

Pt(II)-Coordinated Tricomponent Self-Assemblies of Tetrapyridyl Porphyrin and Dicarboxylate Ligands: Are They 3D Prisms or 2D Bow-Ties?[†]

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Abstract

Thermodynamically favored heteroleptic coordination of Pt(II) ions with one aza- and another oxo-coordinating ligand yield tricomponent supramolecular coordination complexes (SCCs), which possess much greater structural complexity and functional diversity than traditional bicomponent SCCs containing only one of the ligands. Through X-ray crystallography, ¹H, ³¹P, and 2D NMR spectroscopies, mass spectrometry, and computational studies, herein, we demonstrated that heteroleptic coordination of tetrapyridyl porphyrins (M'TPP, M' = Zn or H₂) and various dicarboxylate ligands (XDC) having different lengths and rigidity with *cis*-(Et₃P)₂Pt^{II} corners yielded bow-tie (⌘)-shaped tricomponent [*cis*-(Et₃P)₂Pt]₄(M'TPP)(XDC)₂⁴⁺ complexes featuring a M'TPP core and two parallel XDC linkers held together by four heteroligated Pt^{II}(N,O) corners and ruled out the M'TPP-based tetragonal prism formation. Irrespective of the rigidity and length of the XDC linkers within a certain range (~7–11 Å), they intramolecularly bridged two adjacent pyridyl tips of an M'TPP ligand via Pt^{II}(N,O) corners, which led to the formation of bow-tie complexes instead of prisms. This happened because the angles of projection between the adjacent pyridyl rings of M'TPP cores adapted to accommodate the bridging XDC linkers having different lengths, and the bow-tie formation was entropically favored over tetragonal prisms. This work not only unveiled novel bow-tie-shaped coordination complexes, but also accurately defined the actual structures and compositions of M'TPP-based tricomponent SCCs. Furthermore, a representative bow-tie complex containing an electron-rich ZnTPP core selectively formed a charge-transfer (CT) complex with highly electron deficient 1,4,5,8,9,12-hexaazatriphenylene-2,3,6,7,10,11-hexacarbonitrile (HATHCN) but not with π -donors like pyrene.

Introduction

Owing to the dynamic, directional, and self-selecting/rectifying nature of metal–ligand coordination bonds, metal-driven self-assembly processes have emerged as one of the most attractive and powerful tools of supramolecular chemistry, yielding myriads of supramolecular coordination complexes (SCCs) ranging from discrete metallacycles^{1–10} and cages^{11–23} to extended coordination polymers and metal–organic frameworks^{24–26} over several decades. To obtain the target SCCs and to avoid statistical mixtures of different possibilities, only one rigid organic ligand is combined with a metal ion at appropriate stoichiometry, which usually yield bicomponent coordination complexes. However, the resulting bicomponent SCCs often lack the structural and functional diversity and tunability needed for various advanced applications. Expanding the scope of coordination-driven self-assembly strategies, recently researchers have discovered^{26–56} that *cis*-capped Pt(II) and Pd(II) corners simultaneously bind a carboxylate and a pyridyl ligands, which preferentially yield thermodynamically favored heteroleptic Pt(N,O) complexes instead of two different homoleptic complexes.^{29–34,47–53} Furthermore, when two different homoleptic Pt^{II}(COO[−])₂ and Pt^{II}(pyridyl)₂ complexes were mixed together at an appropriate stoichiometry, they spontaneously reorganized into thermodynamically more stable heteroleptic Pt(N,O) complexes.^{30,32,33} These revelations paved the door for metal-driven self-assembly of tricomponent metallacycles and cages containing two complementary ligands that could further diversify their structures, compositions, properties, and functions.

While it is fairly straightforward to assemble 2:2:4 tricomponent rectangles^{30,32,41} having two parallel dicarboxylate and two parallel dipyridyl arms connected by four heteroligated Pt^{II}(N,O) corners, the formation of 2:4:8 tricomponent tetragonal prisms featuring two cofacial tetratopic (tetrapyridyl or tetracarboxylate) ligands and four complementary ditopic (dicarboxylate or dipyridyl) linkers requires the latter to *intermolecularly* connect the tips of two separate cofacial tetratopic ligands via eight heteroleptic Pt(N,O) corners^{30,32,35,36} instead of *intramolecularly* bridging two adjacent tips of the same tetratopic ligand. If a ditopic linker *intramolecularly* bridges two adjacent binding sites of the tetratopic ligand via shared Pt(N,O) corners, then entropically more favored 1:2:4 bow-tie (⌘) complexes featuring one tetratopic core and two ditopic linkers connected by four heteroleptic corners would be formed instead of 2:4:8 tetragonal prisms containing two cofacial tetratopic faces connected by four ditopic linkers via eight heteroleptic corners.⁵⁴ Nevertheless, previous reports have claimed^{30,33,34} that Pt(II)-driven tricomponent self-assembly of tetrapyridyl porphyrin (M'TPP, M' = Zn- or H₂) and various aromatic and aliphatic dicarboxylate (XDC) linkers having varied length and rigidity yielded tetragonal prisms [*cis*-(Et₃P)₂Pt]₈(M'TPP)₂(XDC)₄]⁸⁺ featuring two parallel M'TPP faces and four XDC pillars connected by eight heteroligated Pt^{II}(N,O) corners.

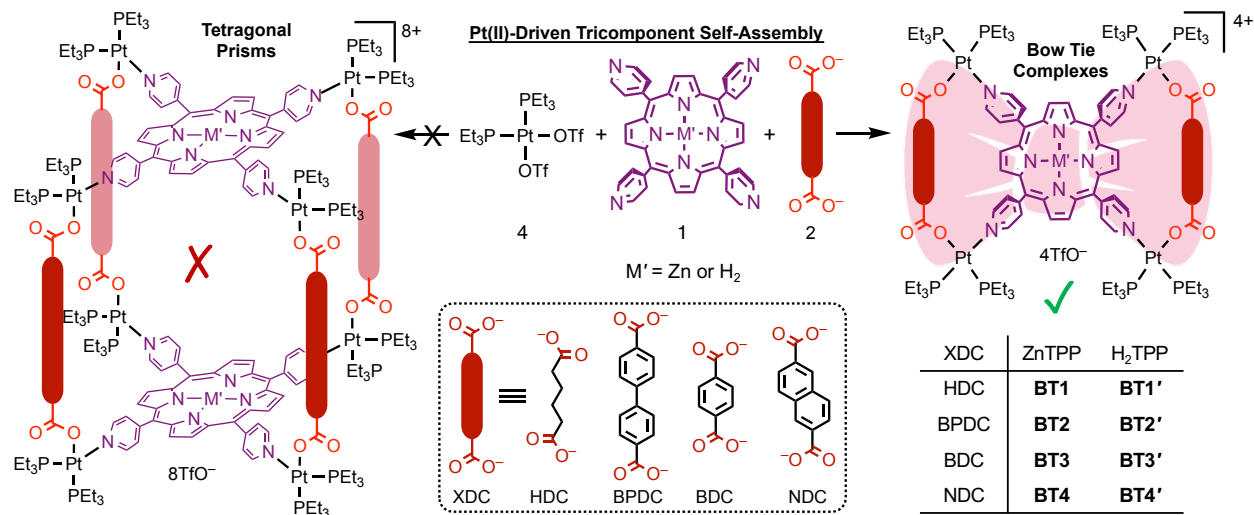
The resulting complexes preserved the photophysical properties of M'TPP chromophores³³ and displayed promising applications in cancer photodynamic therapy³⁹ and guest encapsulation.⁵⁵ Encouraged by these literature reports,^{30,33,34,39,56} we attempted to construct bi-chromophoric tetragonal prisms consisting of two M'TPP faces and four dicarboxylate linkers having complementary redox- and optically active aromatic cores, such as naphthalene- and perylene diimides that could support ligand-to-ligand photoinduced electron and/or energy transfer events. Surprisingly, none of our attempts to construct tetragonal prisms having two parallel M'TPP faces and four XDC linkers based on the reported protocols was successful despite the fact that the lengths of our XDC linkers were much longer than the distances between the two adjacent pyridyl-N atoms of free M'TPP ligands ($d_{N-N/free} = 10.9 \text{ \AA}$), which precluded the possibility of intramolecular bridging of two adjacent pyridyl groups via heteroligated $Pt^{II}(N,O)$ corners.

Prompted by these unexpected outcomes, we took a closer look at the reported 1H NMR spectra of the putative M'TPP-based tetragonal prisms^{30,33,34,56} and recognized that all of them have actually displayed two distinct singlets (1:1 ratio) for pyrrole protons—one set of four pyrrole protons were significantly more shielded than the other four (in contrast, all eight pyrrole protons of free M'TPP ligands are chemically equivalent and show one singlet). These NMR signatures revealed that the pyrrole rings of M'TPP ligands in these complexes were no longer chemically equivalent, i.e., they resided in two completely different environments. However, in the proposed tetragonal prisms, all pyrrole rings of M'TPP should have enjoyed the same chemical environment and displayed a characteristic singlet peak.^{47–49,53,55} Thus, the reported 1H NMR signals were not consistent with the proposed tetragonal prism structures and indicated the formation of 2D bow-tie (\bowtie) structures, a possibility that was previously overlooked. In bow-tie (\bowtie) complexes, the two opposite pyrrole rings of M'TPP would be located inside two isosceles triangles formed by two parallel XDC linkers and therefore shielded accordingly, while the other two opposite pyrrole rings would remain exposed and not shielded by the XDC linkers. These inconsistencies prompted us to carefully examine whether or not the $Pt(II)$ -driven self-assembly processes of M'TPP and XDC ligands indeed produce tricomponent prisms or yield an entirely different architecture, namely bow-ties having the same ratio (4:1:2) of the three components.

Herein, we report self-assembly and in-depth characterization of eight novel bow-tie complexes $[\{ cis-(Et_3P)_2Pt \}_4 (M'TPP)(XDC)_2] \cdot 4(TfO)$ (Scheme 1) composed of M'TPP ligands ($M' = Zn$ and H_2) and four XDC linkers, namely 1,6-hexane-, 4,4'-biphenyl-, 1,4-benzene-, and 2,6-naphthalene- dicarboxylates (HDC, BPDC, BDC, and NDC), having different lengths and rigidity. The most direct and compelling evidence of bow-tie architectures of the M'TPP-based tricomponent SCCs came from their first ever single-crystal X-ray structures, which were fully consistent with their solution phase 1H , ^{31}P , and 2D (COSY and

ROESY) NMR spectra as well as the ESI-MS data. The energy-minimized structures of these complexes were also in good agreement with the experimental results. These mutually corroborating results unequivocally demonstrated that each complex was composed of an M'TPP core and two parallel XDC linkers connected by four heteroligated $(\text{Et}_3\text{P})_2\text{Pt}^{\text{II}}(\text{N},\text{O})$ corners. Interestingly, the formation of M'TPP-based tricomponent bow-tie complexes instead of 3D prisms is also consistent with recent reports^{54,57–59} describing the formation of similar ‘triangular dicycles’ based on other tetratopic cores and ditopic linkers. Furthermore, we demonstrate the ability of an electron rich ZnTPP-based bow-tie complex to form π -donor/acceptor (D/A) charge-transfer (CT) complex with a highly electron deficient 1,4,5,8,9,12-hexaazatriphenylene-2,3,6,7,10,11-hexacarbonitrile (HATHCN) but not with another π -donor like pyrene.

Scheme 1. Pt(II)-driven self-assembly of M'TPP ligands ($\text{M}' = \text{Zn}$ or H_2) and four different XDC linkers (HDC, BPDC, BDC, and NDC) yielded novel bow-tie complexes $[\{\text{cis}-(\text{Et}_3\text{P})_2\text{Pt}\}_4(\text{M'TPP})(\text{XDC})_2] \cdot 4(\text{TfO})$ (BT1–BT4 and BT1'–BT4'). No tetragonal prism was formed irrespective of the length and rigidity of the XDC linkers.



Results and Discussions

Heteroleptic Coordination-Driven Self-Assembly of M'TPP-Based Tricomponent Bow-Ties (⌘). To determine whether heteroleptic coordination of *cis*-(Et_3P)₂Pt^{II} corners with tetratopic M'TPP ligands and ditopic XDC linkers yield tetragonal 3D prisms containing two cofacial M'TPP ligands and four XDC linkers held together by eight Pt(II) corners^{30,33,34} or 2D bow-ties containing a M'TPP core and two parallel XDC linkers connected by four shared Pt(II) corners, herein, we employed four XDC ligands—HDC, BPDC, BDC, and NDC—having different lengths, rigidity, and electron density. For consistency, we

applied the same self-assembly conditions reported in the literature,^{30,33,34} namely (i) 4:1:2 stoichiometry of *cis*-(Et₃P)₂Pt(TfO), M'TPP, and XDC; (ii) solvent mixtures (1:1:1 CH₂Cl₂/MeCN/MeNO₂ or 4:1 Me₂CO/H₂O) that adequately solubilized all components; (iii) temperature (~60 °C), and (iv) reaction time (~18 h) that were used to synthesize the proposed prisms. Our comprehensive 1D and 2D NMR, ESI-MS, single-crystal X-ray, and computational studies unequivocally demonstrated that regardless of their rigidity and lengths (~7–11 Å), all four XDC linkers exclusively afforded bow-tie complexes **BT1–BT4** and **BT1'–BT4'** (Scheme 1, Fig. 1) instead of 3D prisms. Notably, HDC and BDC linkers were also used previously for self-assembly of M'TPP-based tricomponent SCCs,^{30,33,34,39} and the reported NMR spectra of the resulting complexes were same as those displayed by our fully characterized bow-tie structures.

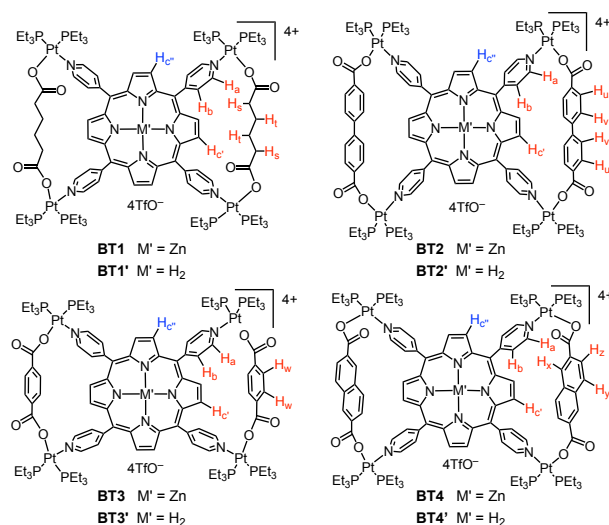


Fig. 1. The chemical structures of bow-tie complexes.

Single-Crystal Structures of Bow-Tie Complexes. The single-crystal X-ray analysis of these tricomponent SCCs presented the most direct and conclusive evidence of their bow-tie structures consisting of a M'TPP core and two parallel XDC linkers held together by four heteroleptic (Et₃P)₂Pt(N,O) corners (Fig. 2). Notably, this was the first time the crystal structures of any M'TPP-based tricomponent SCCs could be determined. The crystals were obtained from the respective NMR solutions (acetone-*d*₆) via either slow evaporation or vapor diffusion of CH₂Cl₂ or Et₂O, assuring that the solid-state crystal structures and solution NMR data belonged to the same materials. The bow-tie complexes based on rigid aromatic BPDC, BDC, and NDC linkers crystallized easily, but those containing flexible aliphatic HDC linkers did not. The structural features of all bow-tie complexes (Table 1) were fully consistent with their respective NMR and ESI-MS data, confirming that the same species were present both in solutions and solid crystals. The crystal structures of these bow-tie complexes shined light on why M'TPP core displayed two distinct ¹H NMR

signals for the enclosed H_c and exposed $H_{c'}$ protons (Fig. 1) and why only H_c protons were shielded by and ROE-coupled to XDC linkers but $H_{c'}$ protons were not.

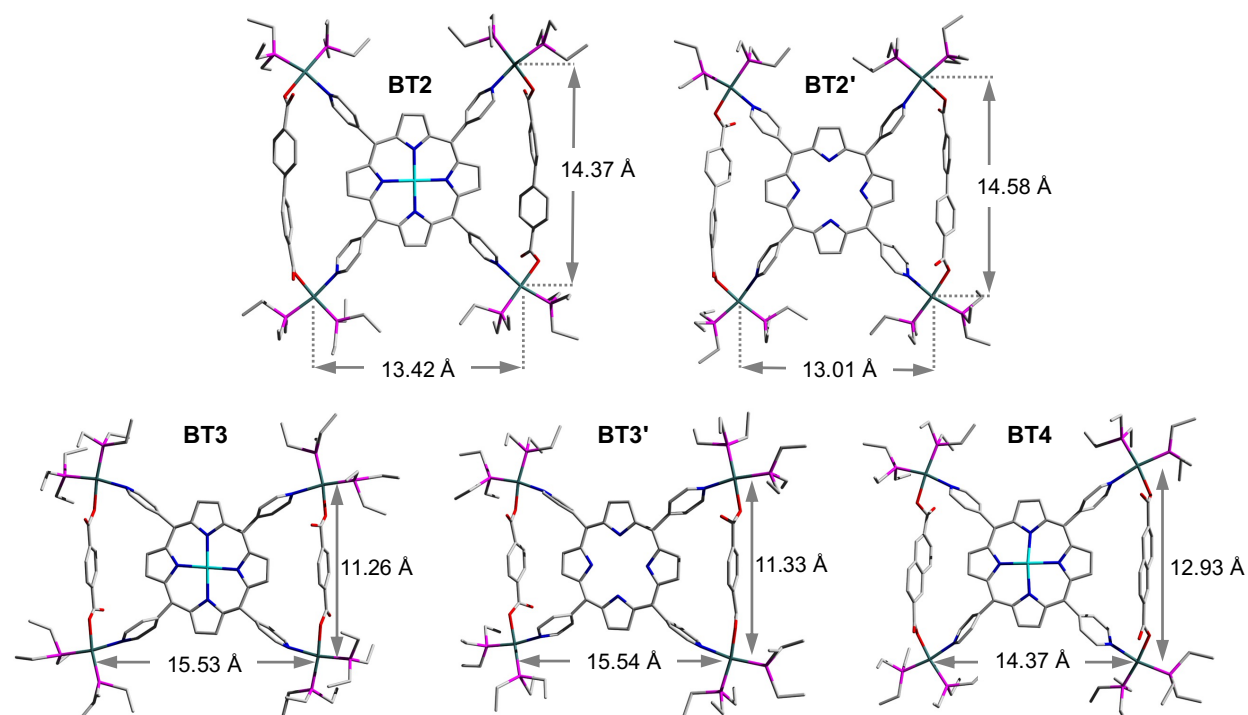


Fig. 2. Single-crystal structures of **BT2**, **BT2'**, **BT3**, **BT3'** and **BT4** bow-tie complexes. Atom legends: green: Pt, cyan: Zn, pink: P, red: O, blue: N, grey: C. The H-atoms and TfO^- anions were omitted for clarity.

Table 1. Key structural parameters of bow-tie complexes obtained from their single-crystal structures. The parameters obtained from the calculated structures are labeled with asterisks (*).

	BT1 calc.*	BT2	BT2'	BT3 (calc.*)	BT3' (calc.*)	BT4
Space Group	—	$P\bar{1}$	$C2/c$	$C2/c$	$I4_1/a$	$P2_1/c$
$d_{Pt-Pt/int}$ (Å)	11.55*	14.37	14.58	11.26 (11.33*)	11.33 (11.33*)	12.93
$d_{Pt-Pt/ext}$ (Å)	15.41*	13.42	13.01	15.53 (15.59*)	15.54 (15.56*)	14.37
$d_{N-N/int}$ (Å)	10.19*	11.38	11.52	10.09 (10.12*)	10.07 (10.20*)	10.83
$d_{N-N/ext}$ (Å)	11.63*	10.37	10.18	11.39 (11.69*)	11.61 (11.68*)	10.80
$d_{H_c-XDC-center}$ (Å)	2.84*	3.37	3.11	3.18 (3.50*)	3.05 (3.44*)	2.95
$d_{H_c-Pyridine-center}$ (Å)	2.96*	3.41	3.51	3.02 (2.98*)	2.96 (3.01*)	3.19
$d_{H_{c'}-Pyridine-center}$ (Å)	3.37*	3.08	3.00	3.41 (3.47*)	3.47 (3.41*)	3.26
$\angle N-Pt-O$ (°)	96*, 98*	82, 85	82, 83	78, 83 (93*)	80, 83 (93*)	82, 82
$\angle (N_{Py}-Center-N_{Py})_{int}$ (°)	82*	95	97	83 (81*)	82 (82*)	88
$\angle (N_{Py}-Center-N_{Py})_{ext}$ (°)	98*	85	83	97 (99*)	98 (98*)	89
$\theta_{Porphyrin/Pyridyl-dh}$ (°)	82*, 83*	76, 77	64, 73	86, 88 (84*)	76, 80 (85*)	65, 76
$\theta_{Pyrrole/XDC-dh}$ (°)	—	—	—	80 (90*)	85 (90*)	87

The bow-tie (\bowtie)-shaped **BT2** and **BT2'** complexes (Fig. 2: $[\{(\text{Et}_3\text{P})_2\text{Pt}\}_4(\text{M}'\text{TPP})(\text{BPDC})_2]^{4+}$, $\text{M}' = \text{Zn}$ or H_2) crystallized in $P\bar{1}$ and $C2/c$ space groups, respectively. The $\angle\text{N-Pt-O}$ angles of slightly distorted square-planar Pt(II) corners ranged between ca. $82\text{--}85^\circ$ (two diagonally opposite angles were the same). In **BT2** and **BT2'**, the distances between two BPDC-bridged Pt(II) corners ($d_{\text{Pt-Pt/int}}$) were ca. 14.5 and 14.6 Å, respectively, whereas those between two adjacent Pt(II) corners not bridged by BPDC ($d_{\text{Pt-Pt/ext}}$) were 13.4 and 13.0 Å, respectively. This happened because the long BPDC linker ($l_{\text{BPDC}} = 11.2$ Å)⁶⁰ positioned the two bridged Pt(II) corners farther away from each other, which in turn shortened $d_{\text{Pt-Pt/ext}}$. Consequently, the distances between the two adjacent pyridyl-N atoms coordinated to two BPDC-bridged Pt(II) corners, i.e., the pyridyl-N atoms that belonged to the same isosceles triangle, were also noticeably longer ($d_{\text{N-N/int}} \approx 11.4$ and Å) than the distance between two adjacent pyridyl-N atoms that were not part of the same triangle ($d_{\text{N-N/ext}} \approx 10.4$ Å). These N–N distances in **BT2** and **BT2'** deviated from the uniform distance between two adjacent pyridyl-N atoms of free M'TPP ligands ($d_{\text{N-N/free}} \approx 10.9$ Å).^{61,62} Furthermore, in **BT2**, the angles of projection between two adjacent pyridyl rings that belonged to the same triangle ($\angle(\text{N}_{\text{Py}}\text{--center--N}_{\text{Py}})_{\text{int}}$) expanded to 95° (97° in **BT2'**), while the angle between two adjacent pyridyl rings that were not part to the same triangle ($\angle(\text{N}_{\text{Py}}\text{--center--N}_{\text{Py}})_{\text{ext}}$) contracted to 85° (83° in **BT2'**). Thus, these angles deviated by $3\text{--}5^\circ$ from the angles of projection ($\sim 90^\circ$) between the adjacent pyridyl rings of free M'TPP in order to incorporate long BPDC linkers within the isosceles triangles of bow-tie structures. The dihedral angles between the porphyrin core and the pyridyl arms were 76° in **BT2** (73° in **BT2'**). The enclosed pyrrole rings of M'TPP and the twisted BPDC core (torsion angle between two phenyl rings $\approx 40^\circ$) located across the bay were not coplanar but formed large dihedral angles. The average distances from the enclosed H_c pyrrole protons located inside the triangles to the center of the closest phenyl ring of BPDC linker ($d_{\text{Hc-XDC}}$) was ca. 3.4 Å in **BT2** (3.1 Å in **BT2'**) and to the center of the nearest pyridyl ring ($d_{\text{Hc-Pyridine-center}}$) was 3.4 Å in **BT2** (3.5 Å in **BT2'**). Both distances were the longest among the bow-tie complexes presented here.

Although the entire single-crystal structures of **BT2** and **BT2'** complexes, including their ancillary Et_3P ligands on the Pt(II) corners, were well-resolved, only the basic bow-tie skeletons of **BT3**, **BT3'**, and **BT4** complexes consisting of the M'TPP core, two parallel BDC and NDC linkers, and four Pt(II) corners were fully resolved, but the fluxional CH_3CH_2 -groups and TfO^- anions were not, which caused large R -values. Nevertheless, since the ancillary ligands and counterions were not integral parts of the bow-tie structures, the poor resolution of these highly disordered components had little effect on key structural features that influenced their respective NMR spectra.

The BDC-based bow-tie complexes **BT3** and **BT3'** (Fig. 2: $[\{(\text{Et}_3\text{P})_2\text{Pt}\}_4(\text{M}'\text{TPP})(\text{BDC})_2]^{4+}$, $\text{M}' = \text{Zn}$ or H_2) crystallized in $C2/c$ and $I4_1/a$ space groups, respectively. The $\angle\text{N-Pt-O}$ angles at distorted square-planar Pt(II) corners ranged between ca. $78\text{--}83^\circ$. The distances between two adjacent Pt(II) corners bridged by short BDC linkers ($l_{\text{BDC}} = 6.9 \text{ \AA}$)⁶⁰ were ca. 11.3 \AA ($d_{\text{Pt-Pt/int}}$), whereas the distances between two adjacent Pt(II) corners not bridged by BDC ($d_{\text{Pt-Pt/ext}}$) were ca. 15.5 \AA . Thus, the short BDC linkers placed the two bridged Pt(II) corners closer to each other while increasing the distances between the two non-bridged Pt(II) corners. Consequently, the distances between the pyridyl-N atoms that belonged to the same triangle were also shorter ($d_{\text{N-N/int}} \approx 10.1 \text{ \AA}$) than those between two pyridyl-N atoms that were not part of the same triangle ($d_{\text{N-N/ext}} \approx 11.5 \text{ \AA}$). Consequently, the projection angles between two adjacent pyridyl rings belonging to the same triangle ($\angle(\text{N}_{\text{Py-center-N}_{\text{Py}}})_{\text{int}}$) shrunk to ca. 82° , while the angles between two adjacent pyridyl rings not belonging to the same triangle ($\angle(\text{N}_{\text{Py-center-N}_{\text{Py}}})_{\text{ext}}$) expanded to ca. 98° (ca. 8° deviations from the ideal $\sim 90^\circ$ angle in free $\text{M}'\text{TPP}$) in order to accommodate short BDC linkers within the triangles. The dihedral angles between the enclosed pyrrole rings of $\text{M}'\text{TPP}$ and the BDC core located across the bay were $80\text{--}85^\circ$, i.e., they were almost orthogonal to each other. The enclosed H_c pyrrole protons were projected toward the center of the BDC ring ($d_{\text{Hc'-XDC}} \approx 3.1 \text{ \AA}$) as well as the adjacent pyridyl rings, which were pulled closer ($d_{\text{Hc'-Pyridine-center}} \approx 3 \text{ \AA}$) by the BDC linker.

The NDC-based bow-tie complex **BT4** $[\{(\text{Et}_3\text{P})_2\text{Pt}\}_4(\text{ZnTPP})(\text{NDC})_2]^{4+}$ (Fig. 2) possessed $P2_1/c$ space group. All four $\angle\text{N-Pt-O}$ angles of distorted square-planar Pt(II) corners were ca. 82° . The distances between two adjacent NDC-bridged Pt(II) corners ($d_{\text{Pt-Pt/int}} = 12.9 \text{ \AA}$) were slightly shorter than those between two adjacent Pt(II) corners not bridged by a NDC linker ($d_{\text{Pt-Pt/ext}} = 14.4 \text{ \AA}$). However, the distances between the N-atoms of two adjacent pyridyl rings belonging to the same triangle ($d_{\text{N-N/int}}$) and those between two adjacent pyridyl-N atoms that were not part of the same triangle ($d_{\text{N-N/ext}}$) were almost same (ca. 10.8 \AA) and close to the uniform distances between two adjacent pyridyl-N atoms of free ZnTPP ligand ($d_{\text{N-N/free}} \sim 10.9 \text{ \AA}$). As a result, the angles between two adjacent pyridyl rings belonging to the same triangle ($\angle(\text{N}_{\text{Py-Center-N}_{\text{Py}}})_{\text{int}}$) and those between two adjacent pyridyl rings not belonging to the same triangle ($\angle(\text{N}_{\text{Py-Center-N}_{\text{Py}}})_{\text{ext}}$) in **BT4** were also close to ideal $\sim 90^\circ$ angle between adjacent pyridyl rings found in free ZnTPP ligand. Thus, the intermediate length of NDC linker ($l_{\text{NDC}} = 9.2 \text{ \AA}$)⁶⁰ caused the least distortion of ZnTPP ligand in order to form the isosceles triangles. In **BT4**, the ZnTPP and the NDC planes were nearly orthogonal to each other (dihedral angles = 87°). The enclosed H_c pyrrole protons located inside the triangles were projected toward the NDC core across the bay ($d_{\text{Hc'-XDC}} \approx 3 \text{ \AA}$), as well as to the center of the adjacent pyridyl rings ($d_{\text{Hc'-Pyridine-center}} \approx 3.2 \text{ \AA}$).

Energy Optimized Structures. Since the crystal structures of **BT1** or **BT1'** complexes containing flexible HDC linker were not available, we calculated their energy minimized structures using Gaussian 09 software (Fig. 3). To verify the accuracy of these calculated structures, we also calculated the energy minimized structures of BDC-based bow-tie complexes, which were in good agreement with their actual crystal structures (Table 1). In **BT1** and **BT1'** complexes, the $\angle\text{N-Pt-O}$ angles at the distorted square-planar Pt(II) corners were ca. $96\text{--}98^\circ$. The two adjacent HDC-bridged Pt(II) corners ($d_{\text{Pt-Pt/int}}$) were located ca. 11.6 \AA apart, whereas the two adjacent Pt(II) corners not bridged by HDC ($d_{\text{Pt-Pt/ext}}$) were located ca. 15.4 \AA apart. The bridged Pt-Pt distances in **BT1** were comparable to those found in **BT3**, suggesting that HDC and BDC have similar lengths. Like BDC, the short HDC linkers also brought the two bridged Pt(II) corners closer to each other, expanding the distances between the two adjacent non-bridged Pt(II) corners. Consequently, the angles between two adjacent pyridyl rings of M'TPP that belonged to the same triangle ($\angle(\text{N}_{\text{Py-center-N}_{\text{Py}}})_{\text{int}}$) shrunk to $\sim 82^\circ$, while the angles between two adjacent pyridyl rings that were not part of the same triangle ($\angle(\text{N}_{\text{Py-center-N}_{\text{Py}}})_{\text{ext}}$) expanded to $\sim 98^\circ$ in order to accommodate short HDC linkers within the isosceles triangles of these bow-tie structures. The distances from the enclosed H_c pyrrole protons to the center of nearest pyridyl ring ($d_{\text{Hc'-Pyridine-center}}$) were ca. 2.96 \AA and to the H_t protons of HDC linker located across the bay were ca. 2.8 \AA . The calculated structures of these complexes were consistent with their respective NMR spectra.

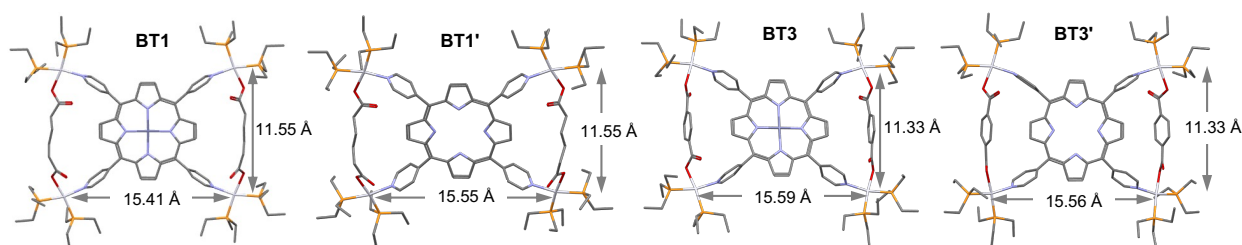


Fig. 3. The optimized structures of **BT1**, **BT1'**, **BT3**, and **BT3'** complexes calculated by PM6 method.

Thus, the single-crystal and energy-minimized structures of M'TPP-based tricomponent SCCs demonstrated that regardless of the length (within the range of $6.9\text{--}11.2\text{ \AA}$) and rigidity of the XDC linkers, the Pt(II)-driven self-assembly of M'TPP and XDC ligands yielded bow-tie complexes instead of 3D prisms. This happened because the pyridyl arms of M'TPP ligands deviated from their original projections in order to incorporate the XDC linkers within the isosceles triangles of bow-tie structures.

^{31}P NMR Spectroscopy. Simultaneous coordination of a pyridyl group of M'TPP and a carboxylate group of XDC with *cis*-(Et_3P) $_2\text{Pt}^{\text{II}}$ was also evident from the ^{31}P NMR spectra of the resulting SCCs (Fig. 4 and S1). While *cis*-(Et_3P) $_2\text{Pt}(\text{TfO})_2$ displayed a characteristic singlet at 11.14 ppm indicating that both Pt(II)-

coordinated P atoms were chemically equivalent, the resulting **BT1–BT4** and **BT1'–BT4'** complexes displayed two distinct doublets—one at ca. 0–2 ppm that corresponded to the P-nucleus *trans*- to the pyridyl-N atom and another at 6–7 ppm ($^2J_{\text{P-P}} \approx 20\text{--}21\text{ Hz}$) corresponding to the P-nucleus *trans*- to the carboxylate O-atom—indicating that the two ancillary Et₃P ligands were not chemically equivalent.^{30,33,34} The fact that no other ³¹P NMR signal was observed further confirmed exclusively heteroleptic coordination of two different ligands with the Pt(II) centers and ruled out the formation of any homoleptic complex. However, the ³¹P NMR spectra shed little light on the actual structures and compositions of the resulting tricomponent SCCs, i.e., whether they were 3D prisms or 2D bow-ties, as both structures would feature the same heteroleptic *cis*-(Et₃P)₂Pt^{II}(N,O) corners.

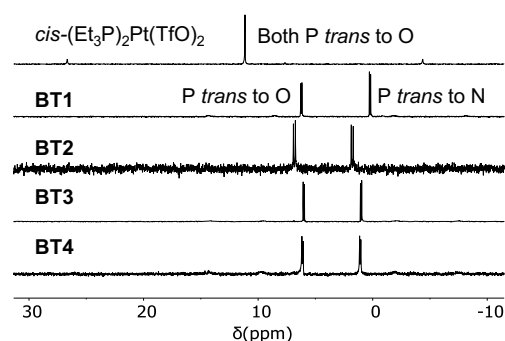


Fig. 4. Partial ³¹P NMR spectra (122 MHz, acetone-*d*₆) of *cis*-(Et₃P)₂Pt(TfO)₂ and **BT1–BT4** complexes.

¹H NMR Spectroscopy. The ¹H NMR spectra of the resulting complexes (Fig. 5, S2, and S3) also presented telltale signs of their bow-tie structures and ruled out tetragonal prism formation. In comparison to *D*_{4h}-symmetric free ZnTPP ligand, which displayed two doublets at 8.58 and 7.95 ppm corresponding to its H_a and H_b pyridyl protons (8 each), respectively, and a sharp singlet at 8.61 ppm for all eight chemically equivalent H_c pyrrole protons (Fig. 5), **BT1–BT4** complexes not only displayed significant downfield shift of H_a and H_b pyridyl protons due to Pt(II)-coordination, but most tellingly, two distinct singlets with 1:1 integral ratio for H_{c'} and H_{c''} (pyrrole) protons indicating that the pyrrole rings of ZnTPP were no longer chemically equivalent. The H₂TPP-based **BT1'–BT4'** complexes (Fig. S2), as well as all previously reported M'TPP-based tricomponent SCCs,^{30,33,34} also displayed two distinct singlets for the pyrrole protons, which indicated that the pyrrole rings of M'TPP resided in two different environments, a telltale sign of bow-tie complexes, not prisms.

The singlets peaks corresponding to four H_{c'} pyrrole protons of **BT1–BT4** complexes appeared at 8.38, 7.73, 7.13, and 6.79 ppm, respectively, which were significantly up-field shifted from the H_c signal (8.61 ppm) of free ZnTPP, while the singlets corresponding to four H_{c''} pyrrole protons appeared at 9.18,

8.52, 9.16, and 8.86 ppm, respectively. Thus, the up-field shifts of H_c signals were directly correlated to the increasing electron cloud, i.e., the shielding effect of XDC linkers ($HDC < BPDC < BDC < NDC$), suggesting that these pyrrole protons were located inside the isosceles triangles formed by two parallel XDC linkers, whereas the $H_{c''}$ pyrrole protons were not. The aliphatic HDC linkers in **BT1** exerted the least shielding effect, causing the smallest up-field shift of the enclosed H_c pyrrole protons located across the bay, followed by the BPDC linkers having two twisted phenyl rings that were not properly aligned with the enclosed H_c pyrrole protons of **BT2**, as seen from its single-crystal structure (*vide supra*). Having the largest π -cloud and the maximum shielding effect among four XDC linkers employed here, the NDC linkers in **BT4** caused the largest up-field shift of enclosed H_c protons, followed by the BDC linkers in **BT3**. Although the exposed $H_{c''}$ pyrrole protons were not shielded by distant XDC linkers, their chemical shifts were affected by the length of the XDC linkers, which controlled their distances from the adjacent pyridyl rings of ZnTPP. This was further evident from their crystal structures and certain structural features summarized in Table 1. For instance, the shorter HDC and BDC linkers ($l \approx 7 \text{ \AA}$) placed the bridged Pt(II) corners closer to each other, which in turn, pulled the pyridyl rings of ZnTPP ligand away from the exposed $H_{c''}$ pyrrole protons of **BT1** and **BT3** complexes. As a result, the shielding effect of pyridyl rings on $H_{c''}$ pyrrole protons in **BT1** and **BT3** was diminished, and they appeared at more downfield positions (ca. 9.2 ppm) than the H_c protons of free ZnTPP ligand. On the other hand, the longest BPDC linkers ($l_{BPDC} = 11.2 \text{ \AA}$) in **BT2** held the bridged Pt(II) corners farther away from each other, which in turn, pushed the pyridyl rings of ZnTPP closer to exposed $H_{c''}$ pyrrole protons making them more shielded and up-field shifted than the H_c protons of free ZnTPP. Having an intermediate length, NDC linkers ($l_{NDC} = 9.2 \text{ \AA}$) in **BT4** pulled the pyridyl rings of ZnTPP slightly away from exposed $H_{c''}$ protons, causing a slight downfield shift. The exact same trends were observed for H_2 TPP-based **BT1'**–**BT4'** complexes (Fig. S2).

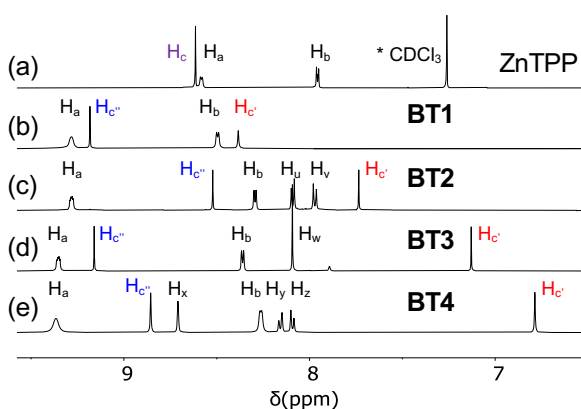


Fig. 5. Partial 1H NMR spectra (500 MHz) of (a) free ZnTPP ligand, (b) **BT1**, (c) **BT2**, (d) **BT3**, and (e) **BT4**. The enclosed H_c pyrrole protons (highlighted in red) located inside the isosceles triangles of bow-tie

structures were shielded proportionately by adjacent XDC linkers, whereas the exposed $H_{c''}$ pyrrole protons (highlighted in blue) were not.

Thus, the splitting of pyrrole protons of M'TPP ligands into two chemically nonequivalent and noncoupled $H_{c'}$ and $H_{c''}$ protons in the tricomponent SCCs and the variable shielding of the former by XDC linkers were the unmistakable signs of bow-tie complexes where two opposite pyrrole rings bearing four $H_{c'}$ protons were located inside the isosceles triangles formed by two parallel XDC linkers while the other two pyrrole rings bearing four $H_{c''}$ protons remained exposed. It is worth noting that previously reported M'TPP-based tricomponent SCCs, including those containing HDC and BDC ligands,^{32,33} essentially displayed the same 1H NMR characteristics as those displayed by our **BT1–BT4** and **BT1'–BT4'** complexes, i.e., two distinct singlets for $H_{c'}$ and $H_{c''}$ pyrrole protons, which are the characteristic signs of bow-tie structures. If the resulting tricomponent SCCs were indeed tetragonal prisms having two cofacial M'TPP planes connected by four XDC linkers via heteroleptic $(Et_3P)_2Pt^{II}(N,O)$ corners,^{30,33,34,56} then all sixteen pyrrole protons of M'TPP faces should have remained chemically equivalent and displayed one singlet peak instead of splitting into two chemically non-equivalent $H_{c'}$ and $H_{c''}$ protons with two distinct singlets, as observed in other porphyrin-based prisms.^{47–49,53,55} However, that was not observed for any $Pt(II)/M'TPP/XDC$ -based tricomponent SCCs,^{30,33,34,56} which ruled out the prism formation.

1H – 1H COSY NMR Spectroscopy. The COSY NMR spectra of these M'TTP-based tricomponent SCCs (Fig. S4) provided further insights into their actual structures by revealing the coupling between the adjacent (α - and β -) protons. Conspicuously missing from the COSY NMR spectra of all these SCCs were any α/β -coupling between the $H_{c'}$ and $H_{c''}$ protons of M'TPP ligand, which further indicated that these two chemically non-equivalent protons did not belong to the same pyrrole ring but to two separate pyrrole rings located in different environments. This scenario was possible only in bow-tie structures where two opposite pyrrole rings carrying the $H_{c'}$ protons were located inside the triangles, whereas the other two pyrrole rings bearing $H_{c''}$ protons remained exposed. If these tricomponent SCCs were indeed tetragonal prisms containing two cofacial M'TPP panels linked by four XDC linkers, then either all the pyrrole protons of M'TPP should have remained chemically equivalent instead of splitting into distinct $H_{c'}$ and $H_{c''}$ protons, or each pyrrole ring would have carried one $H_{c'}$ and one $H_{c''}$ protons involved in α/β -coupling (in case of twisted prisms).⁶³ The absence of such couplings ruled out the prism formation.

ROESY NMR Spectroscopy. Another powerful evidence of bow-tie complex formation was found in the ROESY NMR spectra (Fig. 6 and S5), which revealed long-range coupling between the enclosed $H_{c'}$ pyrrole protons of M'TPP and the XDC protons located across the bay. For example, the enclosed $H_{c'}$ pyrrole

protons were coupled with (i) the H_t protons of HDC in **BT1** and **BT1'**, (ii) the H_u and H_v protons of BPDC in **BT2** and **BT2'**, (iii) the H_w protons of BDC in **BT3** and **BT3'**, and (iv) the H_x , H_y , and H_z protons of NDC in **BT4** and **BT4'**, indicating that the enclosed pyrrole rings were located in close proximity of the XDC linkers. In addition, the enclosed $H_{c'}$ protons were also coupled to adjacent H_b pyridyl protons of M'TPP ligands. However, no such ROE-coupling between the exposed $H_{c''}$ pyrrole protons and the remote XDC protons were observed (the $H_{c''}$ protons were only coupled to H_b protons of the adjacent pyridyl ring in some cases), further verifying that the enclosed $H_{c'}$ and exposed $H_{c''}$ pyrrole protons were located in two different chemical environments. This scenario would be only possible in bow-tie structures, not in prisms.

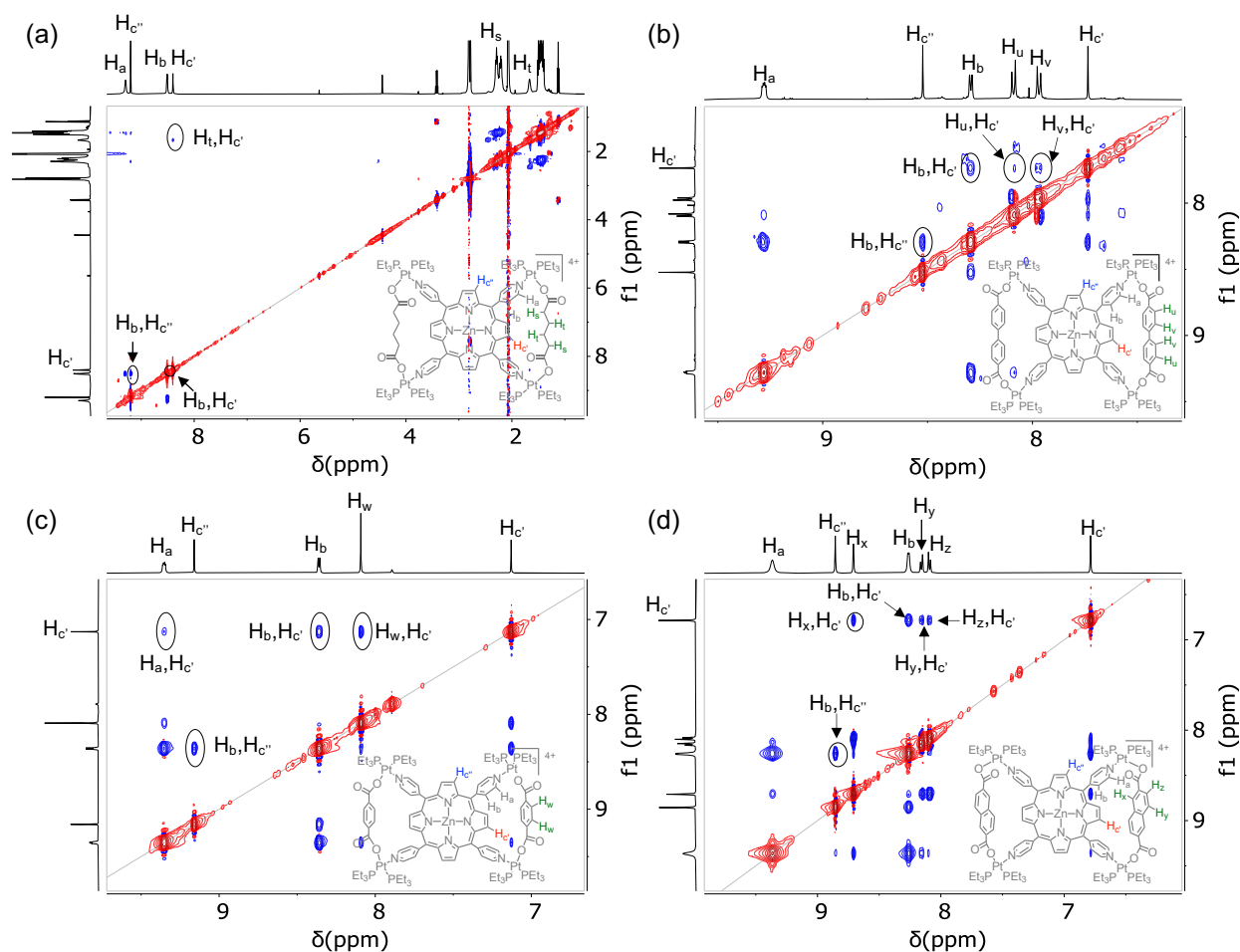


Fig. 6. Partial ^1H - ^1H ROESY NMR spectra (500 MHz, acetone- d_6) of (a) **BT1**, (b) **BT2**, (c) **BT3**, and (d) **BT4** show that the enclosed $H_{c'}$ pyrrole protons of these bow-tie complexes located inside the isosceles triangles are through-space coupled with the protons of adjacent XDC linkers but the exposed $H_{c''}$ pyrrole protons are not coupled with the distant XDC protons.

ESI-MS Analysis. The ESI-MS analysis revealed (Fig. S6) the characteristic m/z peaks of $[M-2TfO]^{2+}$ species of bow-tie complexes. The ZnTPP-based **BT1–BT4** and H₂TPP-based **BT2'**, **BT3'**, and **BT4'** displayed the respective $[M-2TfO]^{2+}$ peaks at $m/z = 1497.09, 1593.58, 1517.01, 1567.05, 1561.50, 1485.52,$ and 1535.60 , but no peak corresponding to any tetragonal prisms. Interestingly, the previously reported³⁰ ESI-MS profile of a tricomponent SCC featuring $(Et_3P)_2Pt^{II}$, H₂TPP, and BDC revealed a prominent m/z peak at 1481.38 , which possibly represented the $[M-2PF_6]^{2+}$ species of the bow-tie complex $[\{(Et_3P)_2Pt\}_4(H_2TPP)(BDC)_2]^{4+} \cdot 4(PF_6^-)$. Although previous reports have assigned certain m/z peaks to $[\{(Et_3P)_2Pt\}_8(M'TPP)_2(XDC)_4]^{8+} \cdot 8(X^-)$ prisms,^{30,33,34,56,64} in the light of our extensive X-ray crystallographic and NMR analyses, those peaks could be attributed to dimers of $[\{(Et_3P)_2Pt\}_4(M'TPP)(XDC)_2]^{4+} \cdot 4(X^-)$ bow-tie complexes (after the loss of certain counterions) instead of any prisms. Thus, in the light of the discovery of M'TPP-based tricomponent bow-tie complexes and a better understanding of the corresponding tetragonal prisms were not formed, it appears that careful 2D NMR and X-ray crystallographic studies are vital for accurate structural characterization of similar tricomponent SCCs.

Molecular Recognition via π -Donor/Acceptor Charge Transfer Interaction. Having assembled and accurately characterized M'TPP-based novel tricomponent bow-tie complexes, we turned our attention to explore their molecular recognition capabilities. We hypothesized that the electron-rich M'TPP core of bow-tie complexes will selectively recognize electron deficient π -systems through π -donor/acceptor CT interaction but not other π -donors. To test this hypothesis, we employed **BT4** ($[\{(Et_3P)_2Pt\}_4(ZnTPP)(NDC)_2]^{4+}$) as a host, which contains the least strained ZnTPP ligand (i.e., $\angle(N_{Py}-Zn-N_{Py})_{int} \approx \angle(N_{Py}-Zn-N_{Py})_{ext} \approx 88-89^\circ$) because of the intermediate length of rigid NDC linker, and a highly electron-deficient HATHCN and an electron-rich pyrene compounds as potential guests.

The **BT4**•HATHCN complex formation was evident from the ¹H NMR titration experiment (Fig. 7a). The characteristic H_{c'} and H_{c''} (pyrrole) protons of ZnTPP core gradually shifted up-field with the increasing amount of HATHCN (it has no proton), revealing that the ZnTPP core was shielded by cofacially aligned HATHCN. The greater up-field shift / shielding of H_{c''} signal ($\Delta\delta = 0.15$ ppm) than H_{c'} signal ($\Delta\delta = 0.06$ ppm) suggests that HATHCN is positioned more above the exposed pyrrole rings than sterically crowded (by NDC) enclosed pyrrole rings. The protons on the pyridyl and NDC rings, which are aligned almost perpendicularly to the Zn-porphyrin and HATHCN cores, shifted downfield, indicating that they were not shielded by HATHCN. The formation constant of **BT4**•HATHCN complex ($K_a = 2.5 \times 10^3$ M⁻¹, 3:7 CD₂Cl₂/CD₃NO₂, 25 °C) calculated from the ¹H NMR titration data (Fig. S7)⁶⁵ is comparable to that of similar CT complexes of other electron-rich Zn-porphyrin and π -acidic HAT-derivatives.⁶⁶ In contrast, during ¹H NMR titration of **BT4** with pyrene (Fig. 7b), none of their signals shifted, indicating the lack of

any meaningful interaction between the two electron-rich species. The ^{31}P NMR spectrum of **BT4** remained unchanged in the presence of HATHCN (Fig. S8), confirming that it did not interfere with the heteroleptic coordination of $(\text{Et}_3\text{P})_2\text{Pt}^{2+}$ corners with ZnTPP and NDC ligands. The ESI-MS analysis also revealed the $[\text{M-TfO}]^{3+}$ peak ($m/z = 1123.36$) of 1:1 **BT4**•HATHCN complex (Fig. S9).

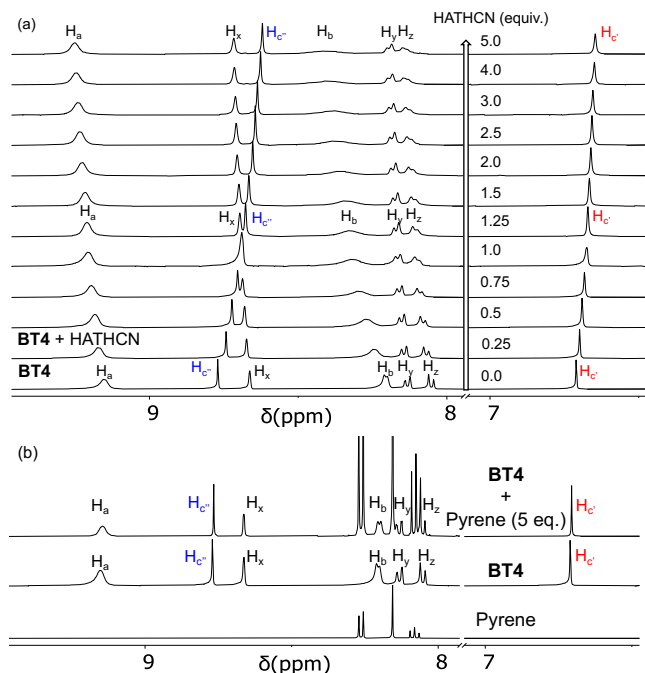


Fig. 8. The ^1H NMR titration data (500 MHz, 3:7 $\text{CD}_2\text{Cl}_2/\text{CD}_3\text{NO}_2$) of **BT4** with (a) HATHCN and (b) pyrene show gradual up-field shift, i.e., shielding of $\text{H}_{c'}$ and $\text{H}_{c''}$ (pyrrole) signals of ZnTPP core by the former but no such change with the latter.

The UV-Vis spectrum **BT4**•HATHCN complex displayed (Fig. 8a) a noticeable decrease of Soret and Q bands intensities (compared to free **BT4** spectrum) with a concomitant appearance of a characteristic broad CT band (650–870 nm) centered at ~ 725 nm, indicating ZnTPP/HATHCN π -D/A CT interaction. In addition, in the **BT4**•HATHCN CT complex, the Q bands of ZnTPP (560 and 602 nm) and the longest wavelength absorption of HATHCN (322 nm), which correspond to $S_0 \rightarrow S_1$ transitions, were noticeably blue-shifted (~ 4 – 6 nm) compared to free species, a characteristic sign of face-to-face π -D/A interaction between these two complementary π -systems.⁶⁷ In contrast, no UV-Vis absorption change of **BT4** was observed in the presence of pyrene (Fig. 8b), indicating the lack of any meaningful electronic interaction between these two electron rich species.

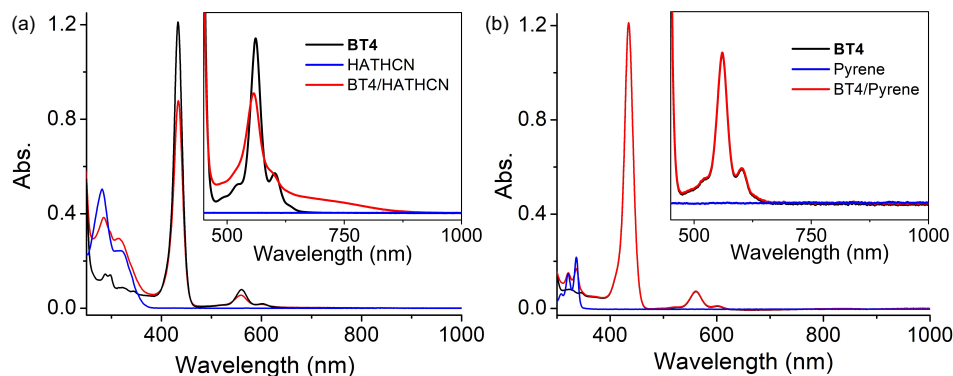


Fig. 8. The UV-Vis spectra of **BT4** (in CH_2Cl_2) in the presence of (a) HATHCN and (b) pyrene. Insets: The amplified 475–1000 nm regions show the appearance of CT band with HATHCN but not with pyrene.

The cyclic voltammetry analysis (Fig. S10) showed that in **BT4**•HATHCN complex, the first oxidation (anodic) peak of ZnTPP core shifted by +40 mV—from +1.14 V for free **BT4** to 1.18 V for **BT4**•HATHCN CT complex (vs. Ag/AgCl in 0.1 M $\text{Bu}_4\text{N}^+\cdot\text{TfO}^-$ / CH_2Cl_2 solution)—suggesting that the CT interaction between the electron-rich ZnTPP core and highly π -acidic HATHCN (LUMO: -4.8 eV)⁶⁸ made the oxidation of π -donor harder in the complex. In contrast, no significant shift of the first oxidation (anodic) peak of ZnTPP core was observed in the presence of electron-rich pyrene due the lack of such interaction.

Conclusions

In summary, we have demonstrated that Pt(II)-driven social self-assembly of a tetratopic M'TPP ligand and ditopic XDC linkers having different lengths (6.9–11.2 Å) and rigidities yielded novel 2D bow-tie complexes $[\{(\text{Et}_3\text{P})_2\text{Pt}\}_4(\text{M'TPP})(\text{XDC})_2] \cdot (4\text{TfO})$ featuring a M'TPP core and two parallel XDC linkers that were held together by four heteroligated $\text{Pt}^{\text{II}}(\text{N},\text{O})$ corners. The most direct and compelling evidence of bow-tie complexes came from the SXRD data, which were fully consistent with their NMR characteristics. The ^{31}P NMR spectra of the resulting tricomponent SCCs revealed the formation of heteroleptic $\text{Pt}^{\text{II}}(\text{N},\text{O})$ corners bearing one carboxylate and one pyridyl groups, while the ^1H and 2D NMR studies presented telltale signs of bow-tie structures by revealing that two opposite pyrrole rings carrying the more shielded H_c protons were located inside two isosceles triangles formed by two parallel XDC linkers, while the other two pyrrole rings bearing less shielded $\text{H}_{c''}$ protons remained exposed. The pyridyl arms of M'TPP deviated from their original projections in order to accommodate different XDC linkers having different lengths (~ 7 –11 Å) and rigidity into isosceles triangles. This led to the formation of 2D bow-tie complexes, which were entropically more favored over 3D tetragonal prisms. Thus, these comprehensive studies not only unveiled novel M'TPP-based bow-tie complexes, but also ruled out the

corresponding prism formation. In addition to assembling and accurately characterizing novel bow-tie-shaped tricomponent coordination complexes based on M'TPP ligands, we have demonstrated that these electron-rich species can bind highly π -acidic HATHCN through strong π -D/A CT interaction, but do not interact with π -donors. While the focus of the foregoing studies was to determine the accurate structures and compositions of M'TPP-based tricomponent SCCs and to demonstrate their molecular recognition capability involving π -donor/acceptor interaction, in the light of these new revelations it appears that some fascinating properties and functions, such as photodynamic cancer therapy that were previously attributed to M'TPP-based prisms, actually belonged to bow-tie complexes. Further studies of potential applications of these bow-tie complexes, such as light-harvesting and energy transduction systems are underway in our laboratory.

Conflicts of Interest

The authors declare no conflicts of interest.

Acknowledgments

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References

1. M. Fujita, J. Yazaki and K. Ogura, *J. Am. Chem. Soc.*, 1990, **112**, 5645.
2. P. J. Stang and D. H. Cao, *J. Am. Chem. Soc.*, 1994, **116**, 4981.
3. P. J. Stang, B. Olenyuk, J. Fan, and A. M. Arif, *Organometallics*, 1996, **15**, 904.
4. P. J. Stang and B. Olenyuk, *Acc. Chem. Res.*, 1997, **30**, 502.
5. A. J. Lees, and S.-S. Sun, *Inorg. Chem.*, 2001, **40**, 3154.
6. F. Würthner, C.-C. You and C. R. Saha-Möller, *Chem. Soc. Rev.*, 2004, **33**, 133.
7. N. C. Gianneschi, M. S. Masar III, and C. A. Mirkin, *Acc. Chem. Res.*, 2005, **38**, 825.
8. B. H. Northrop, Y.-R. Zheng, K.-W. Chi and P. J. Stang, *Acc. Chem. Res.*, 2009, **42**, 1554.
9. H. T. Chifotides, I. D. Giles and K. R. Dunbar, *J. Am. Chem. Soc.*, 2013, **135**, 3039.
10. M. A. Gordillo, P. A. Benavides and S. Saha, *Cryst. Growth Des.*, 2019, **19**, 6017.
11. K. Kumazawa, K. Biradha, T. Kusukawa, T. Okano, and M. Fujita, *Angew. Chemie Int., Ed.* 2003, **42**, 3909.
12. R. Chakrabarty, P. S. Mukherjee and P. J. Stang, *Chem. Rev.*, 2011, **111**, 6810.
13. T. R. Cook, Y.-R. Zheng and P. J. Stang, *Chem. Rev.*, 2013, **112**, 734.
14. T. R. Cook and P. J. Stang, *Chem. Rev.*, 2015, **115**, 7001.

15. S. Zarra, D. M. Wood, D. A. Roberts and J. R. Nitschke, *Chem. Soc. Rev.*, 2015, **44**, 419.
16. S. Fujii, T. Tada, Y. Komoto, T. Osuga, T. Murase, M. Fujita and M. Kiguchi, *J. Am. Chem. Soc.*, 2015, **137**, 5939.
17. W. M. Bloch and G. H. Clever, *Chem. Commun.*, 2017, **53**, 8506.
18. T. K. Ronson, W. Meng and J. R. Nitschke, *J. Am. Chem. Soc.*, 2017, **139**, 9698.
19. W. Brenner, T. K. Ronson and T. K. Nitschke, *J. Am. Chem. Soc.*, 2017, **139**, 75.
20. S. Saha, I. Regeni and G. H. Clever, *Coord. Chem. Rev.*, 2018, **374**, 1.
21. S. Pullen and G. H. Clever, *Acc. Chem. Res.*, 2018, **51**, 3052.
22. S. Chakraborty and G. R. Newkome, *Chem. Soc. Rev.*, 2018, **47**, 3991.
23. D. Bardhan and D. Chand, *Chem. Eur. J.*, 2019, **25**, 12241.
24. S. Kitagawa, R. Kitaura and S. Noro, *Angew. Chem. Int. Ed.*, 2004, **43**, 2334.
25. W. L. Leong and J. J. Vittal, *Chem. Rev.*, 2011, **111**, 688.
26. H. Furukawa, K. E. Cordova, M. O’Keeffe and O. M. Yaghi, *Science*, 2013, **341**, 974.
27. K. W. Chi, C. Addicott, A. M. Arif and P. J. Stang, *J. Am. Chem. Soc.*, 2004, **126**, 16569.
28. M. Wang, Y.-R. Zheng, K. Ghosh, P. J. Stang, *J. Am. Chem. Soc.*, 2010, **132**, 6282.
29. Y. R. Zheng, W. J. Lan, M. Wang, T. R. Cook and P. J. Stang, *J. Am. Chem. Soc.*, 2011, **133**, 17045.
30. Y.-R. Zheng, Z. Zhao, M. Wang, K. Ghosh, J. B. Pollock, T. R. Cook and P. J. Stang, *J. Am. Chem. Soc.*, 2010, **132**, 16873.
31. M. Wang, Y.-R. Zheng, T. R. Cook and P. J. Stang, *Inorg. Chem.*, 2011, **50**, 6107.
32. J. B. Pollock, T. R. Cook, G. L. Schneider and P. J. Stang, *Chem. Asian J.*, 2013, **8**, 2423.
33. Y. Shi, I. Sánchez-Molina, C. Cao, T. R. Cook and P. J. Stang, *Proc. Natl. Acad. Sci.*, 2014, **111**, 9390.
34. Y. Ye, T. R. Cook, S. P. Wang, J. Wu, S. Li, and P. J. Stang, *J. Am. Chem. Soc.*, 2015, **137**, 11896.
35. Y. Sun, C. Chen, J. Liu, L. Liu, W. Tuo, H. Zhu, S. Lu, X. Li and P. J. Stang, *J. Am. Chem. Soc.*, 2020, **142**, 17903.
36. G. Yu, T. R. Cook, Y. Li, X. Yan, D. Wu, L. Shao, J. Shen, G. Tang, F. Huang, X. Chen and P. J. Stang, *Proc. Natl. Acad. Sci.*, 2016, **113**, 13720.
37. Z. Li, X. Yan, F. Huang, H. Sepehrpour and P. J. Stang, *Org. Lett.*, 2017, **19**, 5728.
38. M. Zhang, M. L. Saha, M. Wang, Z. Zhou, B. Song, C. Lu, X. Yan, X. Li, F. Huang, S. Yin and P. J. Stang, *J. Am. Chem. Soc.*, 2017, **139**, 5067.
39. G. Yu, S. Yu, M. L. Saha, J. Zhou, T. R. Cook, B. C. Yung, J. Chen, Z. Mao, F. Zhang, Z. Zhou, Y. Liu, L. Shao, S. Wang, C. Gao, F. Huang, P. J. Stang and X. Chen, *Nat. Commun.*, 2018, **9**, 4335.
40. Y. Sun, Y. Yao, H. Wang, W. Fu, C. Chen, M. L. Saha, M. Zhang, S. Datta, Z. Zhou, H. Yu, X. Li and P. J. Stang, *J. Am. Chem. Soc.*, 2018, **140**, 12819.
41. X. Chang, Z. Zhou, C. Shang, G. Wang, Z. Wang, Y. Qi, Z.-Y. Li, H. Wang, L. Cao, X. Li, Y. Fang and P. J. Stang, *J. Am. Chem. Soc.*, 2019, **141**, 1757.

42. X. Chang, Z. Zhou, C. Shang, G. Wang, Z. Wang, Y. Qi, Z. Y. Li, H. Wang, L. Cao, X. Li, Y. Fang and P. J. Stang, *J. Am. Chem. Soc.*, 2019, **141**, 1757.
43. Y. Sun, C. Chen, J. Liu and P. J. Stang, *Chem. Soc. Rev.*, 2020, **49**, 3889.
44. S. Ghosh and P. S. Mukherjee, *Inorg. Chem.*, 2009, **48**, 2605.
45. A. K. Bar, G. Mostafa and P. S. Mukherjee, *Inorg. Chem.*, 2010, **49**, 7647.
46. D. Samanta, S. Shanmugaraju, S. A. Joshi, Y. P. Patil, M. Nethaji and P. S. Mukherjee, *Chem. Commun.*, 2012, **48**, 2298.
47. C. García-Simón, M. Garcia-Borràs, L. Gómez, I. Garcia-Bosch, S. Osuna, M. Swart, J. M. Luis, C. Rovira, M. Almedia, I. Imaz, D. Maspoch, M. Costas and X. Ribas, *Chem. Eur. J.*, 2013, **19**, 1445.
48. C. García-Simón, M. Garcia-Borràs, L. Gómez, T. Parella, S. Osuna, J. Juanhuix, I. Imaz, D. Maspoch, M. Costas and X. Ribas, *Nat. Commun.*, 2014, **5**, 5557.
49. C. Colombar, G. Szalóki, M. Allain, L. Gómez, S. Goeb, M. Sallé, M. Costas and X. Ribas, *Chem. Eur. J.*, 2017, **23**, 3016.
50. C. Colombar, V. Martin-Diaconescu, T. Parella, S. Goeb, C. García-Simón, J. Lloret-Fillol, M. Costas and X. Ribas, *Inorg. Chem.*, 2018, **57**, 3529.
51. C. Fuertes-Espinosa, A. Gómez-Torres, R. Morales-Martínez, A. Rodríguez-Fortea, C. García-Simón, F. Gándara, I. Imaz, J. Juanhuix, D. Maspoch, J. M. Poblet, L. Echegoyen and X. Ribas, *Angew. Chem. Int. Ed.*, 2018, **57**, 11294.
52. C. Colombar, C. Fuertes-Espinosa, S. Goeb, M. Sallé, M. Costas, L. Blancafort and X. Ribas, *Chem. Eur. J.*, 2018, **24**, 4371.
53. C. García-Simón, A. Monferrer, M. Garcia-Borràs, I. Imaz, D. Maspoch, M. Costas and X. Ribas, *Chem. Commun.*, 2019, **55**, 798.
54. V. Vajpayee, S. Bivaud, S. Goeb, V. Croué, M. Allain, B. V. Popp, A. Garci, B. Therrien and M. Sallé, *Organometallics.*, 2014, **33**, 1651.
55. E. Ubasart, O. Borodin, C. Fuertes-Espinosa, Y. Xu, C. García-Simón, L. Gómez, J. Juanhuix, F. Gándara, I. Imaz, D. Maspoch, M. Delius and X. Ribas, *Nat. Chem.*, 2021, DOI: 10.1038/s41557-021-00658-6.
56. G. Yu, Y. Ye, Z. Tong, J. Yang, Z. Li, B. Hua, L. Shao and S. Li, *Macromol. Rapid Commun.*, 2016, **37**, 1540.
57. C. Li, H. Nian, Y. Dong, Y. Li, B. Zhang and L. Cao, *Inorg. Chem.*, 2020, **59**, 5720.
58. Z. Yang, Y. Wang, X. Liu, R. T. Vanderlinden, R. Ni, X. Li and P. J. Stang, *J. Am. Chem. Soc.*, 2020, **142**, 13689.
59. Y. Hou, Z. Zhang, S. Lu, J. Yuan, Q. Zhu, W. P. Chen, S. Ling, X. Li, Y.-Z. Zheng, K. Zhu and M. Zhang, *J. Am. Chem. Soc.*, 2020, **141**, 18763.
60. M. Eddaoudi, J. Kim, N. Rosi, D. Vodak, J. Wachter, M. O’Keeffe and M. Yaghi, *Science*, 2002, **295**, 469.
61. S. Lipstman and I. Goldberg, *Acta Cryst. Sect. C.*, 2009, **65**, 371.
62. L. R. Dinelli, G. Von Poelhsitz, E. E. Castellano, J. Ellena, S. E. Galembeck and A. A. Batista, *Inorg. Chem.*, 2009, **48**, 4692.
63. A. N. Oldacre, M. R. Crawley, A. E. Friedman and T. R. Cook, *Chem. Eur. J.* 2018, **24**, 10984.

- 64. Z. Zhao, Z. Zhang, H. Wang, X. Li and M. Zhang, *Isr. J. Chem.*, 2019, **59**, 299.
- 65. (a) Thordarson, P. *Chem. Soc. Rev.* 2011, **40**, 1305, (b) www.supramolecular.org.
- 66. T. Aoki, H. Sakai, K. Ohkibo, T. Sakanoue, T. Takenobu, S. Fukizumi and T. Hasobe, *Chem. Sci.*, 2015, **6**, 1498.
- 67. C. R. Martinez and B. L. Iverson, *Chem. Sci.*, 2012, **3**, 2191.
- 68. G. Aragay, A. Frontera, V. Lloveras, J. Vidal-Gancedo and P. Ballester, *J. Am. Chem. Soc.* 2013, **135**, 2620.

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