

Reactant-Induced Photoactivation of a Gold-Catalyzed Csp^2 - Csp Cross-Coupling Leading to Indoles

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Abstract. A novel access to 2,3-disubstituted indoles from *o*-alkynyl aniline and iodoalkyne derivatives *via* a gold-catalyzed sequence under visible-light irradiation and in the absence of an exogenous photocatalyst was uncovered. A wide scope of the process was observed. Of note, 2-iodo-ynamides that have never been used in any organometallic cross-coupling reaction could be used as electrophiles. The resulting *N*-alkynyl indoles lend themselves to post-functionalization affording valuable scaffolds, notably benzo[*a*]carbazoles. Mechanistic studies converged on the fact that a potassium sulfonyl amide generates emissive aggregates in the reaction medium. Static quenching of these aggregates by a vinylgold(I) intermediate yields to an excited state of the latter, which can react with an electrophile *via* oxidative addition and reductive elimination to forge the key C-C bond. This reactant-induced photoactivation of an organogold intermediate opens rich perspectives in the field of cross-coupling reactions.

Introduction. In the quest of synthetic efficiency featuring notably step economy, dual catalysis through the merger of transition metal catalysis and photocatalysis has constituted undoubtedly a major advance.¹ As early as in 2007, Osawa observed an acceleration of a copper-free palladium-catalyzed Sonogashira coupling in the presence of $Ru(bpy)_3^{2+}$.² It was claimed that the latter facilitates the oxidative addition step but no mechanistic evidence was provided. In 2011, the Sanford group highlighted that the photoreduction of an aryldiazonium salt *via* single electron transfer (SET) could be coupled to a Pd(II)/Pd(IV) manifold involving a directed C-H activation to provide bis-aryl products. The design of these transformations relied on the intertwining of a photoredox cycle with the palladium cycle, which ensures the adequate electronic shuttle.³ This seminal report launched the field of metallaphotocatalysis¹ which found valuable applications, notably in gold(I) catalysis a fertile field of investigations over the last two decades.^{4,5} Indeed, in order to access to higher molecular complexity,^{6,7} the frequently encountered protodeauration step has to be circumvented and the post-functionalization of organogold intermediates⁸ *via* C-C bond formation appears highly desirable. Nevertheless, in contrast to the commonly used metals in catalysis and due the high redox potential of the Au(I)/Au(III) couple,⁹ the oxidative addition at gold(I) is difficult and there has been a need to devise pathways to promote it.^{10,11,12} While ligand design for gold(I) has emerged as a possibility,¹³ versatile catalytic cross-coupling reactions, notably thanks to the use of hemilabile (P,N) ligands, have been recently worked out.¹⁴ In addition, Au(I)/Au(III) cycles have also been sustained *via* photoredox catalysis using aryl-diazonium or iodonium partners as SET partners.^{15,16,17,18,19} Synthetically relevant

sequences have resulted but the question to know how those actually work is still pending, notably because some of these reactions proceed in photocatalyst-free conditions.^{20,21,22}

Recently, we have reported a new reactivity paradigm by providing the first example of an energy transfer promoted oxidative addition at gold(I).^{23,24} Upon visible-light irradiation, an iridium photocatalyst triggers *via* triplet sensitization the oxidative addition of an alkynyl iodide onto a vinylgold(I) intermediate to deliver Csp^2 - Csp coupling products after reductive elimination. Mechanistic and modelling studies suggest that no redox pathway is involved but that an energy transfer^{25,26,27} takes place to generate a triplet state of the vinylgold intermediate. The latter can engage in an oxidative addition with an alkynyl or a vinyl iodide. This novel mode of activation in gold homogenous catalysis was applied in several dual catalytic processes. For instance, alkynylbenzofuran derivatives like **3-O** could be obtained from *o*-alkynylphenol **1-O** and iodoalkyne **2a** in the presence of a catalytic mixture of a gold(I) (**[AuCF₃]**: ($[p\text{-CF}_3\text{Ph}]_3\text{PAuCl}$)) and an iridium(III) complex ($[\text{Ir}(\text{dF}(\text{CF}_3)\text{ppy})_2(\text{dtbbpy})\text{PF}_6]$) under blue LEDs irradiation (Fig. 1a).

While we were studying possible synthetic extensions of this chemistry, we made some puzzling observations when comparing the reactivity of *o*-alkynylphenol precursor **1-O** vs *o*-alkynyl *N*-tosyl amide **1a**. In iridium-free conditions, but in the presence of 5 mol% **[AuCF₃]** and K_2CO_3 under blue LEDs irradiation, only a very little amount of **3-O** (5%) was observed (Fig. 1b). In sharp contrast, these iridium-free reaction conditions proved to be perfectly suited for the formation of indole **3aa** (64% yield) from NTs amide **1a**. Structure of **3aa** was confirmed by X-ray diffraction (XRD) analysis.²⁸

This finding led us to examine the generality of this reaction since indoles are key motifs in many relevant natural products and are privileged scaffolds in medicinal chemistry, agrochemical industry, personal cares as well as dyes.^{29,30,31} It was also the occasion to investigate the mechanism and notably how the heteroatomic nucleophilic component (*N*- vs *O*-nucleophile) alters the photochemical activation process in these exogenous photocatalyst-free transformations.

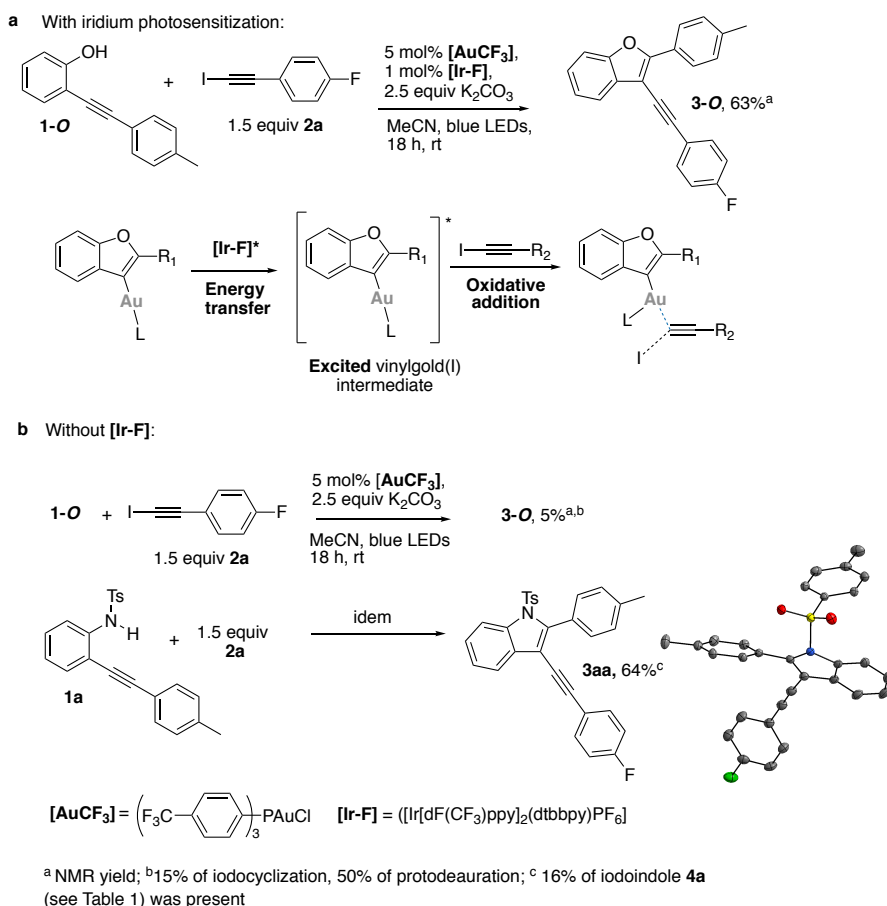


Fig. 1 Gold(I) catalysis under visible-light. **a**, Previous work²³: alkynylation of *o*-alkynylphenols *via* photosensitized oxidative addition. **b**, Contrasting reactivity between the phenol and the sulfonamide precursors **1-O** and **1a**

Results and discussions. We screened some variations on the reaction parameters in order to check if we were working in the most favorable set of conditions (Table 1). We first checked that the use of PPh_3AuCl (entry 1) instead of $[\text{AuCF}_3]$ was detrimental for the reaction since only 50% of conversion was observed and only 8 % of **3aa** was detected on the crude ^1H NMR spectrum, accompanied by 10% of iodoindole **4a** as iodocyclization by-product. The addition of 1 mol% $[\text{Ir-F}]$ was productive in **3aa** but not as much as in photosensitizer free conditions (38% vs 64%, entries 2 & 4) and the selectivity **3aa** vs **4a** was lower. The addition of 10 mol% benzophenone as photosensitizer proved to have no effect in terms of yield and selectivity for **3aa** (entry 3). In order to reduce the quantity of **4a**, we diminished the source of iodine (iodoalkyne **2a**) in the reaction conditions. Engaging a 1:1.1 ratio of **1a** and **2a** (entry 5) proved to be as yielding as previous conditions (1:1.5 ratio, entry 4) but more selective toward **3aa**. Variation of the counter ion gave contrasted results and confirmed the supremacy of the potassium cation over sodium (entry 6), caesium (entry 7) and calcium (entry 8), all three resulting in poor or no conversion. Addition of 18-C-6 caused a consequent yield decrease (only 14% of **3aa**, entry 9). While we showed that the reaction could be run by using a preformed potassic salt **1a-K** resulting from the deprotonation with KH of **1a** (50% of **3aa**, Fig. 2), addition of 18-C-6 was also deleterious in that case (35% of **3aa** accompanied by 10% of protodeauration product **5a**). All attempts with an amide (acetamide, trifluoroacetamide) or a carbamate (Boc) precursor failed to give any product. Overall, this series of findings confirmed us in using a sulfonamide nucleophile and a potassium base for further developments.

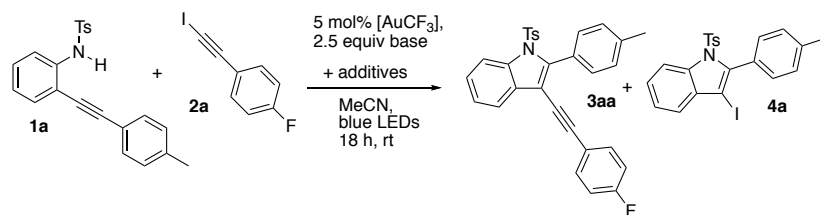


Table 1. Optimization of the formation of indole **3aa**

Entry	1a:2a Ratio	Base	Additives ^[c]	3aa , yield (%) ^[a]	4a , yield (%) ^[a]
1 ^[b]	1:1.5	K ₂ CO ₃	-	8	10
2 ^[c]	1:1.5	K ₂ CO ₃	1 mol% [Ir]	38	18
3	1:1.5	K ₂ CO ₃	10 mol% PhCOPh	61	16
4	1:1.5	K ₂ CO ₃	-	64	16
5	1:1.1	K ₂ CO ₃	-	62	10
6	1:1.1	Na ₂ CO ₃	-	4	10
7	1:1.1	Cs ₂ CO ₃	-	10	6
8	1:1.1	CaCO ₃	-	0	0
9	1:1.1	K ₂ CO ₃	1 equiv 18-C-6	14	16

^[a] NMR yield using 1,3,5-trimethoxybenzene as internal standard; all reactions run on a 0.1 mmol except for entries 3 and 4 (0.3 mmol). ^[b] 5 mol% PPh₃AuCl was used instead of 5 mol% **[AuCF₃]**. ^[c] **[Ir]** = [Ir(dF(CF₃)ppy)₂(dtbbpy)]PF₆

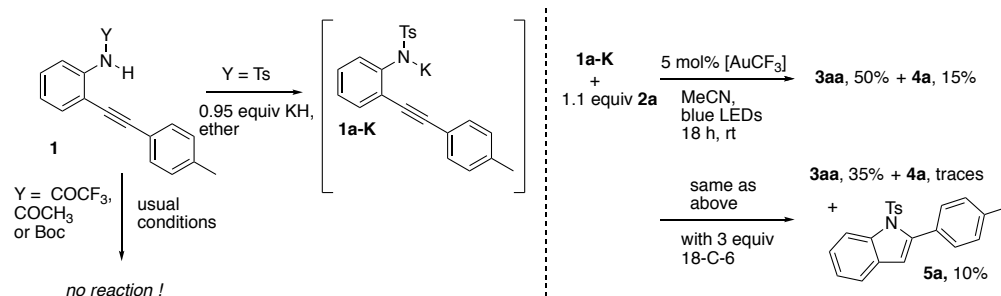


Fig. 2 Complementary elements of reactivity of aniline precursors 1. Preforming the potassic salt **1a-K** is possible to obtain the cross-coupling reaction. Of note, whatever the reaction conditions, no product was observed with an amide or a carbamate precursor.

Mechanistic studies. In order to extend the scope of possible substrates, we wished to get some insight into the mechanism of this transformation. The fact that a photo-triggered event occurs without any photocatalyst logically led us to consider the optical properties of the different partners present in the reaction medium. UV-vis absorption spectra were first recorded (See Fig. S1) and evidenced that the starting materials **1a** and **2a** as well as the cross-coupling product **3aa** at 50 μM were unable to absorb the light of the blue LED used ($\lambda_{\text{max}} = 450 \text{ nm}$). Species presumably formed *in situ* were also studied such as vinylgold(I) intermediate **6** (Fig. 3). The latter was smoothly obtained in 92% yield by reacting an equimolar mixture of **1a** and **[AuCF₃]** in the presence of K₂CO₃ in MeCN. Crystals were collected and allowed a XRD analysis.²⁸ Tosylamide anion **1a-K**, supposedly originating from the deprotonation of **1a** (pKa ≈ 10) in the basic medium was also examined.

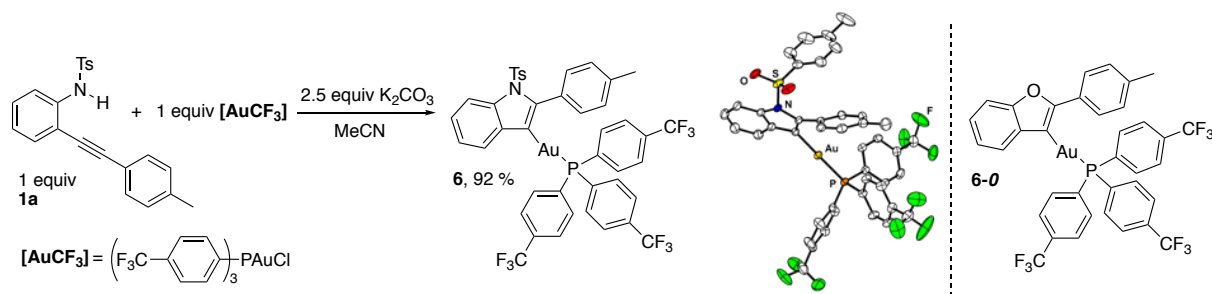


Fig. 3 Synthesis of the putative vinylgold(I) intermediate **6**

Complex **6** did not exhibit absorption above 400 nm at 50 μM and comparison between **6** and **6-O** showed both vinylgold(I) intermediates exhibit a similar absorption profile (See Fig. S2). In contrast, vinylgold **6** proved much more emissive than **6-O**, exhibiting a wide emission band at 400 nm (See Fig. S3). Thorough analysis of the photophysical properties of **1a-K** provided important information. First, we observed a dependence of the absorption properties on the concentration of **1a-K**. Indeed, by increasing the latter from 10 μM to 1 mM (Fig. 4a and S4a), the narrow absorption band at 310 nm collapsed and the appearance of a new wide band centered at 350 nm was observed. Second, upon increasing **1a-K** concentration, the emission band at 383 nm dropped at the benefit of a wide emission band centred at 450 nm (Fig. 4b and S5). All these findings are in line with the formation of aggregates of **1a-K**, which could account for chromatic alteration and excimer formation in a regime of strong coupling between the chromophores.³² In contrast, no such phenomenon was found with a potassic phenolate salt (See S4b).

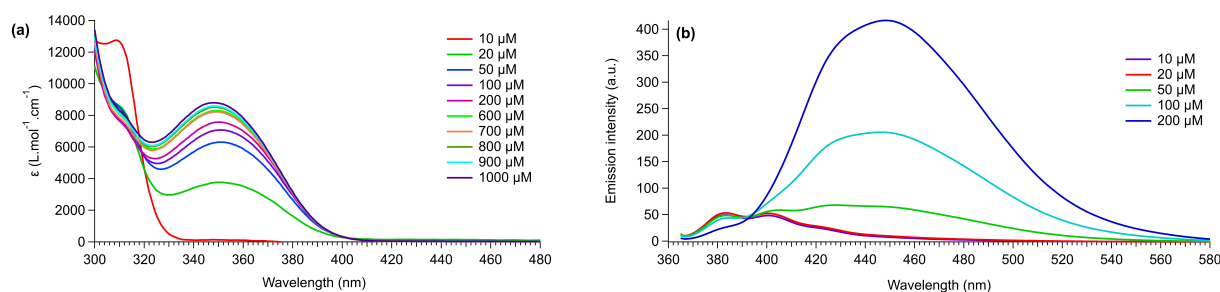


Fig. 4. Dependence of the molar absorption coefficients (a) and the luminescence spectrum (b) on the concentration of **1a-K** in MeCN at 298 K

XRD analysis of **1a-K**^[27] strengthened both the idea of the formation of aggregates and the crucial role of the potassium counterion since it revealed that each potassium cation is octacoordinated (Fig. 5). Besides the electrostatic interaction with the N atom, each potassium connects to four oxygen atoms of four distinct sulfonamide anions and also to two alkyne moieties through η^2 -slipped complexes featuring 3.0973 - 3.2589 Å and 3.3117 - 3.1062 Å distances. Such η^2 -slipped complexes remain rare,^{33,34} but have been invoked in some reactions.³⁵ The last interaction consists in a η^1 -coordination between a carbon of the tolyl ring and the potassium cation with a distance of 3.3074 Å.³³ Conversely, one can understand the detrimental effect of 18-C-6 that would prevent the formation of this aggregate.

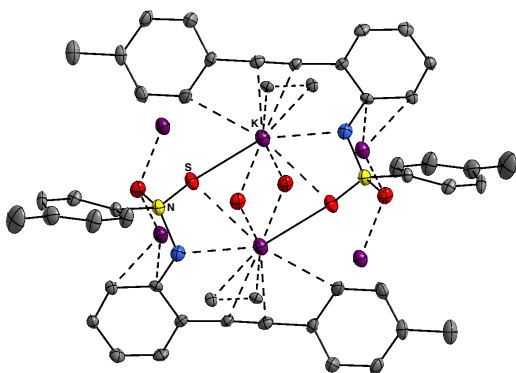


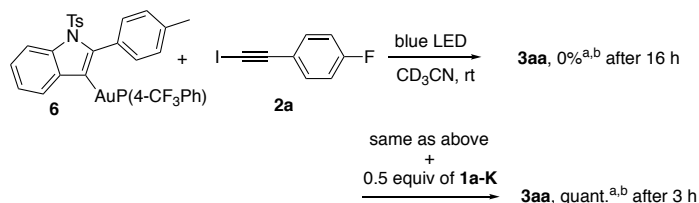
Fig. 5 XRD structure of 1a-K.²⁸ red atom: S, yellow atom: N, purple atom: K

Liquid-state NMR experiments were conducted to monitor the aggregation of **1a-K** in solution. First, the ¹H NMR spectra at concentrations ranging from 10 to 0.05 mM were recorded (Fig. S31). It was observed that upon decreasing the concentration, the four proton signals of the di-*ortho* substituted aniline ring broaden and shift. At 0.05 mM, these broadened signals are at the same chemical shifts than those in pure **1a**. These observations are consistent with monomeric **1a-K** at that low concentration. Hydrogen (in **1a**) and potassium (in **1a-K**) atoms being similarly bounded to nitrogen, quadrupolar potassium (spin 3/2) alone generates signal broadening. The improvement of the signal resolution by increasing the concentration suggests that the molecules adopt progressively a different spatial organization. This is confirmed by acquiring Noe-1D experiments with 0.05, 0.5 and 10 mM solutions of **1a-K**. Clear dipolar interactions are observed between the *ortho*-aromatic proton of the tolyl group at 7.4 ppm with the two aromatic protons of the tosyl group at 7.2 and 7.7 ppm for **1a-K** at 0.5 and 10 mM, while none exists at 0.05 mM (Fig. S32). Such intermolecular contacts are in agreement with the XRD structure proposed in Fig. 5. Finally, the self-diffusion coefficient of **1a-K** was determined at various concentrations (10, 5, 1, 0.5, 0.1 and 0.05 mM) in acetonitrile-d₃ solutions at 295 K by the PGSE-NMR technique (Fig. S33). Its value drops from the most diluted to the most concentrated solution, which suggests the formation of aggregates in compliance with the substantial increases of the estimated hydrodynamic radii and molecular weight (Fig. S34).

The presence of the iodoalkyne **2a** does not alter the absorption properties of **1a-K** (See Fig. S7). Nevertheless, we observed by ¹⁹F NMR a significant shift of the peak associated to the fluorine atom of **2a** upon its titration by a related potassic amide **1g-K** at constant concentration of **2a** (See Fig. S8). This observation may suggest that **1a-K** and **2a** yield a complex engaged in a fast exchange with the free species.³⁶ We obtained a satisfactory fit of the data by applying a 1:1 association model and retrieved $K_a^{298K} = 69 \text{ M}^{-1}$ for the **1aK** : **2a** complex association constant at 298 K (See Fig. S11). This value is consistent with the results of Goroff and Diederich,³⁷ who reported halogen bonding between amines and iodo alkynes of type **2**.^{38,39} DFT calculations also supported the formation of an halogen bond between **1a-K** and **2a** and **2b** (See Fig. S23, S24 and S25).

By analogy with the phenol case,²³ we investigated the reaction of vinylgold **6** with iodoalkyne **2a** under blue LEDs irradiation. Thus, at 5 mM which is close to the maximum concentration of **6** in the preparative reactions (*ie* 3.3 mM of gold species, based on 0.3 mmol of **1a** and 5 mol% of **[AuCF₃]** in 4.5 mL of MeCN, see SI) almost no cross-coupling took place, since only traces of **3aa** were observed after 16 h of reaction (See Fig. 6, S12 S13 and S14). This observation marked a sharp contrast with the one-pot reaction of entry 5 of Table 1 that smoothly delivers **3aa**. We surmised from these findings that a component formed in the reaction medium must promote the

transformation. In view of the preceding series of experiments, we raised our suspicions on intermediate **1a-K**. Indeed, running the same NMR monitoring experiment in the presence of 0.5 equiv of **1a-K** vs **1a** and **6** revealed a contrasting evolution of the reaction (See Fig. S15 and S16). As early as after 10 min, 17 % of **3aa** was formed, 45% of **3aa** after 0.5 h. After 3 h, the NMR yield reached 100 % of **3aa**, based on the limiting one equivalent of **2a**.^{40,41}



^a [2a] = [6] = 5 mM; ^b Determined by ¹H NMR using 1,3,5-trimethoxybenzene as internal standard

Fig. 6 Impact of 1a-K on the direct irradiation of a mixture of 6 and 2a under blue LED irradiation

To gain further insight into the mechanism, we investigated the impact of **6** on the steady-state luminescence of the **1a-K** aggregate excited at 380 nm. We observed a marked luminescence drop when the concentration of **6** was increased (Fig. 7a and S18). Stern-Volmer analysis yielded a slope equal to $9.4 \cdot 10^3 \text{ mol}^{-1} \cdot \text{L}$ at 298 K (See Fig. 7b and S19). To properly interpret this value, we showed that the lifetime of **1a-K*** was below 20 ns (See Fig. S20), which suggests that the excited aggregated **1a-K*** is a singlet state. Hence, by assuming the quenching to be dynamic,²³ one would extract $k_{\text{Q}}^{298\text{K}} = 4.7 \cdot 10^{11} \text{ L} \cdot \text{mol}^{-1} \cdot \text{s}^{-1}$ as the lower limit of the rate constant of **1a-K*** fluorescence quenching.⁴² However, this value is significantly higher than the second-order rate constant for a reaction under control of diffusion evaluated to $k_{\text{MeCN}}^{298\text{K}} = 1.8 \cdot 10^{10} \text{ L} \cdot \text{mol}^{-1} \cdot \text{s}^{-1}$ (See SI, eq. 5).^{43,44} To further eliminate the possibility of a dynamic quenching, we showed that a drop of the temperature increased the quenching efficiency (See Fig. 7b and S21), which could not be explained by a quenching under control of diffusion since the solvent viscosity increases when the temperature drops. Therefore, we concluded the quenching to be static and associated to the formation of a non-emissive complex between the **1a-K** aggregate and **6**. The slope of the Stern-Volmer analysis was interpreted as the thermodynamic association constant at 25°C⁴⁵ and its temperature dependence was analyzed with the Van't-Hoff equation to retrieve $\Delta_r H^\circ = -106.2 \text{ kJ} \cdot \text{mol}^{-1}$ and $\Delta_r S^\circ = -299.44 \text{ J} \cdot \text{mol}^{-1} \cdot \text{K}^{-1}$ for the enthalpy and entropy of the **1a-K-6** complex formation respectively (See S22a,b). This would transpose in the reaction medium in such a way that the blue light-excited **1a-K** aggregates in contact with the vinylgold **6** would promote direct Dexter energy transfer. Further intersystem conversion would yield the triplet state **6***, which undergoes oxidative addition from iodo-alkynes **2** (See Fig. 8 and S26). After reductive elimination, the liberated LAuI could promote another catalytic cycle. In addition, DFT calculations supported the whole catalytic cycle and notably the photosensitized oxidative addition through bending of the iodoalkyne partner (see S26).

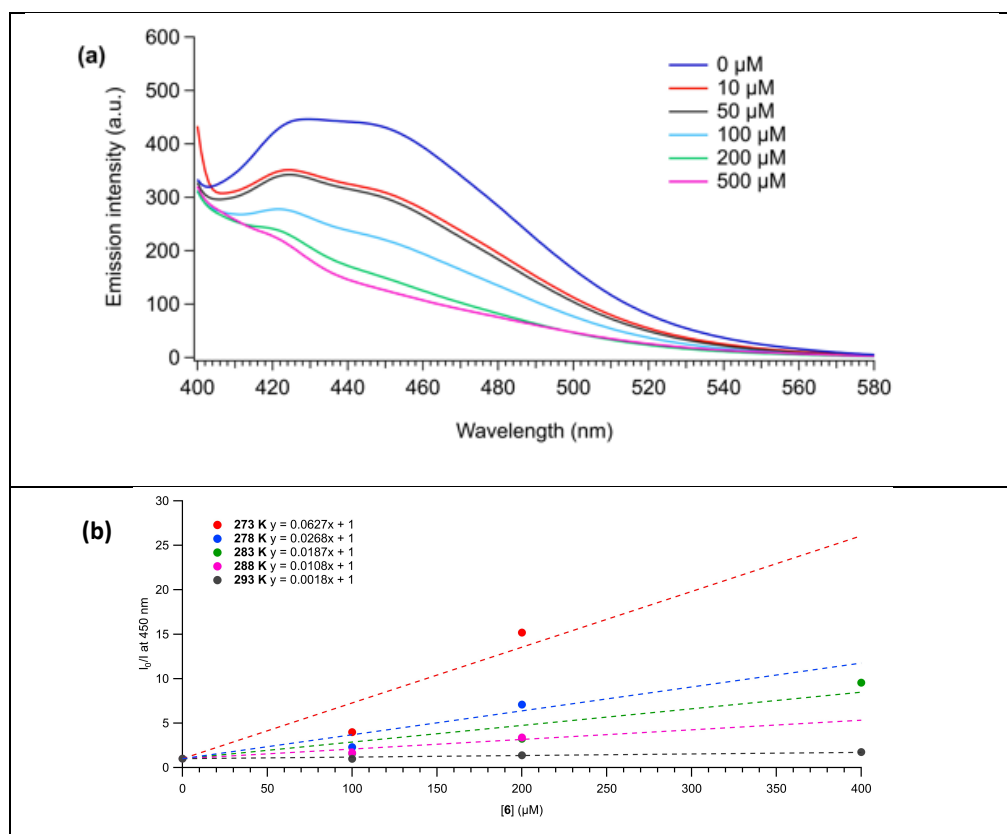
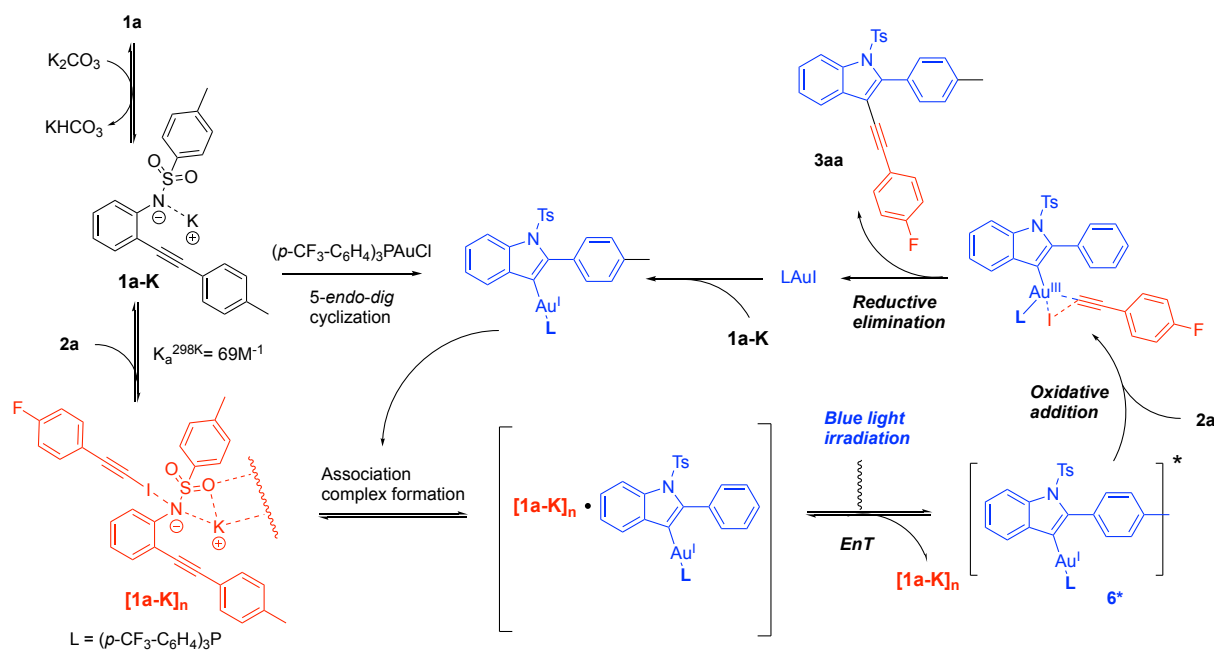


Fig. 7 Luminescence quenching of 1a-K by 6. a, Dependence of the fluorescence spectrum of **1a-K** on the **6** concentration upon light excitation at 380 nm. b, Effect of temperature on quenching efficiency. The slope of $I_0/I_0([6])$ continuously increases from $1.8 \cdot 10^3 \text{ L}\cdot\text{mol}^{-1}$ at 25°C to $6.3 \cdot 10^4 \text{ L}\cdot\text{mol}^{-1}$ at 0°C .



This cross-coupling process enters the examples of photochemical transformations under visible-light for which no exogenous photocatalyst is required to observe a photochemical activation. Several synthetic processes have been shown to follow such pathways and a large part of them have involved photoactive reactants or products,⁴⁶ electron donor-acceptor (EDA) complexes^{47,48,49} or the *in situ* generation of visible-light absorbing intermediates.⁵⁰ Herein, the aggregation of one reactant is the trigger for reactivity. This type of activation remains poorly described and the most closely related example is the aggregation-induced photoreaction of 1,2-diisocynoarenes.⁵¹

Scope of the reaction. Having determined the activation process that governs these transformations, we studied the scope of the reaction. We first varied the alkynyl iodide partners **2** (Fig. 9). The reaction proved to be workable on a series of arylalkynes substrates bearing a donor group on the aromatic ring (67% of **3ad** and 63% of **3ae**, both of them bearing a methoxy group) or with an electron-withdrawing group (56% of **3ai** with an ester group, 58% of **3aj** with a CF₃ group). In contrast, an alkyl chain on the alkyne gave a more modest yield (**3ak**, 38%).

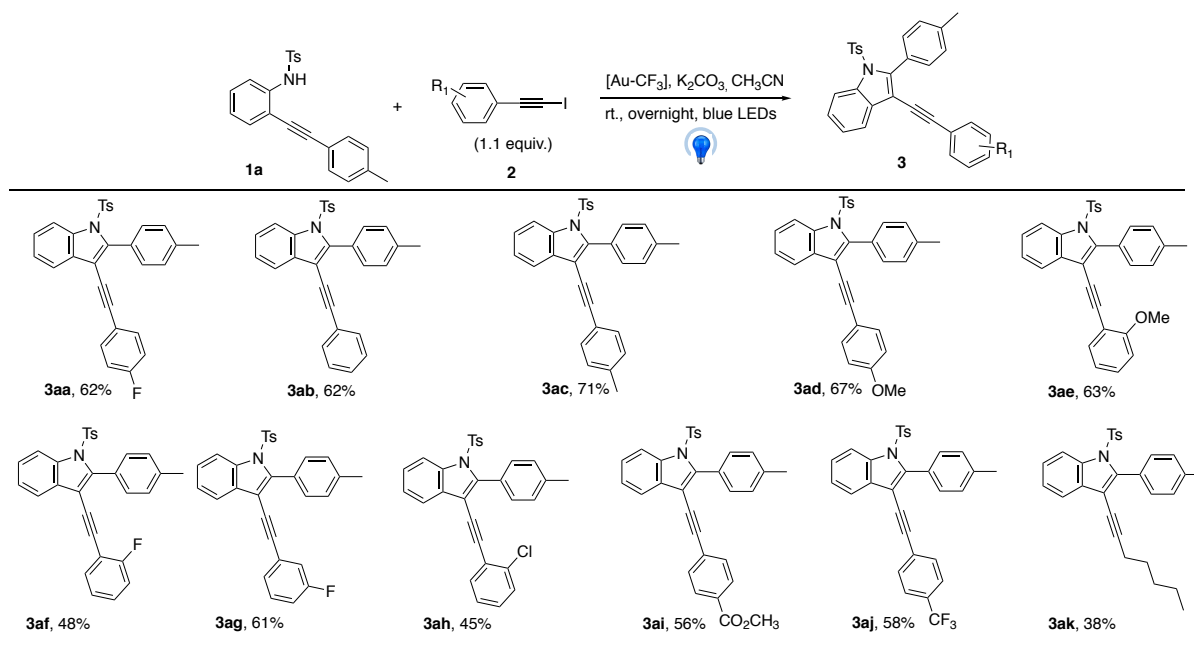
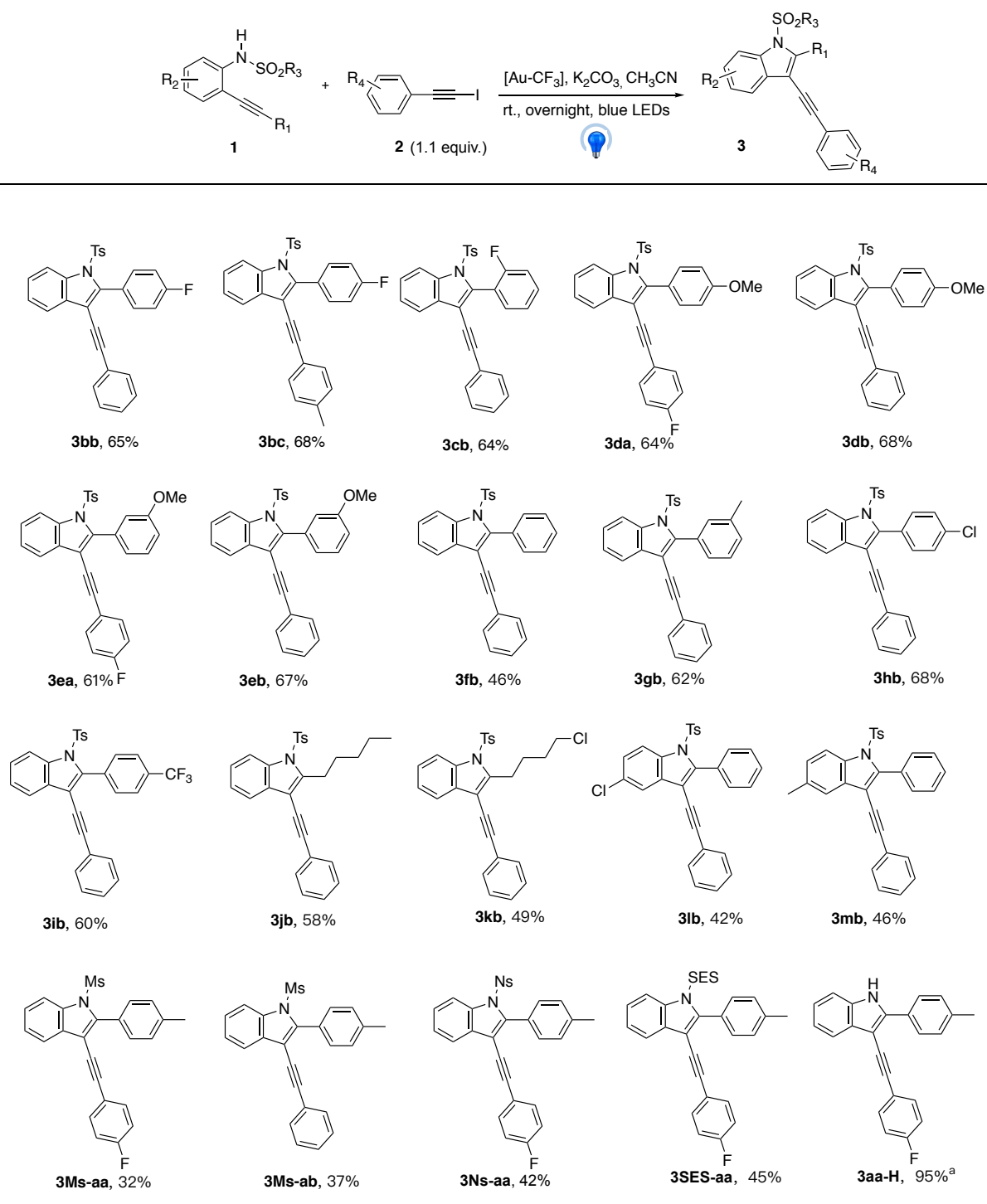


Fig. 9 Scope of alkynes

The aniline platform **1** could also be modulated (Fig. 10). Variation of the alkynyl substituent R₁ was first examined. When dealing with an aromatic alkyne, no much yield variation was observed since most yields of products **3** were above 60% even with more contrasting electronic demand (64% of **da** and 68% of **db** with a 4-methoxy group vs 60% of **3ib** with a 4-CF₃ group). Of note, the presence of a 4-methoxy group on the aromatic ring did not significantly result in protodeauration on the contrary to our previous arylative cyclization of *O*-alkynyl phenols.¹⁹ Donor alkyl groups on the alkyne were also competent providing 2-alkyl indoles **3jb** and **3kb** respectively in 58% and 49% yields. Introduction of a chlorine or a methyl group in 4-position of the aniline precursor gave average yields (< 50%) of products **3ib** and **3mb**. We also wished to survey the importance of a sulfonyl group attached on the nitrogen atom of the aniline precursor by varying the R₃ group. In general more modest yields of indole products were obtained with mesylate precursors that gave the corresponding products **3Ms-aa** in 32% yield or **3Ms-ab** in 37% yield. Although the *N*-detosylation of indoles is known,⁵² even more easily disposed sulfonamide groups were tested and 42% of nosyl-indole **3Ns-aa** and 45% of 2-

(trimethylsilyl)ethanesulfonyl (SES) indole **3SES-aa** were obtained. The latter could be easily converted into the free NH indole **3aa-H** by treatment with Bu₄NF as shown in Fig. 10.



^aAfter treatment with Bu₄NF from **3SES-aa**

Fig. 10 Scope of anilines

In order to extend the synthetic scope of this process and also to provide substrates for valuable post-functionalizations, we focused our attention on 2-iodo-ynamide **7** as a new electrophilic partner (Fig. 11). Interestingly, this substrate has been only described once by Danheiser who used it as a [2+2] cycloaddition

partner.⁵³ From a broader perspective, very few examples of metal-catalyzed cross-coupling of such halo-ynamides have been published, all of them from bromo- or chloro-ynamides with copper^{54,55,56} but none with gold. Gratifyingly, though 2-iodo-ynamide **7** reacted a little bit less efficiently than iodoalkynes **2** with anilines **1**, a series of functionalized scaffolds **8** (for instance **8b**²⁸) was obtained. Of note, a naphthyl aniline precursor could also be engaged to give adduct **8g** in 48%.

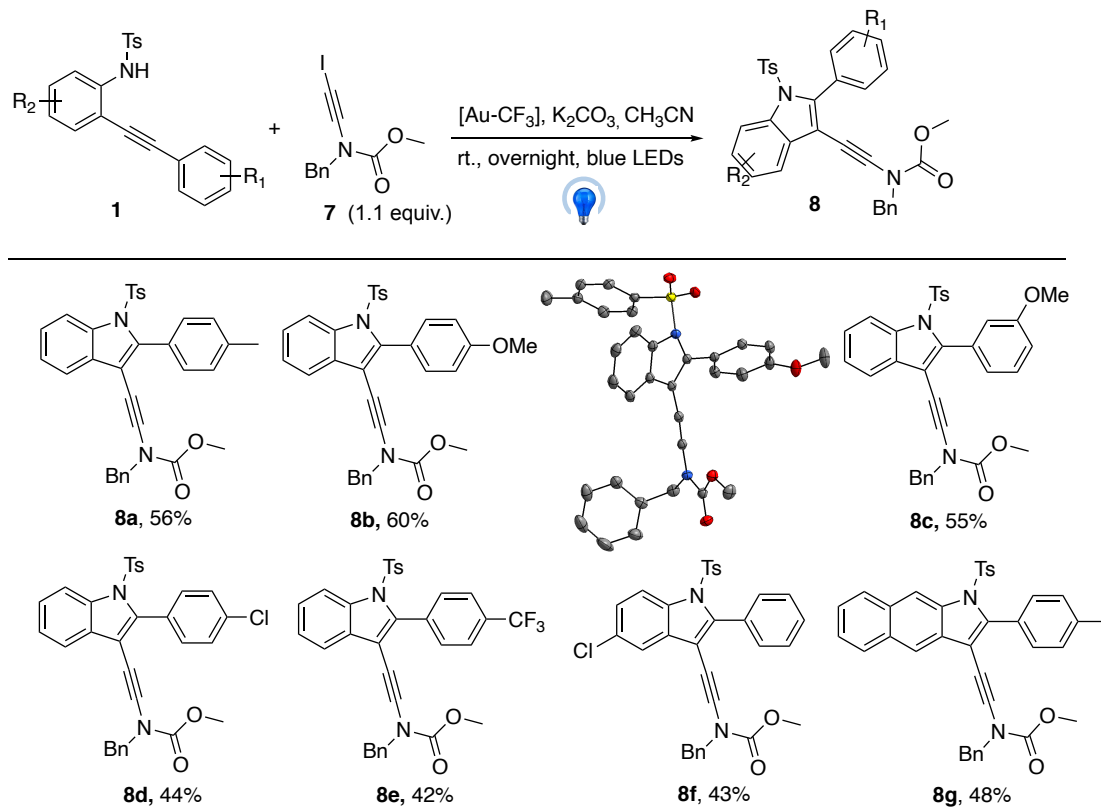
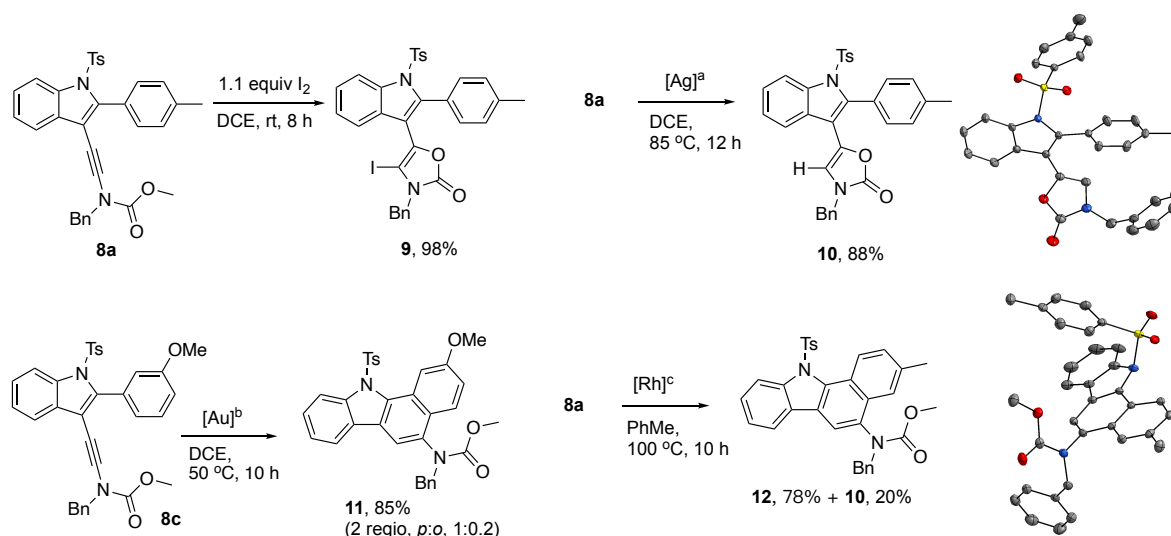


Fig. 11. 2-Iodo-ynamides as electrophiles

Taking advantage of the electron-richness of the ynamide moiety, we devised several transformations on polyfunctionalized substrates **8** (Fig. 12). While iodine smoothly promoted the formation of iodo-oxazolone **9**, a so far unknown type of derivatives, the corresponding Ag(I)-catalyzed cyclization in (wet) DCE provided the corresponding oxazolone **10**²⁸ via protodesilveration.^{57,58} Engaging the pendant aromatic ring in hydroarylation reaction was possible whether by gold(I) catalysis from **8c** or by cationic rhodium catalysis⁵⁹ from the less activated substrate **8a** to provide the corresponding carbazoles **10** and **12**.²⁸ Benzo[a]carbazoles derivatives are important structural units found in many natural products⁶⁰ and biologically active molecules^{61,62} as well as organic materials.⁶³



^a [Ag] = 20 mol% AgSbF₆; ^b [Au] = 20 mol% Ph₃PAuNTf₂; ^c [Rh] = 20 mol% RhCl(Ph₃P)₃, 20 mol% AgSbF₆

Fig. 12 Post-functionalizations

In conclusion, this study uncovers an unprecedented access to 2,3-disubstituted indoles from *o*-alkynyl aniline derivatives *via* a new type of gold-catalyzed sequence under visible-light irradiation. Subsequent to the formation of a vinylgold(I) intermediate, a photosensitized oxidative addition at gold(I) of an alkynyl iodide yields to Csp²-Csp cross-coupling reactions. Of note, no exogenous photocatalyst is required in these processes. On the contrary, it was found that one of the *in-situ* formed reactant, the potassium sulfonyl amide generates emissive aggregates in the reaction medium. Variation of the temperature in the fluorescence quenching studies established a static quenching of these aggregates by the vinylgold(I) intermediate to promote its excited state. After ISC to the triplet state of the latter, reaction with an electrophile *via* oxidative addition and reductive elimination forges the key C-C bond. Of note, 2-iodo-ynamides that have never been used in any organometallic cross-coupling reaction could be used as electrophiles. The resulting *N*-alkynyl indoles lend themselves to various post-functionalizations providing valuable scaffolds including benzo[*a*]carbazoles. The present reactant-induced photoactivation of an organogold intermediate confirms the versatility of the photosensitized oxidation process and it opens rich perspectives in the field of cross-coupling reactions for the assembly of complex frameworks.

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Keywords: gold catalysis • photosensitization • iodynamides • aggregates • indole

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