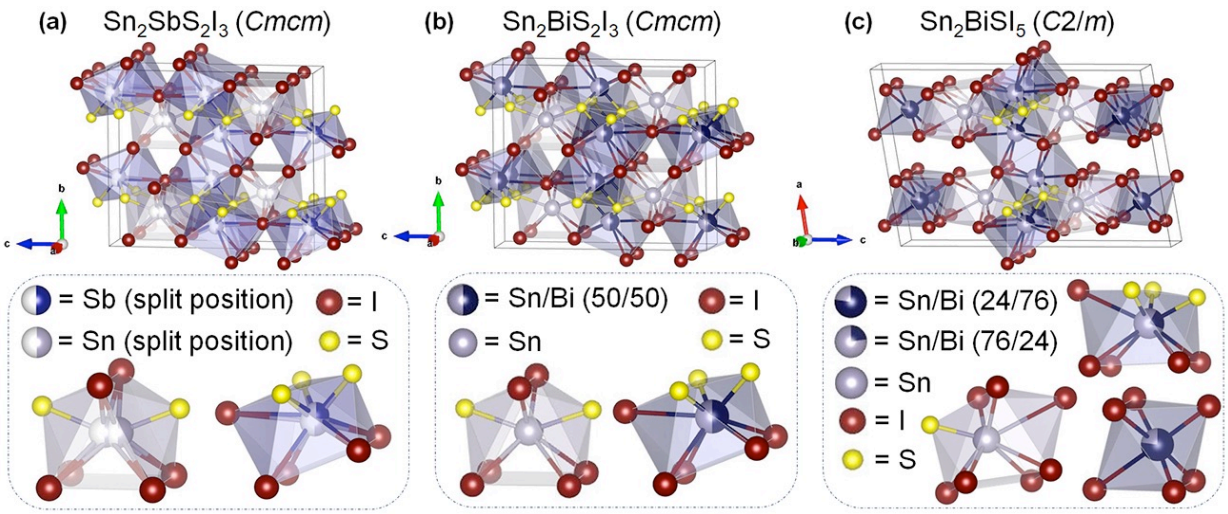

Figure 1. Unit cells of quaternary tin chalcohalides: (a) *Cmcm* $Sn_2SbS_2I_3$, (b) *Cmcm* $Sn_2BiS_2I_3$, and (c) *C*2/*m* Sn_2BiSI_5. (Note: $Sn_2SbS_2I_3$ displays both compositional and positional disorder with Sn and Sb,[21] while $Sn_2BiS_2I_3$ and Sn_2BiSI_5 display only compositional disorder with mixed occupancy sites of Sn and Bi.[17])

Inspired by recent work on the solution-phase synthesis of lead-based chalcohalides, we hypothesized that the solution-phase synthesis of both quaternary and ternary tin chalcohalides could be attained using thiocyanate precursors. Thiocyanates have proven to be useful in the preparation of both chalcogenides[41] and chalcohalides[42,43,44] due to the presence of a preformed metal-sulfur bond, which aids in controlling particle nucleation and growth.[43,44] However, even with this particular chemical strategy in hand, the synthetic phase space in which multinary chalcohalides form at relatively low temperatures (< 300 °C) from solution is expected to be complex due to numerous competing binary, ternary, and quaternary phases. Consequently, difficulties often arise in isolating the specific multinary compositions unless their phase evolution is thoroughly understood.

In this study, we examine the phase-specific synthesis of quaternary and ternary tin chalcohalides $Sn_2SbS_2I_3$, $Sn_2BiS_2I_3$, Sn_2BiSI_5, and Sn_2SI_2. We demonstrate that using different types of thiocyanate precursors affects their formation as well as the presence of lower order impurities. We employ ^{119}Sn solid-state (ss) NMR spectroscopy to probe the local tin environment in each semiconductor and further assess their phase purity. Importantly, experiments demonstrate that chalcohalides have excellent thermal and moisture stability compared to standard halide perovskites, further highlighting their suitability for device applications.

Results and Discussion.

Precursor screening: Where thiocyanate matters. Precursor selection is key to achieving phase and shape selectivity as well controlling particle size in solution.[45,46,47,48,49] Successful synthesis requires precursors and conditions that favor the nucleation and growth of the desired, multinary phase, often in a series of multiple steps,[50,51,52,53,54] while suppressing the nucleation and preventing the runaway growth of competing binaries or lower order phases and impurities. With this in mind, we resorted to metal thiocyanates as precursors to simultaneously introduce the desired Sn or Pn (Sb or Bi) and S elements. We specifically investigated supplying the thiocyanate along with the pnictide ($Bi(SCN)$[55]) or tin ($KSn(SO_4)SCN$[56] or $Sn(SCN)_2$[57]) precursors, with initial concentrations of 20–60 mM in a 10 mL mixture of 1-octadecene (ODE) and oleic acid (see Experimental).

Powder X-ray diffraction (XRD) of the solids isolated from the reaction of $Bi(SCN)_3$ with SnI_2 across a wide temperature range consist of ternary $Bi_{13}S_2I_{18}$,[31,32] without evidence for quaternary formation (Figure 2a); see Supporting Information available (SI). Given that SnI_2 appears to insufficiently deliver tin to the medium, we introduced tin acetate ($Sn(OAc)_2$) as an additional tin precursor. This approach succeeds in producing $Sn_2BiS_2I_3$ at 300 °C, although *ca.* 10% of the crystalline solids are still made of $Bi_{13}S_2I_{18}$ (Figure 2b). Prolonged reaction times at 300 °C failed to increase the yield of $Sn_2BiS_2I_3$; instead, the quaternary is quickly replaced by $Bi_{13}S_{18}I_2$ and, eventually, Bi metal. Therefore, while $Bi(SCN)_3$ serves as a precursor to quaternary, $Bi_{13}S_{18}I_2$ impurities are difficult to avoid. Because of this, and given the limited availability of $Sb(SCN)_3$, we decided to investigate supplying the thiocyanate with the tin precursor.

Reaction of $KSn(SO_4)SCN$[56] with SbI_3 results in $Sn_2SbS_2I_3$ that is free of crystalline impurities. However, scanning electron microscopy (SEM) and energy dispersive spectroscopy (EDS) reveal a significant amount of amorphous K_2S is present throughout the sample (see SI). This impurity, which is hard to detect by powder XRD alone, originates from the potassium-containing thiocyanate precursor. Interestingly, reaction of $KSn(SO_4)SCN$ with BiI_3 gradually produces the ternary bismuth chalcohalides BiSI and $Bi_{13}S_{18}I_2$ and, above 275 °C, quaternary $Sn_2BiS_2I_3$ (Figure 2c). However, a significant amount of BiSI is still present at 300 °C, even when SnI_2 is introduced as an additional tin precursor (Figure 2d). Therefore, while promising, $KSn(SO_4)SCN$ has limitations as a suitable precursor for synthesizing quaternary chalcohalides.

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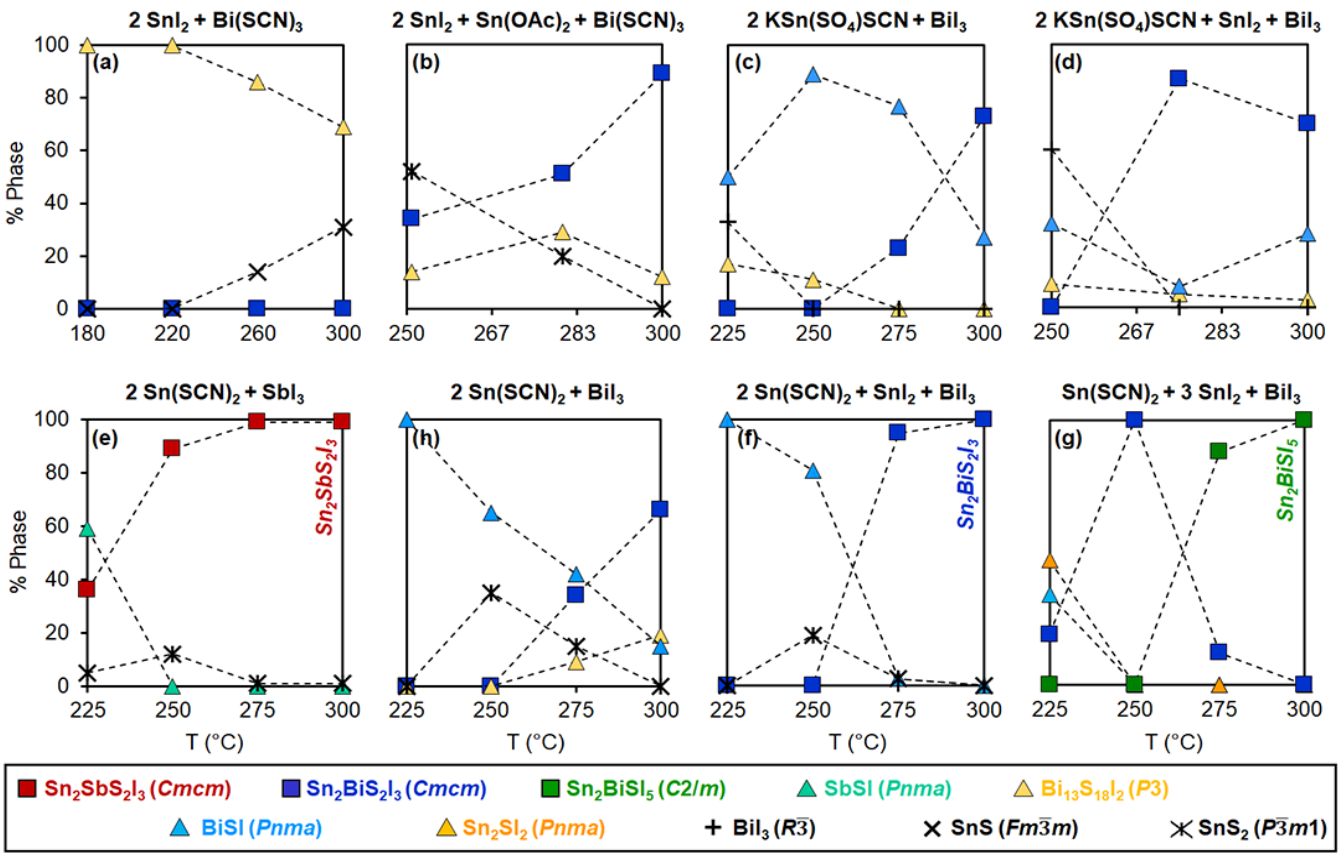

Figure 2. Phase evolution of quaternary (□) and ternary (△) tin chalcohalides using thiocyanate, halide, and acetate precursors under comparable reaction conditions (20–60 mM SCN precursor, 10 mL ODE + oleic acid; see Experimental).

By contrast to the aforementioned results, reaction of $Sn(SCN)_2$[57] with SbI_3 initially forms a mixture of binary, ternary, and quaternary phases below 250 °C, but cleanly transforms into phase pure $Sn_2SbS_2I_3$ (*Cmcm*) at 300 °C (Scheme 1, Figure 2e). Similarly, reaction between $Sn(SCN)_2$ and BiI_3 initially forms a mixture of BiSI and SnS_2, which then transforms into $Sn_2BiS_2I_3$—albeit, some $Bi_{13}S_{18}I_2$ forms above 275 °C (Figure 2f). Interestingly, using both $Sn(SCN)_2$ and SnI_2 along with BiI_3 enables the preparation of phase-pure $Sn_2BiS_2I_3$ (*Cmcm*) (Figure 2g).

Scheme 1. Solution-phase synthesis of tin-based, quaternary $Sn_2PnS_2I_3$ chalcohalides (Pn = Sb or Bi).

$$2Sn(SCN)_2 + PnI_3 + 6O_2 \xrightarrow[\text{ODE, oleic acid}]{\text{300 °C, 10 s}} Sn_2PnS_2I_3 + 2SO_2 + 4CO_2 + 2N_2$$

Pn = Sb or Bi

Synthesis of iodine-rich quaternary Sn_2BiSI_5. Further increasing the initial concentration of SnI_2 (20 mM) by two or three enables the preparation of a different, iodine-rich quaternary Sn_2BiSI_5 ($C2/m$) phase (Figure 2h). Increasing the oleic acid concentration to 2.5 M further facilitates the isolation of this different quaternary (Scheme 2), which may be attributed to increased solubility of SnI_2 in oleic acid (see Experimental and SI).

To further evaluate the importance of the thiocyanate precursor, we attempted to synthesize quaternary chalcohalides using a combination of metal carboxylates, elemental sulfur, and oleylammonium iodide (oleyl$NH_3^+I^-$, sometimes also referred to as 'I-OLAM').[58] Using the synthesis of $Sn_2BiS_2I_3$ as a test bed, we find that tin- and bismuth-carboxylates produce a mixture of $Bi_{13}S_{18}I_2$ and SnS (see SI). Using bismuth carboxylates and elemental sulfur while also introducing the more reactive SnI_2 instead of oleyl$NH_3^+I^-$ results in more quaternary formation, yet binary SnS or SnS_2 impurities persist. These results further demonstrate that a combined, thiocyanate/metal halide approach is among the most effective for synthesizing quaternary tin chalcohalide semiconductors.

Scheme 2. Solution-phase synthesis of additional tin-based, quaternary Sn_2BiSI_5 and ternary Sn_2SI_2 chalcohalides.

(a)

$$Sn(SCN)_2 + xSnI_2 + BiI_3 + 3O_2 \xrightarrow[\text{300 °C, 10 s}]{\text{ODE, oleic acid}} Sn_2BiSI_5 + SO_2 + 2CO_2 + N_2$$

x = 2 or 3

(b)

$$Sn(SCN)_2 + xSnI_2 + 3O_2 \xrightarrow[\text{300 °C, 10 s}]{\text{ODE, oleic acid}} Sn_2SI_2 + SO_2 + 2CO_2 + N_2$$

x = 1, 2, or 3

SEM imaging shows that $Sn_2SbS_2I_3$, $Sn_2BiS_2I_3$, and Sn_2BiSI_5 adopt highly anisotropic morphologies with high aspect ratios (l/w), such as rods and needles, such as those

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observed for their lead-based analogues (Figure 3).[43] The average rod widths range from *ca.* 700 nm to 2 µm, while the rod lengths range from 5 to 15 µm, yielding large particles. EDS confirms that the elemental composition of the phase-pure tin chalcohalides (Figure 3, Table 1).

Molecular basis for phase selectivity. A possible explanation for the relative reactivity and phase selectivity observed is based on the stability of the different metal-sulfur and metal-halide bonds that are present in the different molecular precursors.[59] The strongest among these is the Sn-S bond at 464 kJ/mol, followed by Sb-S and Bi-S bonds at 379 kJ/mol and 316 kJ/mol, respectively.[59] Stronger bonds are harder to make and harder to break, because both M-S bond forming and breaking processes involve a higher energy of activation.[60]

Indeed, our data show that to form quaternary tin chalcohalides, it is best to employ a precursor that already contains preformed Sn-S bonds and can deliver premade [SnS] building blocks to the reaction, such as $Sn(SCN)_2$. Instead, only ternary bismuth chalcohalides form when $Bi(SCN)_3$ is used instead. In comparison to M-S bonds, M-halide (M-X) bonds are relatively weak at Sn-I and Bi-I at 235 kJ/mol and 186 kJ/mol, respectively. Metal-halide bonds are relatively easy to make and break, thus a variety of M-I precursors are able to provide $I(SnI)_n$ (n = 0 (I^-), 1, 2 ...) units to cap and link [SnS] units during nucleation and growth.

Solution-grown ternary tin chalcohalides. Our synthetic approach succeeds in the preparation of ternary chalcohalides as well. Valuable targets include tin-based ternary chalcohalides which previously lacked a reported solution-phase synthesis.[38,39,40] Phase pure, orthorhombic (*Pnma*) Sn_2SI_2 forms at 300 °C by reacting together 20 mM $Sn(SCN)_2$ with a one-to-three fold excess of added SnI_2 (Scheme 2, Figure 4). Interestingly, during synthetic optimization we observed the disordered monoclinic ($C2/m$) Sn_2SI_2 polymorph in a highly concentrated (2.5 M) oleic acid solution with a four-fold excess of SnI_2. XRD showed mostly $C2/m$ Sn_2SI_2 (91%) with a small amount of the orthorhombic (*Pnma*) polymorph (Table 1, see SI).

This finding is unexpected because a higher SnI_2 concentration continues to favor the formation of Sn_2SI_2 polymorphs rather than the more iodine-rich composition Sn_4SI_6 ($C2/m$). SEM-EDS shows that the Sn_2SI_2 polymorphs adopt rod-like morphologies. Interestingly, the aspect ratio (l/w) of the particles decreases from ~25 to ~4 when the concentration of SnI_2 precursor used increases from one- to three-fold (20 to 60 mM) that of $Sn(SCN)_2$ (20 mM) (Figure 4). This coincides with a slight increase in the iodine-content compared to the theoretical ternary stoichiometry and demonstrates a possible entry for controlling the particle size and tuning the optoelectronic properties of Sn_2SI_2 and other chalcohalides.

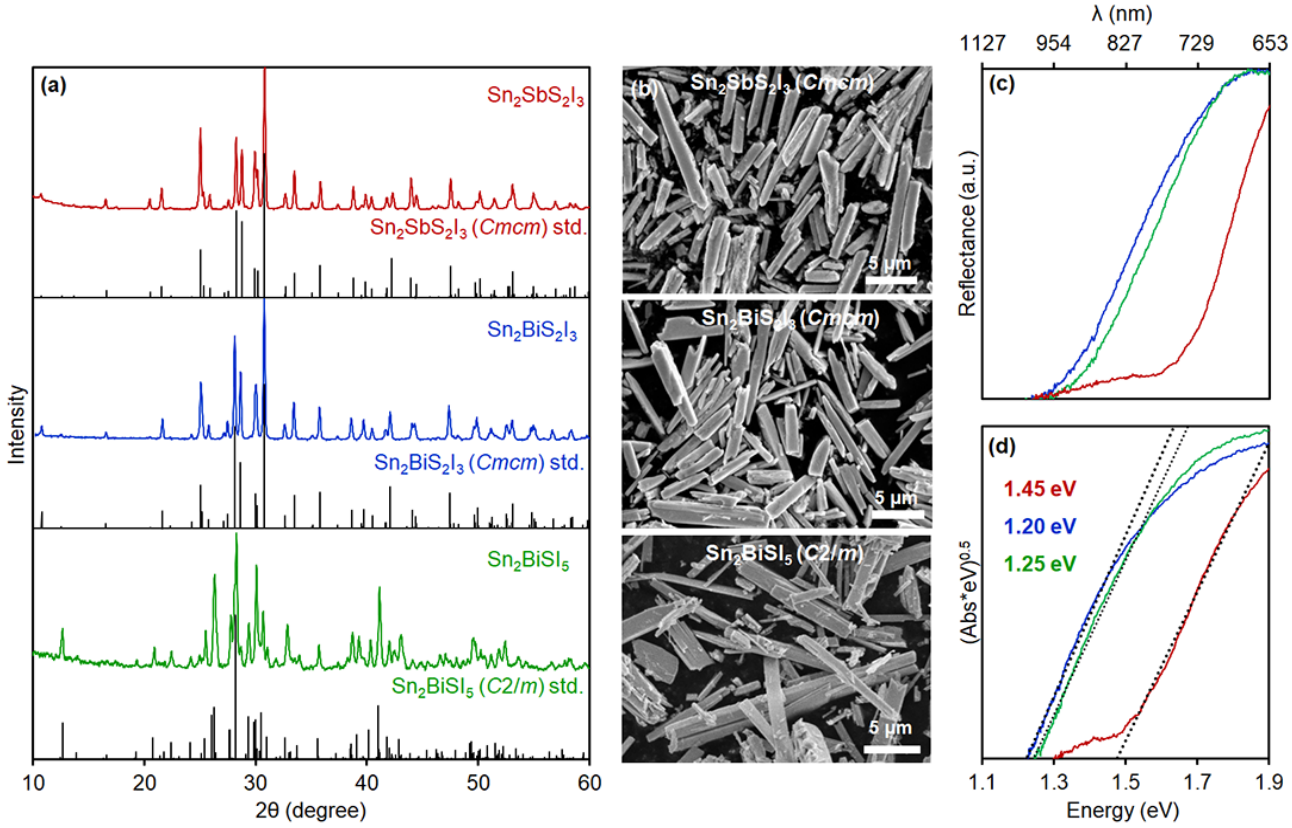

Figure 3. Powder XRD patterns (a), SEM images (b), diffuse reflectance spectra (c), and direct band Tauc plots of $Sn_2SbS_2I_3$, $Sn_2BiS_2I_3$, and Sn_2BiSI_5 (20–60 mM SCN precursor, 10 mL ODE + oleic acid; see Experimental).

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Table 1. Synthesis of quaternary and ternary tin chalcohalides.[a,b]

Sn prec(s). (mmol)	Pn prec. (mmol)	Ligands/ Solvents[a] (mL)	Product (%)[c]	Exp. gap Abs, CV / eV (nm)	Width[d] (w, µm)	Length[d] (l, µm)	Aspect ratio (l/w)
$Sn(SCN)_2$ (0.4)	SbI_3 (0.2)	ODE (8) OA (2)	$Sn_2SbS_2I_3$ (*Cmcm*)	1.45 (855), 1.55	1.5±0.5	9.7±2.5	6
$Sn(SCN)_2$ (0.4) SnI_2 (0.2)	BiI_3 (0.2)	ODE (8) OA (2)	$Sn_2BiS_2I_3$ (*Cmcm*)	1.20 (1033), 1.32	0.7±0.2	6.7±2.1	10
$Sn(SCN)_2$ (0.2) SnI_2 (0.6)	BiI_3 (0.2)	ODE (2) OA (8)	Sn_2BiSI_5 (*C*2/*m*)	1.25 (992), 1.25	1.2±0.4	14±6.0	12
$Sn(SCN)_2$ (0.2) SnI_2 (0.2)	none	ODE (8) OA (2)	Sn_2SI_2 (*Pnma*)	2.11 (588)	0.68 ± 0.15	17.5 ± 6.1	25
$Sn(SCN)_2$ (0.2) SnI_2 (0.4)	none	ODE (8) OA (2)	Sn_2SI_2 (*Pnma*)	2.03 (605), 1.97	1.3 ± 0.33	19 ± 6.4	15
$Sn(SCN)_2$ (0.2) SnI_2 (0.6)	none	ODE (8) OA (2)	Sn_2SI_2 (*Pnma*)	1.92 (646)	3.2 ± 1.1	14 ± 4.6	4
$Sn(SCN)_2$ (0.2) SnI_2 (0.8)	none	ODE (2) OA (8)	Sn_2SI_2 (*C*2/*m*, 91%) Sn_2SI_2 (*Pnma*, 9%)	1.96 (633)	*2.5*±1.1	15±4.7	6

[a]General conditions: T = 300 °C, t = 30 sec. [b]Total reaction volume = 10 mL. [c]From Match! [d]From SEM. (See Experimental.)

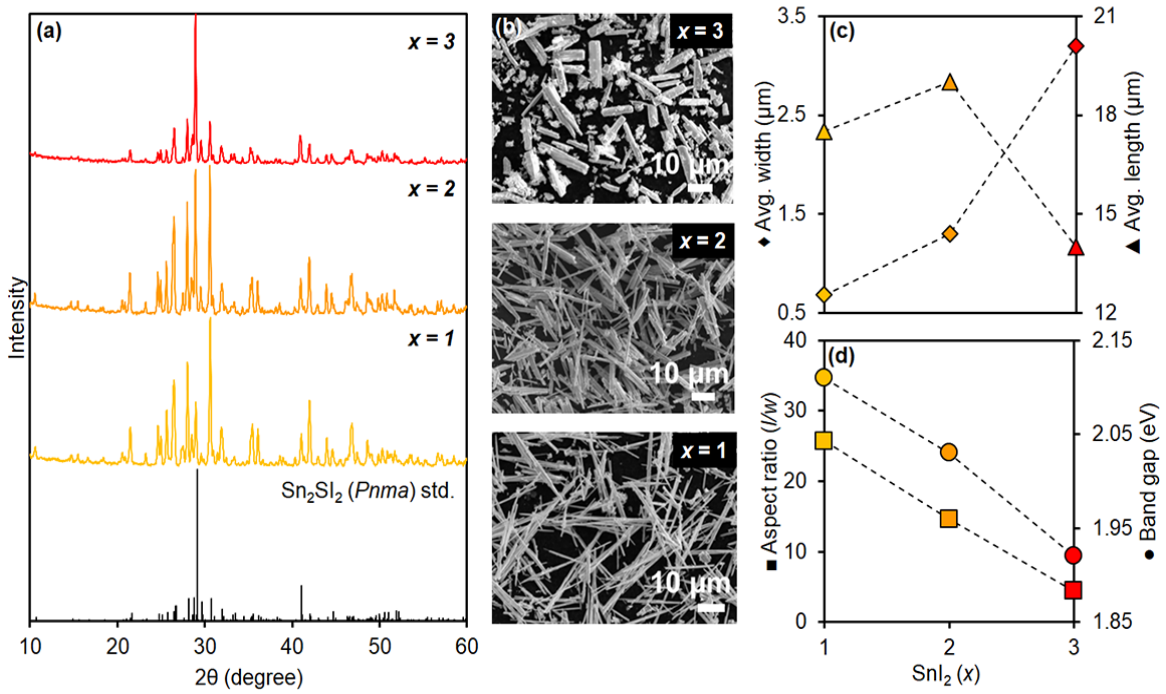

Figure 4. Powder XRD patterns (a), SEM images (b), particle dimensions (c), and band gap to aspect ratio comparison (d) of solution-grown *Pnma* Sn_2SI_2 (20–80 mM Sn, 10 mL ODE + oleic acid; see Experimental).

Opto-electronic characterization. Previous studies indicate that $Sn_2PnS_2I_3$ (Pn = Sb or Bi) chalcohalides are direct band gap semiconductors. Here, we evaluated optical gaps of the solution-grown materials optically *via* diffuse reflectance spectroscopy, and electrochemically using cyclic voltammetry. Direct gap Tauc plots[61] derived from the diffuse reflectance spectra yield 1.45 eV (855 nm), 1.20 eV (992 nm), and 1.25 eV (1033 nm) gaps for $Sn_2SbS_2I_3$, $Sn_2BiS_2I_3$, and Sn_2BiSI_5, respectively (Figure 3). These values closely agree with other experimental and calculated values and further support the suitability of these materials as semiconductors for light harvesting devices.[17]

Unlike its quaternary counterparts, Sn_2SI_2 is predicted to be an indirect band gap semiconductor.[38,39] Diffuse reflectance measurements reveal a band of 1.96 eV (633 nm) for the monoclinic (*C*2/*m*) Sn_2SI_2 polymorph (see SI). For the orthorhombic (*Pnma*) polymorph, the band gap gradually red shifts from 2.11 to 1.92 eV (588–646 nm) as the concentration of SnI_2 used in the synthesis increases (Figure 4). This change in optical properties is directly proportional to the decreasing aspect ratio (*l*/*w*) of the Sn_2SI_2

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particles observed from SEM. However, it is unlikely that this phenomenon is related to quantum confinement effects due to the large particle dimensions. Nonetheless, this shows that it is still possible to synthetically tune both the particle size and the optical properties of Sn_2SI_2.

Cyclic voltammograms (CVs) were obtained for the quaternary and ternary tin chalcohalides to better assess their electronic band alignment in addition to their band gap energies. The anodic (E_{pa}) and cathodic (E_{pc}) peak potential values were derived from well-defined oxidation (E_{ox}) and reduction (E_{red}) waves in the CV (with respect to a saturated calomel electrode), and used to extrapolate the HOMO-LUMO energies using relationships:

$$E(HOMO) = -I_p = -(E_{ox} + 4.20)\ eV$$

$$E(LUMO) = -E_a = -(E_{red} + 4.20)\ eV$$

The electrochemically determined ionization potentials (*Ip*) and electron affinities (*Ea*) for the tin chalcohalides reveal narrowing band gap values from 1.97 to 1.55, 1.32, and 1.25 eV for Sn_2SI_2, $Sn_2SbS_2I_3$, $Sn_2BiS_2I_3$, and Sn_2BiSI_5, respectively (Figure 5). These agree well with those obtained from diffuse reflectance and are consistent with general trends in the lead chalcohalide series.[43]

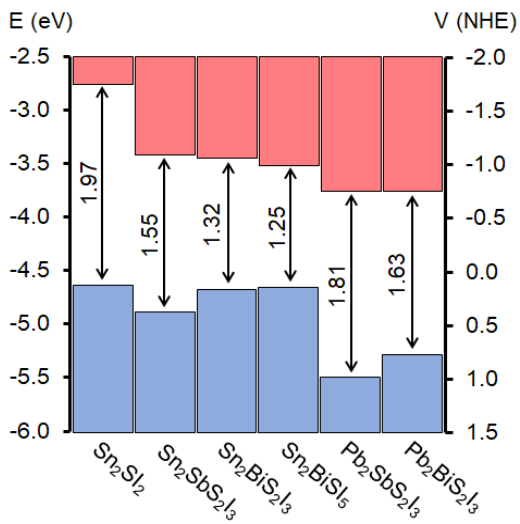

Figure 5. Electrochemically determined band energy levels and alignment of tin and lead[43] chalcohalide semiconductors (see Experimental).

^{119}Sn ssNMR spectroscopy. Direct excitation ^{119}Sn solid-state NMR spectroscopy was employed in this study to investigate the local environment of Sn. ^{119}Sn is a spin ½ nucleus that has a natural isotopic abundance of 8.6% and is known to possess an isotropic chemical shift range of 1000 ppm to –2000 ppm for diamagnetic compounds, making ^{119}Sn NMR spectra extremely sensitive to changes in Sn local environment.[62] ^{119}Sn NMR spectra of the precursor materials $Sn(SCN)_2$ and $KSn(SO_4)(SCN)$ were also measured (see SI). These spectra contain narrow peaks and have an isotropic chemical shift of -812 and -978 ppm, respectively. Neither peak is seen in the Sn chalcohalides, indicating that the reactants were fully consumed or converted to other products. During synthetic development, a common amorphous impurity that was observed by ^{119}Sn NMR—but not by powder XRD—was SnS_2.[63]

All of the tin iodide materials studied here give rise to broad ^{119}Sn NMR signals. For comparison, quaternary lead chalcohalides containing mixed-halides (*i.e.*, Pb_3SBrI_3) also exhibit a single broad ^{207}Pb ssNMR signal despite having multiple Pb environments in their structure.[44] Kubicki and co-workers reported ^{119}Sn solid-state NMR spectra of mixed halide Sn perovskites.[64] The compound perovskite methylammonium tin iodide ($CH_3NH_3SnI_3$) gives rise to a broad ^{119}Sn NMR signal that has a peak width of *ca.* 26 kHz. We observed similarly broadened ^{119}Sn NMR spectra here (Figure 6), with peak widths on the order of *ca.* 70 kHz at 9.4 T and *ca.* 100 kHz at 14.1 T—the origin of the broadening is further discussed below.

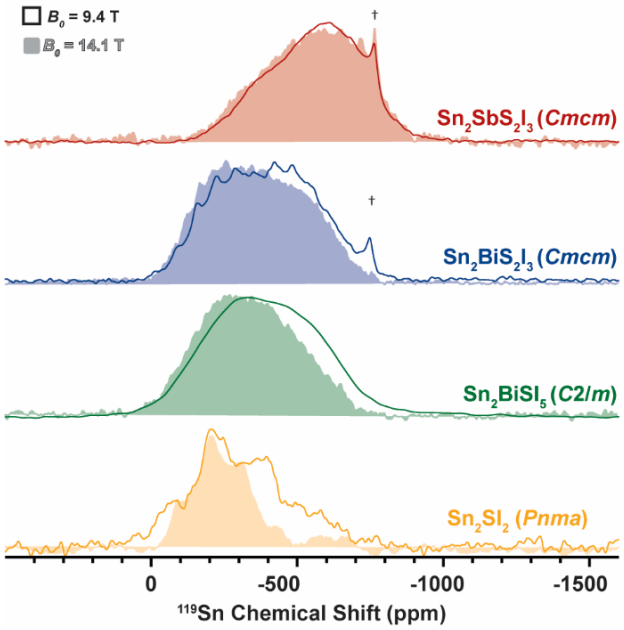

Figure 6. ^{119}Sn spin echo solid-state NMR spectra of $Sn_2SbS_2I_3$, $Sn_2BiS_2I_3$, Sn_2BiSI_5, and Sn_2SI_2 at 9.4 T (lines) and 14.1 T (shaded). 9.4 T NMR spectra were collected using an MAS frequency of 12.5 kHz for $Sn_2SbS_2I_3$, $Sn_2BiS_2I_3$, and Sn_2BiSI_5. Spectra collected at 14.1 T were measured with a 10 kHz MAS spinning frequency. A 25 kHz MAS frequency was used for Sn_2SI_2 at both fields. († indicates a possible SnS_2 impurity at -762 ppm.[63])

The simplest compound studied here, $Sn_2SbS_2I_3$, gives rise to a ^{119}Sn NMR signal that was 65 kHz broad at 9.4 T, despite the fact that only a single Sn site was reported in the original crystal structure of this material.[28] A comparison of the ^{119}Sn NMR spectra of $Sn_2SbS_2I_3$ obtained at 9.4 T and 14.1 T shows that the spectra are similar in appearance when plotted in units of ppm. This observation suggests that the broadening of the ^{119}Sn NMR spectrum is primarily inhomogeneous. Inhomogeneous broadening occurs from structural disorder that leads to a distribution in the isotropic chemical shift and/or from the effects of chemical shift anisotropy (CSA). Fitting the ^{119}Sn NMR

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spectra of $Sn_2SbS_2I_3$ to a single Sn site suggests the span (Ω) is *ca.* 360 ppm. In order to fit the spectrum to a single Sn site, we used 29 kHz and 46 kHz of Gaussian line broadening at 9.4 T and 14.1 T, respectively, which accounts for inhomogeneous broadening. If the span is 360 ppm, the fits suggest that part of the ^{119}Sn NMR spectrum could correspond to spinning sidebands. The fit suggests that the true isotropic chemical shift is *ca.* -547 ppm. For both fields, the fits used a constant Lorentzian broadening of 5 kHz (see SI). The amount of Lorentzian broadening used in the fit was estimated from a measurement of the homogenous ^{119}Sn transverse relaxation time constant (T_2') at 9.4 T, which suggests T_2' is less than 200 μs. Similarly short T_2' were observed for the other Sn compounds. The homogeneous broadening likely occurs due to coupling to ^{127}I, which is a spin 5/2 quadrupolar nuclei that is 100% naturally abundant. It has been shown before that when a heavy spin ½ nucleus such as ^{119}Sn or ^{207}Pb is dipole and/or scalar coupled to an abundant quadrupolar nucleus such as ^{127}I, there can be efficient transverse and longitudinal relaxation.[65,66] The efficient relaxation of the spin ½ nucleus occurs because some combination of rapid longitudinal relaxation of the quadrupolar nucleus or chemical exchange (diffusion) of the iodide anions.[44,64,66]

Thus, acquisition of the ^{119}Sn NMR spectra of $Sn_2SbS_2I_3$ at 9.4 T and 14.1 T confirms that the NMR spectrum of $Sn_2SbS_2I_3$ is primarily inhomogeneously broadened due to an isotropic chemical shift distribution and partly due to spinning sideband overlap. The observed inhomogeneous broadening is consistent with the presence of significant structural disorder, perhaps giving support to recent crystallographic models where $Sn_2SbS_2I_3$ exhibits multiple S and I positions as well as Sn and Sb site mixing.[21]

Moving down to the heavier pnictide (Pn), the crystal structure of to $Sn_2BiS_2I_3$ shows two unique Sn sites, one of which exhibits mixed Pn occupancy (50% Sn:50% Bi). Similar ^{119}Sn NMR linewidths were observed for $Sn_2SbS_2I_3$ and $Sn_2BiS_2I_3$. The inhomogeneous broadening persists in the ^{119}Sn NMR spectrum of Sn_2BiSI_5, which contains the largest number of Sn sites, including two mixed occupancy sites (24% Sn:76% Bi and 76% Sn:24% Bi). For Sn_2BiSI_5, the peak width was approximately 75 kHz at 9.4 T.

Considering the work of Kubicki on tin halide perovskites and related phases, perovskites generally showed more positively shifted ^{119}Sn NMR signals as the amount of iodine in the lattice was increased. A similar trend is observed here, where the more iodine rich phases tend to exhibit ^{119}Sn NMR signals shifted to higher (more positive) frequencies.

The ^{119}Sn NMR spectrum of Sn_2SI_2 gives rise to narrower NMR signals. In this compound, the ^{119}Sn T_1 is approximately four times longer (T_1=30 s) than the ^{119}Sn T_1 of the other compounds (T_1=8 s), suggesting that relaxation due to coupling with ^{127}I is not as efficient. The 14.1 T ^{119}Sn NMR spectrum of Sn_2SI_2 shows narrowing (when plotted in ppm) as compared to the 9.4 T spectrum. This observation suggests that the broadening of the ^{119}Sn NMR spectra is primarily homogeneous in nature. Compared to the aforementioned, pnictide-containing quaternary compositions, which show either positional or compositional disorder (or even a combination of the two), *Pnma* Sn_2SI_2 is completely ordered, with six different Sn sites.[39] Three of these Sn sites show coordination environments of SnS_3I_5, while the remaining sites adopt coordination environments of $SnSI_5$, $SnSI_7$, and $SnSI_6$, respectively (see SI). Interestingly, only three, significantly less broad, NMR peaks are present in the NMR spectrum, suggesting that Sn_2SI_2 is more crystalline than the other tin chalcohalides.

Superior moisture and thermal stability. To better assess the durability of tin-based chalcohalides for various electronic applications, we exposed them to excess water and compared their moisture stability to freshly made samples of halide perovskites as reference materials.[67,68] Briefly, we vigorously stirred each chalcohalide or perovskite sample in water for two days (48 h). Remarkably, the powder XRD of the tin chalcohalides lacks any signs of degradation or formation of crystalline impurities after exposure to water (Figure 7). In contrast, both tin- and lead-based halide perovskites show major signs of decomposition after the same water treatment. Specifically, $CsPbI_3$ decomposes to PbI_2 over 48 h, while $CsSnI_3$ visibly decomposes to a completely amorphous, white-colored product within the first 1 h in water. This highlights one of the major common challenges with other tin-based semiconductors, such as tin halide perovskites, which are susceptible to moisture-induced decomposition—often involving the oxidation of Sn^{2+} to Sn^{4+}.[75]

Lastly, we studied the thermal stability of the tin chalcohalides using thermogravimetric analysis (TGA) and differential scanning calorimetry (DSC) under different environments. Under an inert (N_2) atmosphere, the onset of decomposition for all materials occurs between 400–450 °C (Figure 7). Among the quaternaries, general stability appears to increase with increasing formula weight, specifically on going from $Sn_2SbS_2I_3$ to the heavier pnictide-containing quaternary compositions $Sn_2BiS_2I_3$ and Sn_2BiSI_5. Notably, ternary Sn_2SI_2 exhibits a decomposition onset around 500 °C, the highest out of the tin chalcohalide series in our study. Under air, the thermal decomposition of the chalcohalides occurs at slightly lower temperatures, with mass loss beginning around 250 °C. This data are strikingly different from the thermal profile of halide perovskites: While there are a number of factors that influence their thermal stability—*i.e.*, cation- and halide-composition, particle size, surface treatment, etc. —many halide perovskites decompose between 200–400 °C under an inert atmosphere,[69,70,71,72,73,74] and as low as 180 °C under air.[75] Therefore, combined with their superior moisture stability mentioned above, tin chalcohalides are a much more robust and stable alternative to perovskites as materials for energy conversion and catalytic devices.

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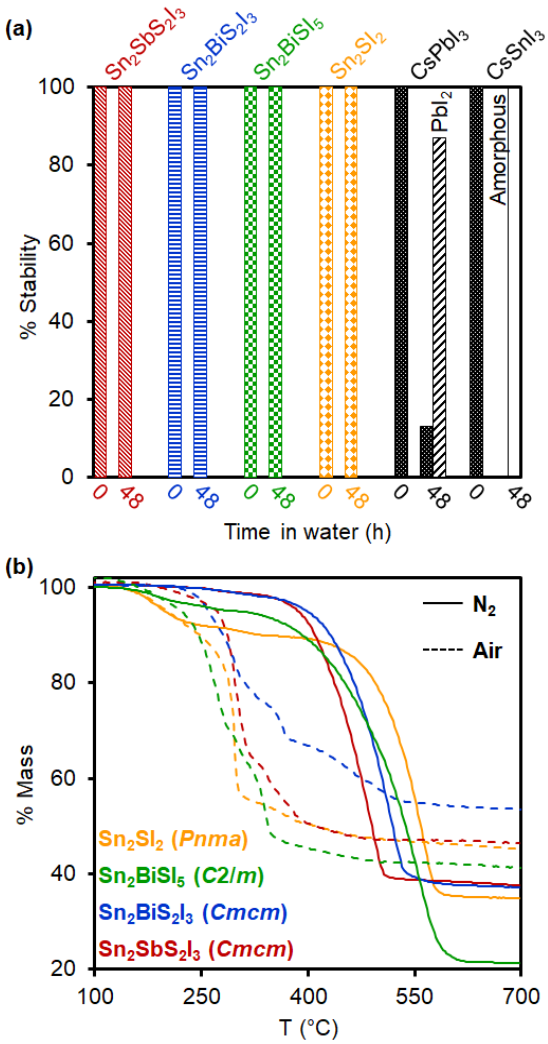

Figure 7. (a) Stability of tin chalcohalides stirred in water for 48 h (21 °C) compared to $CsSnI_3$ and $CsPbI_3$ perovskites.[67,68] (b) TGA analysis of tin chalcohalides under N_2 (solid line) or air (dashed).

Conclusions.

This work advances the fundamental synthetic chemistry of lead-free chalcohalides starting from readily available and air-stable precursors and conditions. Understanding the complex phase space associated with quaternary compositions reveals optimal conditions to prepare specific chalcohalides while avoiding impurities. Despite the fact that each of the thiocyanate precursors tested provides a source of sulfur and a cation (Sn or Sb/Bi), each elicits the phase evolution of various binary, ternary, and quaternary products. The formation of each quaternary chalcohalide is often preceded by the formation of lower order binary and ternary products. ^{119}Sn ssNMR helps distinguish between different quaternary chalcohalide phases and detect trace amounts of amorphous impurities that are otherwise silent by XRD or hard to detect by SEM-EDS. In addition to their direct optical band gap character, the chalcohalides display great resilience to moisture and heat, which is favorable for semiconductors intended for numerous applications including photovoltaics, thermoelectrics, and catalysis.

Future studies will expand upon this knowledge to synthesize colloidal nanoscale versions of these materials that are more conducive to thin film deposition and device fabrication. Additionally, other multinary tin chalcohalide compositions such as $Sn_7Br_{10}S_2$ and $CdSnSX_2$ (X = Cl or Br) exhibit properties that are highly suitable for either nonlinear optics (NLO)[76] or photocatalysis,[77] respectively, and their colloidal chemistry remains underexplored. Additionally, halide-mixing and chalcogen-mixing are yet to be studied in tin chalcohalides, potentially offering a wider range of compositional, structural, and optical tunability for a variety of applications.[16]

Experimental.

Materials. 1-Octadecene (ODE, technical grade, 90%), potassium thiocyanate (KSCN, 99%) sodium thiocyanate (NaSCN, 98%), and tin(II) acetate ($Sn(OAc)_2$) from Sigma Aldrich; antimony(III) iodide (SbI_3, 99.9%), bismuth(III) iodide (BiI_3, 99.999%), tin(II) iodide (SnI_2, 99%), tin(II) sulfate ($SnSO_4$, 95%), and bismuth(III) chloride ($BiCl_3$, 99%) from Strem; oleic acid (technical grade, 90%) from Alfa Aesar; hexanes (99.9%) and methanol (99.9%) from Fisher. All chemicals were used as received without further purification.

Synthesis of thiocyanate precursors. $Bi(SCN)_3$, $KSn(SO_4)SCN$, and $Sn(SCN)_2$ were prepared according to literature methods. *$Bi(SCN)_3$.* Inside a N_2-filled drybox, $BiCl_3$ (534 mg, 1.7 mmol), KSCN (576 mg, 4.7 mmol), and anhydrous THF (60 mL) were added to a round bottom flask (R.B.) and stirred for 18 h at room temperature (R.T., 21 °C). After filtration to remove the formed precipitate (KCl), the supernatant was evaporated under vacuum to yield bright orange crystals (Yield: 0.291 g, 44%).[55] *$KSn(SO_4)SCN$.* A solution of $SnSO_4$ (4 g, 19 mmol) and concentrated H_2SO_4 (3 drops) in deionized water (25 mL) was filtered into a solution of KSCN (4 g, 41 mmol) in deionized water (12.5 mL). After sitting at 0 °C for 24 h, the white precipitate was collected by filtration, washed with deionized water, and dried over vacuum (Yield: 1.627 g, 27%).[56] *$Sn(SCN)_2$.* A solution of $SnSO_4$ (4 g, 19 mmol) and 1.0 M H_2SO_4 (4 mL) in deionized water (21 mL) was filtered into a solution of NaSCN (6 g, 74 mmol) in deionized water (25 mL). After 5 h at R.T., then the mixture was left for 5 h at 0 °C. White crystals were isolated by filtration, washed with deionized water, and dried under vacuum (Yield: 1.383 g, 31%).[57]

Synthesis of chalcohalides. All syntheses were performed in air under standard ambient conditions. *Quaternaries: $Sn_2SbS_2I_3$.* A mixture of $Sn(SCN)_2$ (94 mg, 0.4 mmol), SbI_3 (100 mg, 0.2 mmol), ODE (8 mL, 25 mmol), and oleic acid (2 mL, 6.3 mmol) was stirred in a three-neck R.B. flask for *ca.* 5 min at 140 °C until all solids dissolved or suspended. The mixture was heated to 300 °C, let sit for 30 s, and allowed to cool to R.T. by removing the heating mantle. *$Sn_2BiS_2I_3$.* Prepared in a similar manner using a mixture of $Sn(SCN)_2$ (94 mg, 0.4 mmol), BiI_3 (118 mg, 0.2 mmol), SnI_2 (75 mg, 0.2 mmol), ODE (8 mL, 25 mmol), and oleic acid (2 mL, 6.3 mmol). *Sn_2BiSI_5.* Prepared in a similar

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manner using a mixture of $Sn(SCN)_2$ (47 mg, 0.2 mmol), SnI_2 (0.4 or 0.6 mmol, see discussion), BiI_3 (118 mg, 0.2 mmol), ODE (2 mL, 6.3 mmol), and oleic acid (8 mL, 25 mmol). *Ternaries: Sn_2SI_2 (Pnma).* A mixture of $Sn(SCN)_2$ (47 mg, 0.2 mmol), SnI_2 (0.2, 0.4, or 0.6 mmol, see discussion) ODE (8 mL, 25 mmol), and oleic acid (2 mL, 6.3 mmol) was stirred in a three-neck R.B. flask at 140 °C until all solids dissolved or suspended. The mixture was heated to 300 °C for 30 s, and allowed to cool to R.T. by removing the heating mantle. *Sn_2SI_2 (C2/m).* Prepared in a similar manner using SnI_2 (300 mg, 0.8 mmol), with OA (8 mL, 25 mmol), and ODE (2 mL, 6.3 mmol). *Purification.* Crude chalcohalide solutions were initially centrifuged at 4500 rpm for 5 min. After discarding the supernatant, the solids were resuspended in hexanes (5 mL), re-precipitated with methanol (5 mL) and centrifuged again. This process was repeated at least three times or until the supernatant was colorless.

Optical characterization. Diffuse-reflectance spectra were collected using a SL1 Tungsten Halogen lamp (vis–IR), a SL3 Deuterium Lamp (UV), and a BLACK-Comet C-SR-100 spectrometer (200–1080 nm). Band gap values were estimated by extrapolating the linear slope of Tauc plots of $(Ah\nu)^r$ versus $h\nu$ (A = absorbance, $h\nu$ = incident photon energy in eV, with r = 1/2 for direct and r = 2 for indirect semiconductors).[61]

Structural characterization. Powder X-ray diffraction (XRD) was measured on a Rigaku Ultima IV diffractometer (40 kV, 44 mA) using Cu Kα radiation on a zero-background quartz sample holder. Scanning electron microscopy (SEM) images were acquired on a JEOL JSM-IT200 scanning electron microscope.

Cyclic voltammetry. Electrochemical studies were carried out using cyclic voltammetry (CV) with the potentiostat Workstation CHInstrument that was controlled by the software CHI660e V14.08. A three-electrode system and a thermostatic electrochemical cell were used for all measurements. Glassy carbon electrodes were used as the working (WE, 3 mm diameter) and counter (CE, 3 mm diameter) electrodes, and a saturated calomel electrode (SCE) was used as reference electrode (RE). The working electrode was polished with 1 mm diamond paste (DP-Paste, P) and rinsed with ethanol after each set of measurements. The semiconductors were each dispersed in n-hexanes (1 mg/mL) followed by sonication for 1-2 minutes. Afterwards, 50 μL of each sample was drop-cast onto the glassy carbon electrode surface and the hexanes was fully evaporated. Prior to recording the CVs for the chalcohalides, blank experiments were conducted to ensure that the working electrode was polished and clean. No electrochemical signals were detected for non-modified working electrodes in the selected potential range (from -2.5 to 1.5 V). CV experiments were performed under an inert atmosphere using the modified working electrode in a pure acetonitrile + 0.1 M tetrabutylammonium hexafluoroborate (TBA BF_6,) solution at different scan rates (v = 0.1 to 0.7 V s^{-1}).

Solid state ^{119}Sn NMR spectroscopy. Solid state NMR experiments were performed at two different fields (B_0 = 9.4 T and 11.4 T. ^{119}Sn was referenced indirectly to ^{1}H in a mixture of tetramethylsilane in $CDCl_3$ and adamantane using IUPAC recommend frequencies.[78] NMR experiments at 9.4 T (400 MHz) were performed with a wide-bore magnet equipped with a Bruker AVANCE III HD console. NMR spectra were recorded on a 4.0 mm HXY MAS probe at 12.5 kHz MAS. ^{119}Sn pulse length for π/2 and π pulses were 1.75 μs and 3.5 μs respectively, corresponding to a 143 kHz RF field. Faster MAS frequencies was needed for experiments on Sn_2SI_2, thus a 2.5 mm HXY MAS probe was used. The ^{119}Sn pulse length for π/2 and π pulses were 1.92 μs and 3.84 μs respectively, corresponding to an *ca.* 130 kHz Rf field. Experiments at 14.1 T (600 MHz) were performed with a Bruker wide-bore magnet equipped with a Bruker AVANCE NEO console. Experiments were recorded on a 4.0 mm HXY MAS probe at 10 kHz MAS. π/2 and π pulse lengths were 3.02 μs and 6.04 μs, corresponding to an *ca.* 83 kHz Rf field. The ^{119}Sn NMR spectrum of Sn_2SI_2 was measured on a 2.5 mm HXY MAS probe with a pulse length of 1.96 μs and 3.92 μs for π/2 and π pulses. The ^{119}Sn Rf field was *ca.* 128 kHz. All experiments were performed with the probes configured in double resonance (HX) mode to maximize sensitivity.

Thermal analysis. TGA and DSC were carried out on chalcohalide samples (5–20 mg) using a Netzsch DSC/TGA (STA449 F1) and alumina (Al_2O_3) crucibles. The temperature program was a 10°C/min ramp from 40–700 °C under either a nitrogen or air atmosphere.

ASSOCIATED CONTENT

Supporting Information. Phase evolution plots, XRD, SEM-EDS, diffuse reflectance, cyclic voltammograms, ssNMR, and TGA/DSC are provided in the supporting information. This material is available free of charge via the Internet at http://pubs.acs.org."

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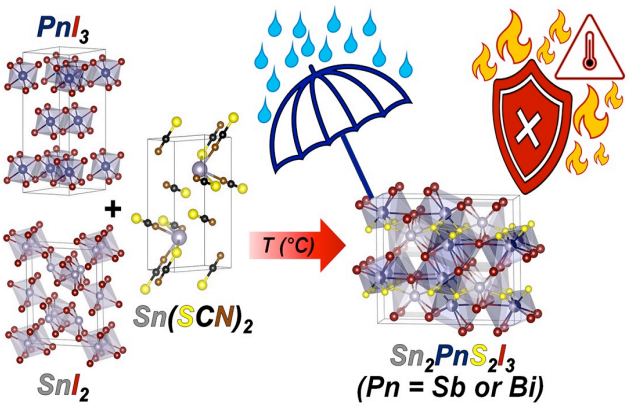
TOC Figure



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