In-Silico-Assisted Derivatization of Triarylboranes for the Catalytic Reductive Functionalization of Amino Acids with H_2

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ABSTRACT: Cheminformatics-based machine learning (ML) has assisted in determining optimal reaction conditions, including catalyst structures, in the field of synthetic chemistry. However, such ML-focused strategies have remained largely unexplored in the context of catalytic molecular transformations using Lewis-acidic main-group elements, probably due to the absence of a candidate library and effective guidelines (parameters) for the prediction of the activity of main-group elements. Here, the construction of a triarylborane library and its application to an ML-assisted approach for the catalytic reductive alkylation of amino acids with aldehydes and H_2 is reported. The obtained results suggest that the deformation energy serves as a useful parameter for Gaussian process regression to construct adequate models for predicting the turnover frequencies of triarylboranes under the applied model reaction conditions, while Gaussian progress regression based on the energy levels of the lowest unoccupied molecular orbitals (LUMOs) including the empty p-orbitals of boron tend to underestimate the turnover frequency. The optimal borane, i.e., B(2,3,5,6-Cl_4-C_6H)(2,6-F_2-3,5-$(CF_3)_2$-$C_6H)_2$, effectively catalyzes the reductive functionalization of aniline-derived amino acids and C-terminal-protected peptides in the presence of 4-methyltetrahydropyrane and H_2, generating H_2O as the sole byproduct.

Introduction

Catalysis is a fact our daily lives. A wide variety of important commercial chemical substances are currently produced on both the fine and bulk scales in the presence of molecular catalysts that have been optimized based on specific factors such as efficiency, toxicity, cost, or a combination thereof. Recent advancements in cheminformatics-based machine learning (ML) offer chemists a way to bypass traditional Edisonian empiricism and develop more efficient approaches to optimizing catalysts.[1–4] Several groups have reported successful demonstrations of ML-driven optimizations of homogeneous catalysts such as phosphoric acids[5–8] and Lewis-basic ligands for metal-based catalysts involving phosphines, *N*-heterocyclic carbenes, and nitrogen-based ligands.[9–18]

Recent progress in frustrated Lewis pairs (FLPs)[19,20] has expanded the practical and sustainable application of main-group catalysis, e.g., enabling the hydrogenation of unsaturated molecules without toxic/precious metals.[21–26] In this context, the main-group-catalyzed reductive alkylation of amines with carbonyl compounds and H_2 via the generation of FLP species has been widely accepted as a waste-minimizing process that generates valuable *N*-alkylated amines, whereby H_2O is the only by-product.[23,27–31] Our group[30] and that of Soós[28] have independently shown that triarylboranes effectively catalyze the reductive alkylation of a variety of amines with aldehydes in the presence of H_2. Moreover, we have demonstrated that an FLP that consists of Soós' borane, i.e., B(2,6-Cl_2-C_6H_3)(2,3,5,6-F_4-$C_6H)_2$ (**B**$^{\mathbf{1a}}$),[27] and tetrahydrofuran (THF) exhibits good functional-group tolerance for aniline derivatives including halogens, hydroxyl, and amide groups (Figure 1).[30] However, the direct reductive alkylation of amino acids with H_2 has remained challenging, and such reactions have proceeded in only low-to-moderate yields even under forcing conditions. In terms of toxicity, the solvent THF, which also acts here as a Lewis base to generate FLPs with boranes, should be replaced with a less hazardous chemical.[32,33] Given the central role of the reductive alkylation of amines using carbonyl compounds in the synthesis of e.g. pharmaceuticals, bio-active molecules, and agrochemicals,[34–36] the development of a straightforward and greener protocol for derivatizing amino acids and peptides would be worthwhile.

To this end, we envisioned an ML-assisted approach to identify a suitable triarylborane that is able to efficiently catalyze the reductive alkylation of amino acids and peptides with H_2 through the construction of an in-silico library that includes a variety of unprecedented triarylboranes. Moreover, through the construction of this *in-silico* library, we aimed to contribute to the structural diversification of triarylboranes beyond the archetypical B($C_6F_5)_3$,[37–40] which should expand the utility of this compound class in catalysis, materials science, and other areas. It should be noted that Dyson and Corminboeuf et al. have

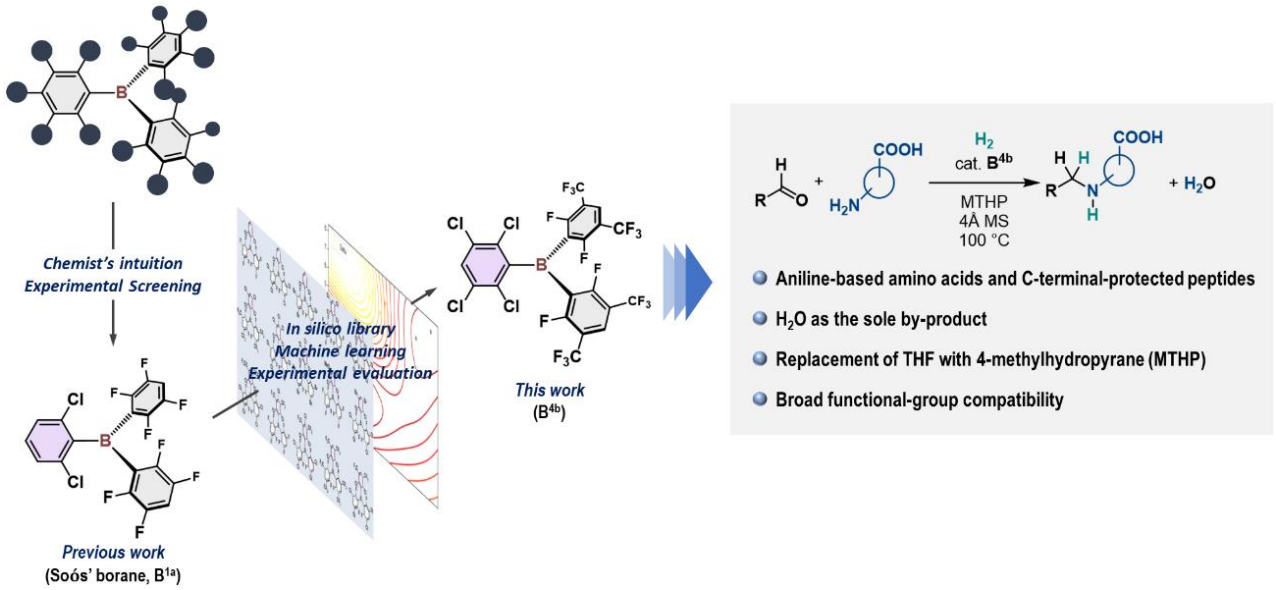

Figure 1. Schematic representation of the concept of this study.

recently demonstrated the hydrogenation of CO_2 to yield a formate salt [DBU–H][H–COO], which was catalyzed by an FLP consisting of tris(p-bromo)tridurylborane and 1,8-diazabicyclo[5.4.0]undec-7-ene (DBU). They identified this combination of borane and DBU using a cheminformatics-assisted approach that profiled the theoretically predicted catalyst activity based on the intrinsic acidity and basicity of the Lewis components.[41] Herein, we report the construction of an *in-silico* library with 54 triarylboranes, including 46 unprecedented structures, which was used for the ML-assisted identification of the optimal borane for the catalytic reductive alkylation of amino acids and peptides with aldehydes and H_2 (Figure 1). We also explored the functional-group compatibility of the present system using the functional group evaluation (FGE) kit recently proposed by Morimoto and Oshima et al.,[42] which is based on the concept of robustness screening that has been proposed by Glorius et al.[43,44]

Results and discussion

We started our investigation with the construction of an *in-silico* library of triarylboranes using the strategy described below. Generally, we explored triarylboranes that seemed to be synthetically accessible using common procedures.[37,40] Optimization of the gas-phase structures of the 54 boranes shown in Figure 2A was accomplished using DFT calculations at the *ω*B97X-D/6-311+G(d,p)//*ω*B97X-D/6-31G(d,p) level. The 53 explored heteroleptic boranes $\mathbf{B}^{xy}$ (x = **1**–**6**, y = **a**–**w**) include two 2,6-F_2-3,5-R_2-C_6H groups, and their core structures are classified as $\mathbf{B}^{1}$–$\mathbf{B}^{6}$ depending on the R groups, i.e., $\mathbf{B}^{x}$(R) = $\mathbf{B}^{1}$(F), $\mathbf{B}^{2}$(Cl), $\mathbf{B}^{3}$(Br), $\mathbf{B}^{4}$(CF_3), and $\mathbf{B}^{5}$(H), whereas $\mathbf{B}^{6}$ includes 2,6-F_2-3-Cl-C_6H_2 groups. With these core structures, we combined 23 aryl groups (**a**–**w**), and completed the construction of the borane library with the addition of $B(C_6F_5)_3$. It should be noted that boranes $\mathbf{B}^{xa}$ (x = 1–3, 5–6), $\mathbf{B}^{1v}$, and $\mathbf{B}^{1w}$ were known before the construction of this library,[27,30,45–47] and their reactivity in the hydrogenation of unsaturated molecules inspired us when designing the library molecules. In fact, $\mathbf{B}^{1a}$ has previously been employed for the catalytic reductive alkylation of aniline derivatives to generate active FLP species with THF,[30] and thus, *in-silico* derivatization of $\mathbf{B}^{1a}$ was carried out via substitution of the *meta* and/or *para* H atoms in the 2,6-Cl_2-C_6H_3 group with Cl, Br, CF_3, OMe, OCF_3 or C_6F_5 groups. We further conducted extensive *in-silico* derivatization of $\mathbf{B}^{2a}$ and $\mathbf{B}^{3a}$, given that these boranes exhibit far superior catalytic activity than $\mathbf{B}^{1a}$, $\mathbf{B}^{6a}$, and $B(C_6F_5)_3$ in the hydrogenation of *N*-heteroaromatics using a gaseous mixture of H_2/CO/CO_2/CH_4.[45] In these cases, we envisioned that the modulation of the intrinsic Lewis acidity of the triarylboranes, i.e., the energy levels of the LUMO, which includes the p orbital on the boron center,[48] as well as the remote back-strain[49] that influences the stability of the four-coordinated tetrahedral Lewis base–borane adducts. The introduction of 2,6-Br_2-C_6H_3 (**j**) and its derivatives (**k**–**r**) into the $\mathbf{B}^{2}$ core was also explored, as we expected an increase in the front strain that influences the accessibility of the Lewis bases to the boron centers.[37]

We subsequently obtained the theoretical parameters. We thought that the use of structural parameters obtained from the gas-phase optimization of $\mathbf{B}^{xy}$ would not play a critical role in predicting the reactivity of the triarylboranes under the chosen conditions, given that no substantial differences were observed among them (Table S3). Thus, we obtained the following energetic parameters: (i) the energy levels of the LUMOs [eV], which include the p orbitals on the boron atoms, (ii) the energy barriers ($\Delta G_H^{\ddagger}$ in kcal mol^{-1}) for the heterolytic cleavage of H_2 with the combination of $\mathbf{B}^{xy}$ and THF, and (iii) the relative Gibbs energy values ($\Delta G_w°$) [kcal mol^{-1}] for the formation of the H_2O–$\mathbf{B}^{xy}$ adducts with respect to [H_2O + $\mathbf{B}^{xy}$]. These theoretical values are shown in Figure 3A for some selected boranes. For parameter (ii), we have previously proposed that the heterolytic cleavage of H_2 by FLPs should be involved in the rate-determining event of the $\mathbf{B}^{1a}$-catalyzed reductive alkylation of amines.[30] For parameter (iii), H_2O could be a potential quencher of the triarylborane catalysts via the formation of adducts followed by proto-deboronation.[50] In this context, we envisioned that boranes $\mathbf{B}^{xy}$ that exhibit larger $\Delta G_w°$ and smaller $\Delta G_H^{\ddagger}$ values should show superior performance as Lewis acids for the generation of more active FLPs with ethereal components. It should also be mentioned here that we theoretically optimized a structure that included a sole imaginary frequency related to the H–H bond cleavage; however, to reduce the calculation costs, we performed an IRC calculation only for selected cases

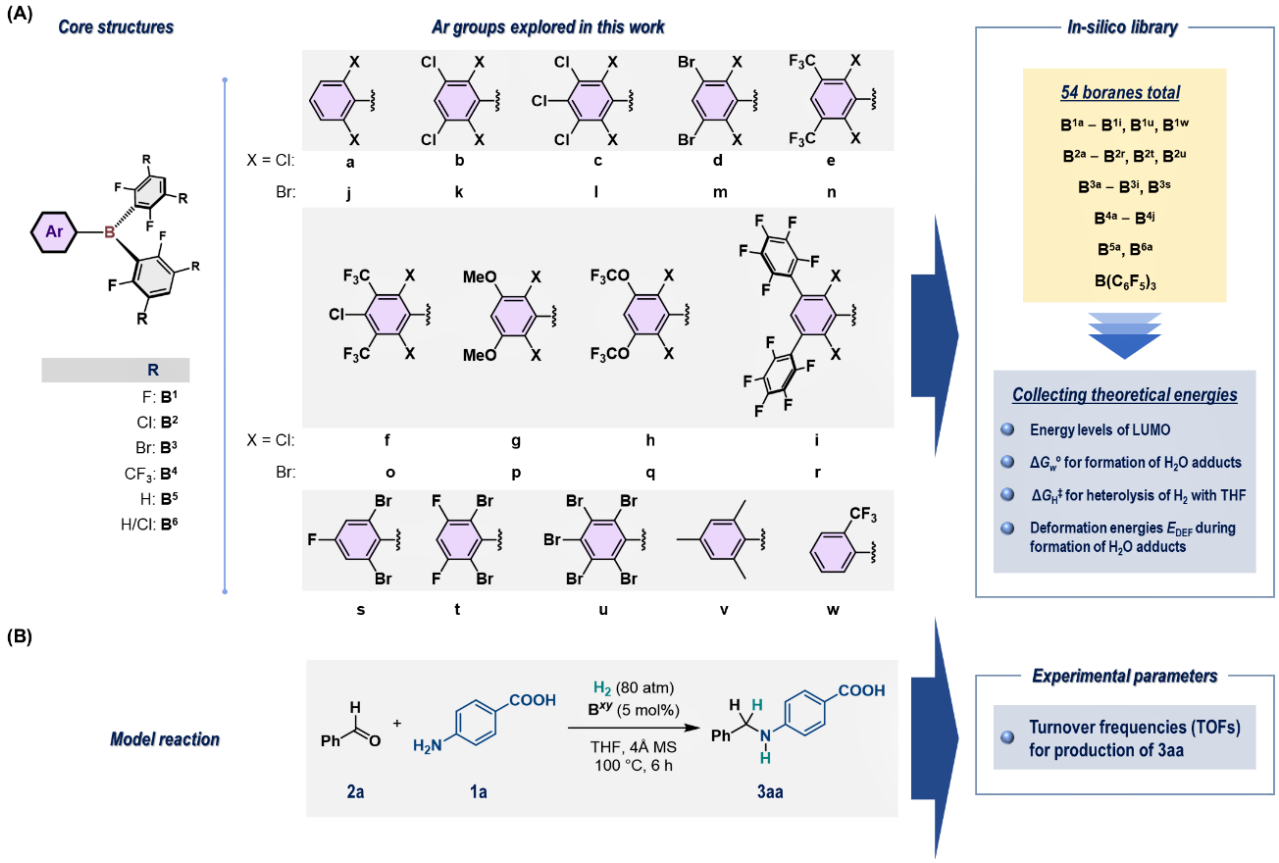

Figure 2. (A) The *in-silico* library of triarylboranes explored in this work. The structure of $\mathbf{B^6}$ includes 2,6-F_2-3-Cl-C_6H_2 groups. (B) The model reaction for obtaining the experimental parameters (turnover frequencies per hour) used in this work.

and confirmed their validity as a possible transition state structure.

Next, we turned our attention to collecting experimental data for the reported triarylboranes ($\mathbf{B^{xa}}$, x = 1–3, 5, 6; $\mathbf{B^{1v}}$; $\mathbf{B^{1w}}$; $B(C_6F_5)_3$) and newly synthesized boranes ($\mathbf{B^{1f}}$; $\mathbf{B^{2y}}$, y = b, c, e; $\mathbf{B^{3y}}$, y = b, c, s); the latter compounds were used to analyze the influence of the derivatization of the 2,6-Cl_2-C_6H_3 structure. We obtained the turnover frequency (TOF in h^{-1}; Figure 3A) as an experimental parameter calculated based on the yield of **3aa** from the reductive alkylation of amino acid **1a** with benzaldehyde (**2a**) and H_2 in the presence of 5 mol% $\mathbf{B^{xy}}$ under the shown conditions as a model reaction (Figure 2B). The design of this model reaction is based on the potential of $\mathbf{B^{1a}}$ to produce **3aa** in approximately 50% yield after 6 h (i.e., TOF = 1.67 h^{-1}), as this consideration allows for a clearer evaluation of the positive and negative effects of structural derivatization during the optimization of $\mathbf{B^{xy}}$.

With the theoretical and experimental parameters in hand, Gaussian process regression (GPR), which is a radial basis function kernel-based statistical-learning algorithm, using GPy (a programming library for GPR),[51] was applied to construct a model to predict the TOF values for the production of **3aa** (it should be noted here that unless stated otherwise, mean values are presented for the theoretically predicted TOF values).[52] GPR using GPy constructs a regression model using a limited number of observed data through ML and searches for a subsequent adequate parameter value of $\mathbf{B^{xy}}$ using the surrogate model. We evaluated the accuracy of each GPR model based on the coefficient of determination (Q^2) of the leave-one-out (LOO) cross-validation using either training or validation data. It is noteworthy that Q^2 values range between 0 and 1, whereby higher values indicate a more robust explanation for the data. We carried out the initial GPR analysis using the experimental TOF values and pairs of two of the three theoretical parameters (LUMO energy level, $\Delta G_w°$, and $\Delta G_H^‡$) obtained for the 15 boranes shown in Figure 3A as training data. This GPR analysis resulted in three distinct models, labelled **Model I** ($\Delta G_w°$ vs $\Delta G_H^‡$; Q^2 = 0.75), **II** (LUMO level vs $\Delta G_H^‡$; Q^2 = 0.76), and **III** (LUMO level vs $\Delta G_w°$; Q^2 = 0.18) (Figure 3B). Subsequently, we used the theoretical parameters of the other 39 triarylboranes to predict their TOF values using these three models. We found intriguing inconsistencies among the TOF values for the $\mathbf{B^4}$ derivatives that contain *meta*-CF_3 groups predicted using **Model I** and those predicted using **Model II** or **III** (Figure 3C). For example, when **Model I** was applied, the TOF values of $\mathbf{B^{4b}}$, $\mathbf{B^{4c}}$, and $\mathbf{B^{4e}}$ were predicted to be 3.50, 3.66, and 3.59, respectively; however, using **Model II** (or **III**), the corresponding TOF values were predicted to be 0.55 (2.10), 0.23 (2.03), and 0.55 (1.92) respectively (Figures 3C and 3D). A critical difference between **Model I** and **Model II** or **III** is the use of the LUMO energy levels in the GPR analysis. Thus, to evaluate whether the LUMO levels can serve here as a critical parameter for the prediction of the TOF for the production of **3aa**, we additionally synthesized $\mathbf{B^{4b}}$, $\mathbf{B^{4c}}$, and $\mathbf{B^{4e}}$, and confirmed that these boranes demonstrate excellent activity for the reductive alkylation of **1a** with **2a** under the model reaction conditions (Figure 3D). Subsequently, we evaluated the root mean square errors (RMSE) of **Model I**, **II**, and **III** using the experimental TOF values for $\mathbf{B^{4b}}$, $\mathbf{B^{4c}}$, and $\mathbf{B^{4e}}$ as test data; a model with a smaller RMSE value is considered more accurate. The accuracy of **Model I** (RMSE =

(A)

	B^{1a}	B^{1f}	B^{1v}	B^{1w}	B^{2a}	B^{2b}	B^{2c}	B^{2e}	B^{3a}	B^{3b}	B^{3c}	B^{3s}	B^{5a}	B^{6a}	$B(C_6F_5)_3$
LUMO	−1.33	−1.64	−0.92	−1.25	−1.31	−1.48	−1.53	−1.62	−1.28	−1.45	−1.50	−1.35	−1.03	−0.71	−1.93
$\Delta G_w°$	+5.2	+4.1	+7.7	+4.4	+5.6	+6.3	+6.5	+7.0	+5.5	+6.5	+6.6	+6.9	+6.1	+6.6	−0.9
$\Delta G_H^‡$	+25.5	+23.3	+30.1	+28.7	+24.4	+24.1	+24.5	+24.7	+24.5	+24.5	+24.5	+25.9	+25.4	+28.7	+21.3
TOF	1.87	1.80	0	0.90	2.70	3.20	2.87	2.67	2.70	2.93	2.87	1.63	1.97	0.10	1.33

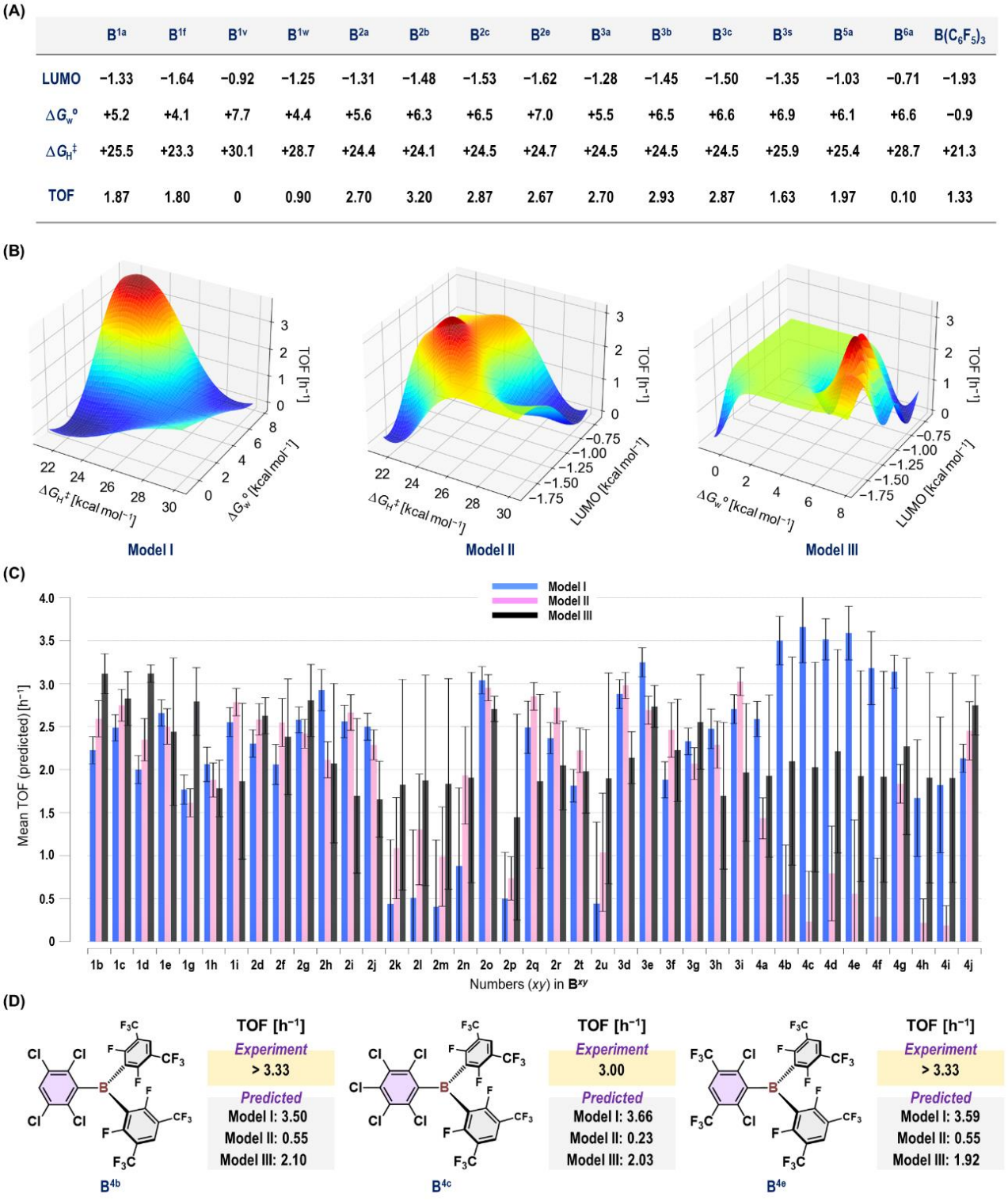

Figure 3. (A) LUMO energy level [eV], $\Delta G_w°$ [kcal mol^{-1}], and $\Delta G_H^‡$ [kcal mol^{-1}] for selected boranes $\mathbf{B}^{xy}$. TOFs [h^{-1}] calculated based on the yield of **3aa** under the model conditions are also shown. (B) Gaussian process regression using the programming library GPy for the prediction of the TOF values. The theoretical values of the parameters $\Delta G_w°$ and $\Delta G_H^‡$ (**Model I**), the LUMO level and $\Delta G_H^‡$ (**Model II**), or the LUMO level and $\Delta G_w°$ (**Model III**) were used. (C) Comparison of the TOF values predicted using **Models I**, **II**, or **III**. Error bars represent 1σ standard deviation. (D) Comparison of the experimental and predicted TOF values for $\mathbf{B}^{4b}$, $\mathbf{B}^{4c}$, and $\mathbf{B}^{4e}$.

0.42 [h^{-1}]) was confirmed to be greater than that of **Model II** (RMSE = 2.78 [h^{-1}]) and **Model III** (RMSE = 1.22 [h^{-1}]). These results indicate that **Model I** is the most adequate model for the prediction of the catalytic activity of $\mathbf{B}^{xy}$ under the applied conditions. Based on synthetic accessibility considerations, we finally decided to employ $\mathbf{B}^{4b}$ as the optimal catalyst in the following experiments. These results also suggest that the employment of a theoretical parameter related to the intrinsic Lewis

(A)

	B^{1a}	B^{1f}	B^{1v}	B^{1w}	B^{2a}	B^{2b}	B^{2c}	B^{2e}	B^{3a}	B^{3b}	B^{3c}	B^{3s}	B^{5a}	B^{6a}	$B(C_6F_5)_3$
E_{DEF}	+21.6	+25.4	+17.6	+16.2	+20.6	+22.1	+22.0	+23.6	+20.0	+21.7	+21.7	+21.4	+21.0	+20.3	+18.4

(B)

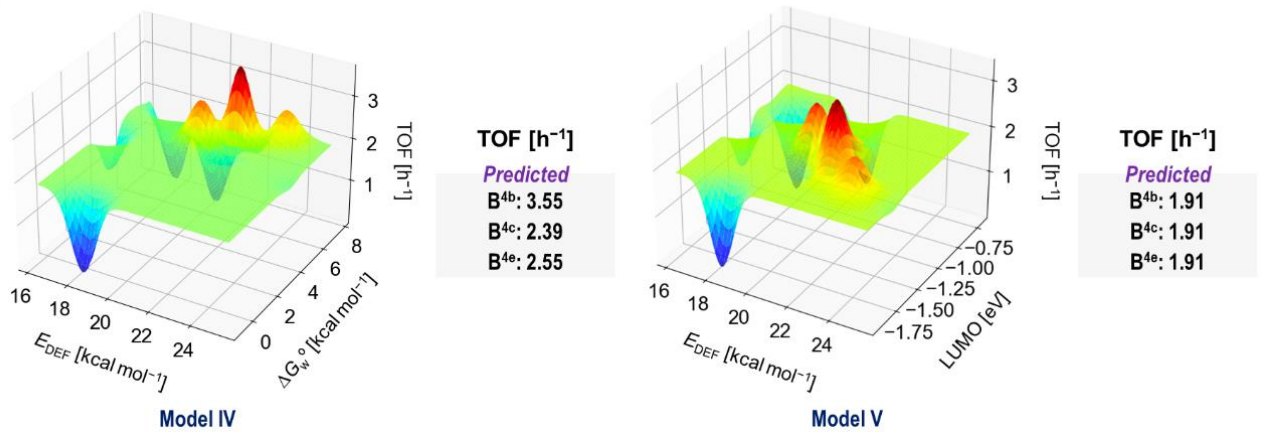

Figure 4. (A) Deformation energies (E_{DEF}) [kcal mol^{-1}] for selected boranes $\mathbf{B}^{xy}$. (B) Gaussian process regression with GPy for the prediction of TOF values [h^{-1}], using either E_{DEF} and $\Delta G_w°$ (**Model IV**) or E_{DEF} and the LUMO level (**Model V**).

acidity of triarylboranes, such as the LUMO energy level, is inappropriate for the prediction of the catalytic activity of triarylboranes under the chosen conditions.

Although we successfully developed **Model I** for the prediction of the catalyst activity of $\mathbf{B}^{xy}$, the calculation of $\Delta G_H^{‡}$ via the DFT-based optimization of a possible transition state is relatively inefficient in terms of costs. We thus pursued the use of a different theoretical parameter for the construction of a regression-based model for the prediction of the TOF values. This parameter should be related to the steric properties of the *meta*-substituents with respect to the boron centers in $\mathbf{B}^{xy}$, given that the GPR analysis using the intrinsic electronic properties of the boranes failed to construct reliable prediction models (*vide supra*). In this context, we have recently proposed that the deformation energy (E_{DEF}) [kcal mol^{-1}] can be used to evaluate the degree of remote back-strain,[49] where E_{DEF} represents the energetic penalty associated with the change in the conformation at the boron center from trigonal planar to tetrahedral upon the formation of LB–borane adducts.[48,53,54] These results prompted us to examine whether the E_{DEF} value could serve as an adequate index for designing boranes $\mathbf{B}^{xy}$ for the catalytic reductive alkylation of amino acids and peptides. We calculated the E_{DEF} values for the boranes shown in Figure 3A via the gas-phase optimization of their H_2O adducts followed by energy-decomposition analysis at the RI-DSD-PBEP86-D3BJ/ma-Def2-QZVPP//PBEh-3c/Def2-SVP level (Figure 4A). The GPR analysis with E_{DEF} and $\Delta G_w°$ using the experimental TOF values shown in Figure 3A as training data resulted in the construction of **Model IV** (Q^2 = 0.20), which exhibits nearly adequate predictions for the TOF values in the production of **3aa**, as exemplified by the cases with $\mathbf{B}^{4b}$ (3.55), $\mathbf{B}^{4c}$ (2.39), and $\mathbf{B}^{4e}$ (2.55) (Figure 4B, left). In contrast, **Model V** (Q^2 = 0.13), which was generated using the identical training data and the E_{DEF} and LUMO level, obviously underestimates the TOF values (1.91 in all cases), similar to the predictions of **Model III** (Figure 4B, right). Identical conclusions were confirmed when comparing the RMSE values of **Model IV** (0.59) and **V** (1.33) using the experimental TOF values for $\mathbf{B}^{4b}$, $\mathbf{B}^{4c}$, and $\mathbf{B}^{4e}$ as test data. These results indicate that the E_{DEF} can be a useful indicator to predict the activity of triarylboranes, at least under the present reaction conditions. Moreover, **Model V** again demonstrates that the inclusion of LUMO levels leads to underestimation of the TOF values.

With the optimal triarylborane $\mathbf{B}^{4b}$ in hand, we modified the reaction conditions to reduce its environmental impact and thus establish a greener and more sustainable system. In this context, the use of an alternative reaction solvent that could also act as a Lewis base to generate an FLP with $\mathbf{B}^{4b}$ was initially explored, given the recent demand for the replacement of hazardous THF with alternative ethereal compounds that exhibit lower toxicity combined with high chemical and thermal stability.[32] In the presence of 40 atm H_2, the reductive alkylation of **1a** with **2a** was carried out using THF, 2-methyltetrahydrofurane (2-MeTHF), cyclopentyl methyl ether (CPME), 4-methyltetrahydropyrane (MTHP), or 2,2,4-trimethyl-1,3-dioxolane (TMD) (Figure 5A). While THF provided a superior result (50%) compared to 2-MeTHF (38%) and CPME (18%), **3aa** was generated in 59% yield when MTHP was used. Prolongation of the reaction time to 24 h resulted in the formation of **3aa** in 72% yield; however, the removal of the 4Å MS caused a significant decrease in the yield of **3aa** to 42%. Finally, increasing the H_2 pressure to 60 atm resulted in the formation of **3aa** in 95% yield. The use of TMD did not furnish any **3aa**. It should be noted that MTHP can be easily separated from water (its solubility in H_2O is ~1.5 wt%) and removed under reduced pressure due to its strong hydrophobicity and low heat of vaporization, although its employment as a greener solvent has been limited in organic synthesis compared with the use of 2-MeTHF and CPME.[32,33] Moreover, to clarify the benefit of using MTHP over THF, we compared the activation energies for the heterolytic cleavage of H_2 by the combination of $\mathbf{B}^{4b}$ and THF or MTHP at the ωB97X-D/6-311+G(d,p)//ωB97X-D/6-31G(d,p) level (Figure 5B). A possible transition state was found in both cases, and that in the case of MTHP was found to be more stabilized (TS_{MTHP} = +21.8

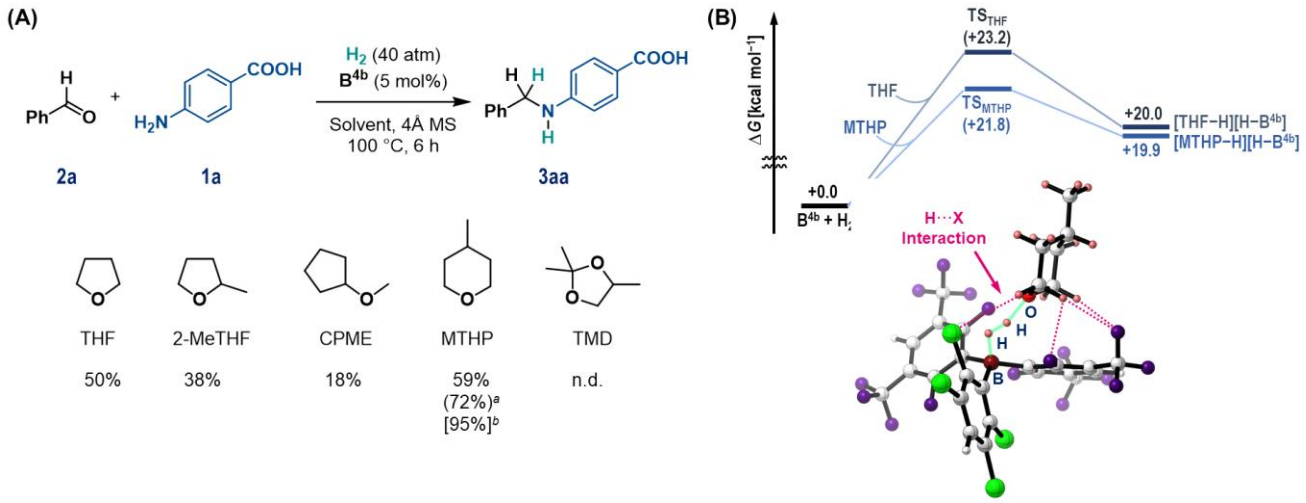

Figure 5. (A) Exploration of greener Lewis-basic solvents. Reaction conditions: **1a** (0.4 mmol, 0.05 M), **2a** (1.0 equiv.), $\mathbf{B^{4b}}$ (5 mol%), and 4Å MS (100 mg) were mixed in the solvent, followed by pressurization with H_2 (40 atm). The yield of **3aa** was determined via ^{1}H NMR analysis. [a]24 h. [b]60 atm H_2. (B) Relative Gibbs free energies [kcal mol^{-1}] with respect to [$\mathbf{B^{4b}}$ + H_2 + LB], where LB is either MTHP or THF, calculated at the ωB97X-D/6-311+G(d,p)//ωB97X-D/6-31G(d,p) level. The structure of TS_{MTHP} is also shown. Pairs of atoms involved in H···X (F/Cl) interactions that were found using AIM analysis are indicated by dashed lines (H: pink; B: brown; C: gray; O: red; F: purple; Cl: light green).

kcal mol^{-1}) than that in the case involving THF (TS_{THF} = +23.2 kcal mol^{-1}). We attribute this stabilization to the increased structural flexibility of the tetrahydropyrane motif relative to THF, which allows the formation of efficient non-covalent interactions (NCIs) between the F/Cl atoms in $\mathbf{B^{4b}}$ and the H atoms in MTHP. The participation of such NCIs was confirmed using the quantum theory of atoms in molecules (AIM) method (for details, see Figure S13).[55,56]

The $\mathbf{B^{4b}}$-catalyzed reductive alkylation of **1a** with **2a** in MTHP using H_2 (40 atm) demonstrated remarkable compatibility toward a variety of additives (**A0**–**A21**) (Figure 6). All these experiments were carried out twice, and mean values [%] are given for the yield of **3aa**, the recovered additive, and the imine intermediate 4-(benzylideneamino)benzoic acid formed *in situ* through the $\mathbf{B^{4b}}$-catalyzed condensation of **1a** and **2a**. Initially, we carried out a control experiment using **A0**, which confirmed that the yields of **3aa** and the remaining imine were consistent (72% and 28%, respectively) with those of the reaction conducted without **A0** (Figures 5A and 6). Relative to the control experiment, the reductive alkylation among **1a**, **2a**, and H_2 proceeded without a significant change in the yield of **3aa**, the recovered additives, or imine intermediates for additives with ketone (**A3**/**A13**), primary amide (**A4**), aryl bromide/iodide including alkyl ether (**A6**/**A7**), allylic ether (**A8**), terminal alkyne (**A9**), enone (**A12**), nitrile (**A15**), and ester (**A20**) moieties. It is also noteworthy that additives including sulfhydryl (**A16**) and sulfide (**A17**) moieties did not affect the present reaction, whereas such sulfur-containing compounds can be critical inhibitors in transition-metal-based catalysis and organocatalysis.[42] On the other hand, the hydrogenation of the imine intermediates was obviously suppressed in the presence of aliphatic/aromatic carboxy (**A1**/**A14**), aliphatic hydroxyl (**A2**), bulky silyl ether (**A10**), pinacolatoboryl (**A18**), and *N*-*tert*-butoxycarbonyl (Boc) (**A21**) moieties, as these functional groups include either a Lewis-basic or -acidic site that can kinetically inhibit the formation of the FLP consisting of $\mathbf{B^{4b}}$ and MTHP. In fact, the quantitative recovery of the additives after a period of 24 h was confirmed, without the generation of any other significant byproduct, i.e., the sum of the yields of **3aa** and the imine was always ~95%). In contrast, the formation of **3aa** was largely suppressed under reaction conditions including *N*-heteroaromatic moieties such as an imidazole (**A5**) and an indole (**A19**), as these heteroaromatic units can react with $\mathbf{B^{4b}}$ to form classical Lewis adducts and/or aldehydes to complicate the system. In the case of **A11,** which includes an epoxide moiety, **3aa** was only produced in 13% yield, and a significant loss of **A11** was confirmed after the reaction. Given that Lewis-acidic triarylboranes mediate the ring-opening transformation of epoxides,[57,58] the $\mathbf{B^{4b}}$-catalyzed ring-opening reaction of **A11** to give the corresponding aldehyde can be expected to compete with the targeted reaction (see Supporting Information for details).

Finally, we applied the combination of $\mathbf{B^{4b}}$ and MTHP for the reductive alkylation of amino acids and peptide derivatives in the presence of H_2 (Figure 7). Aminosalicylic acids **1b** and **1c** were effectively alkylated under the optimized conditions, and **3ba** and **3ca** were obtained in 90% and 93% yield, respectively; 60 atm of H_2 was used in the latter case. For comparison, under a pressure of 80 atm of H_2 in THF, **3ba** and **3ca** were furnished in 47% and 70% yield in the presence of 10 mol% and 15 mol% $\mathbf{B^{1a}}$, respectively, which again demonstrates the advantages of the present system using $\mathbf{B^{4b}}$ and MTHP in terms of synthetic efficiency and sustainability. The reductive alkylation of anthranilic acid (**1d**), which is also known as vitamin L1, 5-aminoisophthalic acid (**1e**), and 4-aminophenylacetic acid (**1f**) afforded **3da**, **3ea**, and **3fa** in excellent yield using 40–60 atm of H_2. In contrast, we recognized that aliphatic amino acids (or their imine derivatives) and substrates insoluble in MTHP were not suitable. For example, aspartic acid (**1g**) is insoluble in MTHP, and no reaction took place when **1g** was employed under otherwise identical conditions. Esterification of the carboxy group in 3-amino-4,4-dimethylpentanoic acid effectively improved its solubility in MTHP. Moreover, the hydrogenation of the imine derived from **1h** and **2a** did not occur. Based on these results, we prepared alanine- and *S*-methylcysteine-based

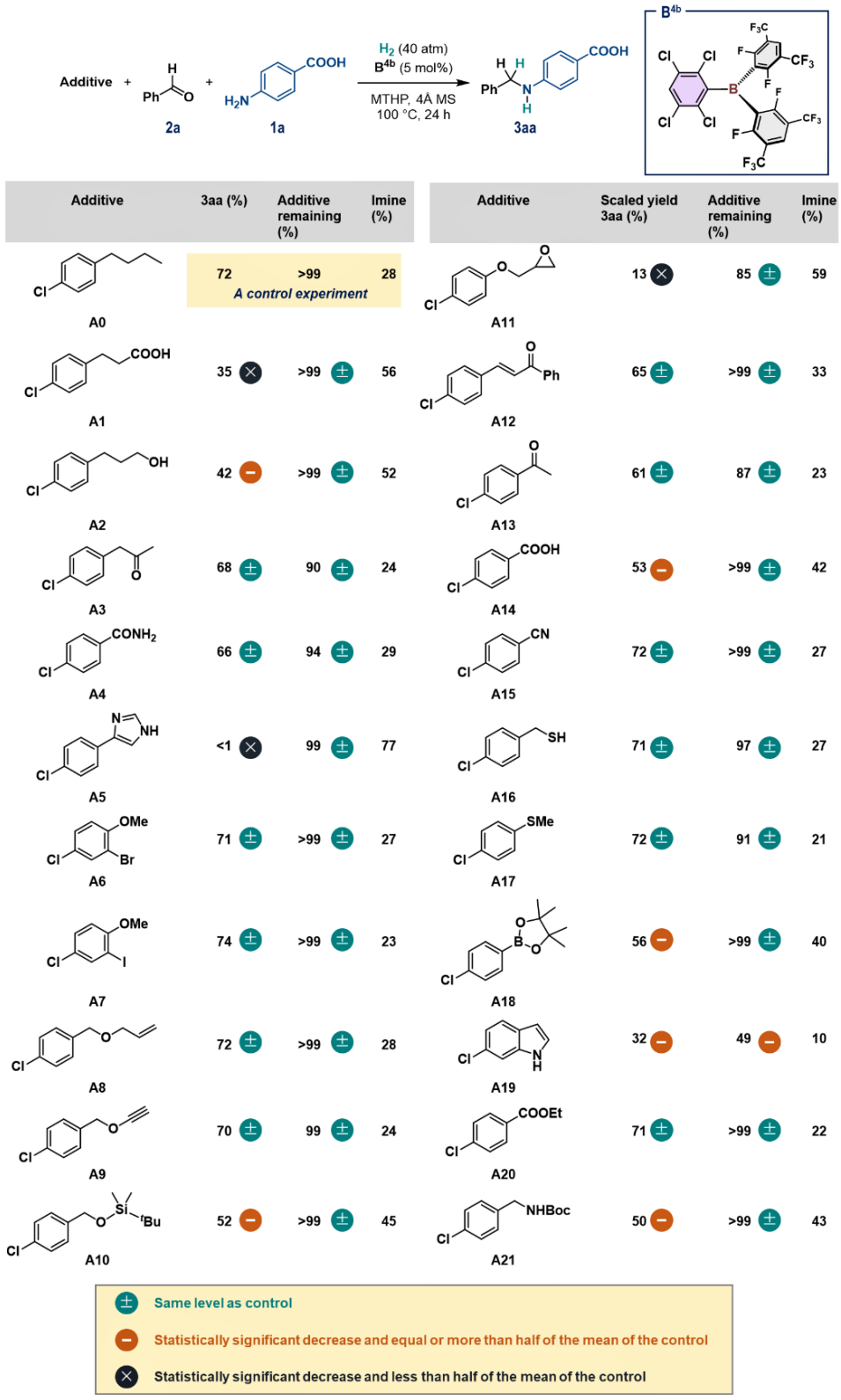

Figure 6. Exploring functional-group compatibility. Reaction conditions: **1a** (0.4 mmol, 0.05 M), **2a** (1.0 equiv.), additive (1.0 equiv.), **B**^{4b} (5 mol%) and 4Å MS (100 mg) were mixed in MTHP, followed by pressurization with H_2 (40 atm). The mean values of two experiments are given, which were determined via ^{1}H NMR analysis.

peptides **1i** and **1j**, and subjected them to the optimal reaction conditions; alkylated peptides **3ia** and **3ja** were obtained in 95% and 55% yield, respectively. In terms of the scope of aldehydes, 3,5-bis(trifluoromethyl)benzaldehyde (**2b**) furnished **3ab** in 93% yield, but 3,5-di-*tert*-butylbenzaldehyde (**2c**) gave **3ca** in merely 38%. In the latter case, a significant amount of **2c** remained unreacted, indicating difficulties associated with the formation of the imine intermediate due to the decreased

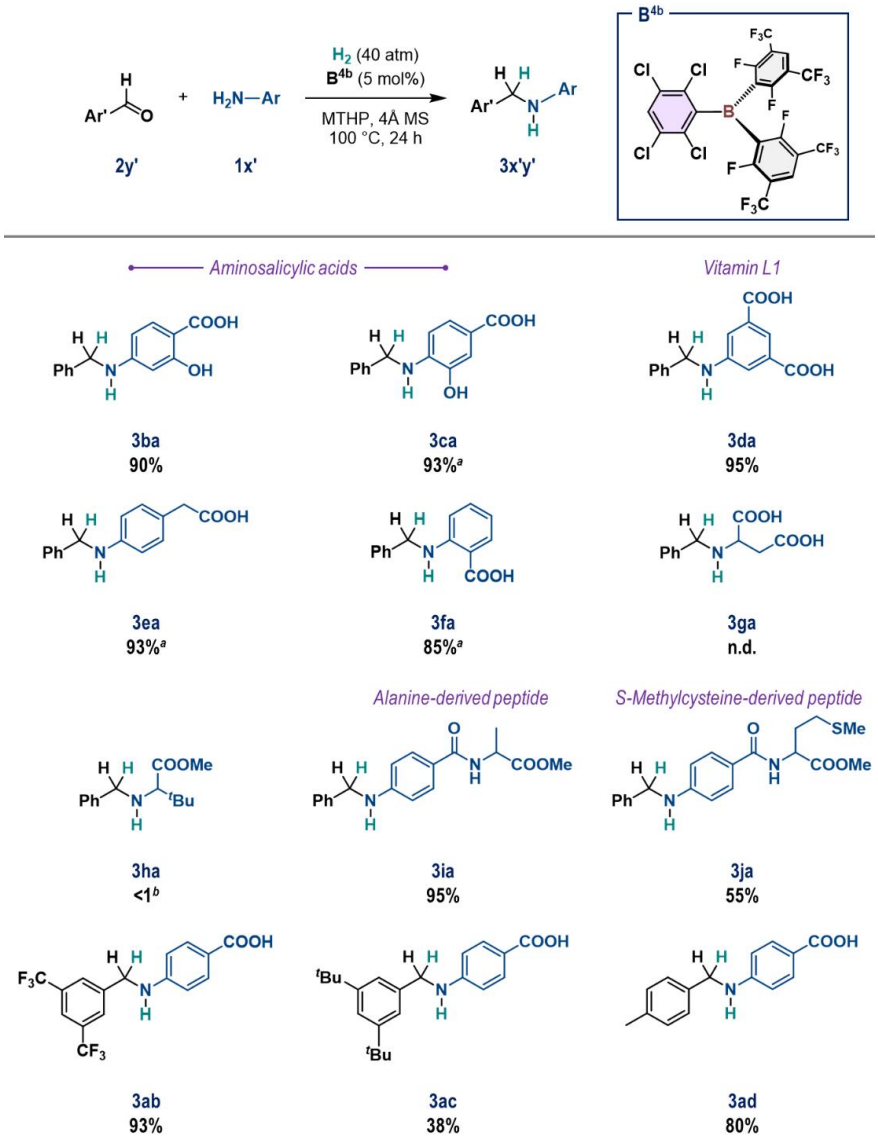

Figure 7. The $\mathbf{B^{4b}}$-catalyzed reductive alkylation of amino acids (**1a**–**1h**) and peptides (**1i** and **1j**) with aldehydes (**2a**–**d**) using H_2 in MTHP. General conditions: **1** (0.4 mmol, 0.05 M), **2** (1.0 equiv.), $\mathbf{B^{4b}}$ (5 mol%), and 4Å MS (100 mg) were mixed in MTHP, followed by pressurization with H_2 (40 atm). Yields of isolated products are given. [a]60 atm H_2. [b]The formation of the imine in >99% was confirmed.

electrophilicity of the aldehyde moiety. A comparable result was obtained when *p*-tolualdehyde (**2d**) was used with respect to the case using **2a**, and **3da** was afforded in 80% yield.

Conclusion

The present study demonstrates an *in-silico*-assisted approach to design triarylboranes that exhibit promising reactivity as main-group catalysts for the reductive alkylation of amino acids with H_2. We constructed an *in-silico* library of triarylboranes, including 46 unprecedented boranes, and obtained their theoretical parameters using DFT calculations. We initially synthesized seven of these unprecedented boranes to confirm their reactivity under the chosen model reaction conditions. Guided by Gaussian process regression using these theoretical and experimental parameters, we subsequently synthesized three more boranes and identified the optimal triarylborane, i.e., $B(2,3,5,6\text{-}Cl_4\text{-}C_6H)(2,6\text{-}F_2\text{-}3,5\text{-}(CF_3)_2\text{-}C_6H)_2$ ($\mathbf{B^{4b}}$). Through the evaluation of the regression-based models, we confirmed that the deformation energy (E_{DEF}) may serve as a potentially useful parameter to construct an adequate model, while the use of LUMO energy levels tends to lead to an underestimation when predicting the catalyst activity (TOF in h^{-1}) under the model reaction conditions. We also identified that 4-methyltetrahydropyrane (MTHP) is a superior Lewis-basic solvent for not only the generation of FLP species with $\mathbf{B^{4b}}$, but also for the realization of a more practical and sustainable reaction system compared to a system using THF. In fact, the $\mathbf{B^{4b}}$-catalyzed reductive alkylation using aldehydes as an alkylating reagent and H_2 in MTHP proceeded efficiently even in the presence of a variety of additives, showcasing its broad functional-group compatibility. Aniline-derived amino acids and C-terminal-protected peptides were alkylated in good-to-excellent yields under the optimized conditions with the concomitant generation of H_2O as the sole byproduct, demonstrating the significant advancement in terms of practicality and sustainability for the main-group-catalyzed reductive functionalization of valuable amines using H_2.

ASSOCIATED CONTENT

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Notes

The authors declare no competing financial interest.

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