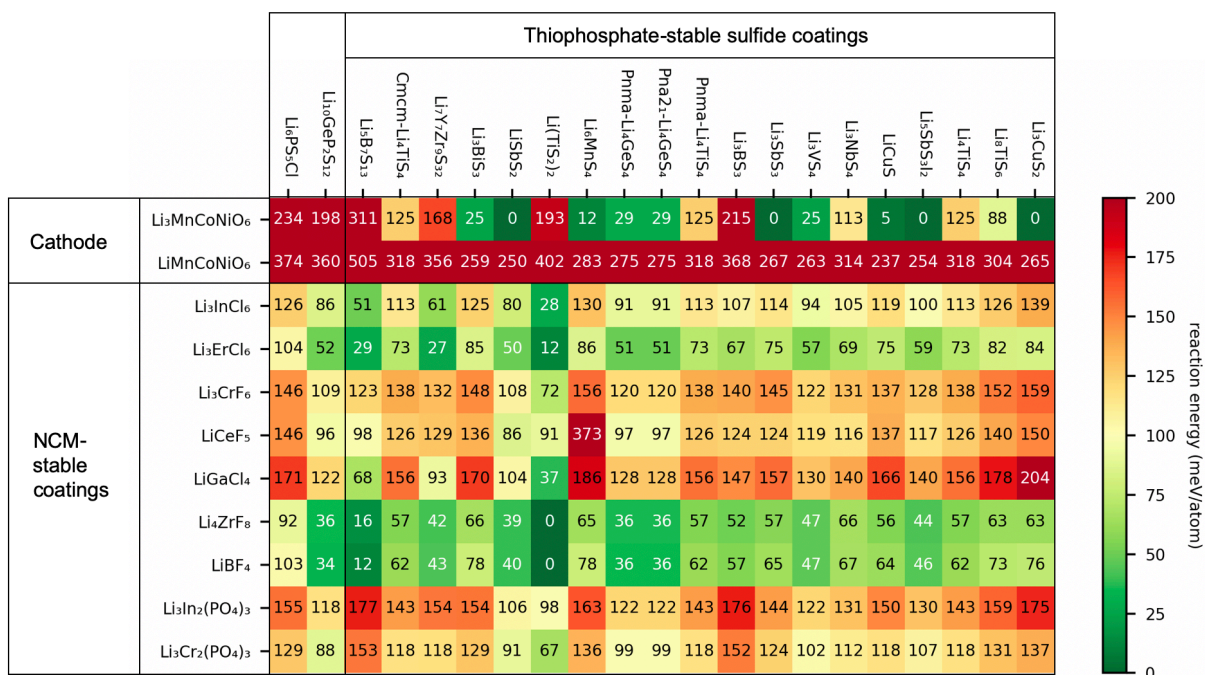


# Computational design of double-layer cathode coatings in all-solid-state batteries

Chuhong Wang,<sup>1</sup> Koutarou Aoyagi,<sup>1</sup> Tim Mueller<sup>1,\*</sup>



<sup>1</sup> Department of Materials Science and Engineering, Johns Hopkins University, Baltimore, Maryland 21218, United States

\*Corresponding Author

Tim Mueller

Department of Materials Science and Engineering

Johns Hopkins University

3400 N. Charles St.

Baltimore, MD 21218 USA

(410)516-5779

[tmueller@jhu.edu](mailto:tmueller@jhu.edu)

## ABSTRACT

All-solid-state lithium-ion batteries have great potential for improved energy and power density compared to conventional lithium-ion batteries. With extensive research efforts devoted to the development of inorganic superionic conductors, lithium thiophosphates stand out due to their high ionic conductivity and room-temperature processability. However battery rate performance still suffers from increased impedance attributed to the interfacial reactions between thiophosphate electrolyte and oxide electrodes. Stabilizing the interfaces with a protective coating layer has been proposed as a solution to the interfacial problem, but it is rare for a material to simultaneously exhibit fast ionic conductivity and chemical stability at battery interfaces. Here, we propose a double-layer coating design comprising a sulfide-based layer adjacent to the thiophosphate electrolyte accompanied by a layer that is stable against the oxide cathode. Based on a high-throughput thermodynamic stability screen and active learning molecular dynamics simulations, we identify several sulfide + halide couples that potentially outperform the known coating materials in interfacial stability as well as ionic conductivity. Several halides we identify have been recently identified as novel solid electrolyte candidates. We highlight the integration of fast ionic conductors  $\text{Li}_5\text{B}_7\text{S}_{13}$  ( $137 \text{ mS cm}^{-1}$ ),  $\text{Li}_7\text{Y}_7\text{Zr}_9\text{S}_{32}$  ( $6.5 \text{ mS cm}^{-1}$ ), and  $\text{Li}(\text{TiS}_2)_2$  ( $0.0008 \text{ mS cm}^{-1}$ ) which potentially reduces interfacial reactivity with minor loss of charge transfer rate through the thiophosphate electrolyte.

## 1.INTRODUCTION

Rechargeable batteries are a crucial component of efforts to reduce carbon emissions. Lithium ion batteries have dominated the market for portable electronics due to their high efficiency, low cost, high capacity, lack of memory effect, and long cycle life.<sup>1, 2</sup> The introduction of new solid-state electrolyte chemistries has the potential to further improve battery performance. Solid electrolytes enable the construction of bipolar cells with improved power and energy density.<sup>3, 4</sup> In addition, by mechanically suppressing dendrite formation and growth, solid electrolytes may enable higher specific energy through the use of Li-metal anodes.<sup>5, 6</sup>

Several structural families of solid Li-ion conductors have been identified over the past several decades.<sup>7</sup> Among the solid electrolyte candidates, lithium thiophosphates (e.g.  $\text{Li}_6\text{PS}_5\text{Cl}$ ,  $\beta\text{-Li}_3\text{PS}_4$ ,  $\text{Li}_{10}\text{GeP}_2\text{S}_{12}$ , and various thiophosphate glasses) have attracted the most attention for their superior Li-ion conductivity of  $10^{-3}$  to  $10^{-2} \text{ S cm}^{-1}$  at  $25^\circ\text{C}$  with a high lithium transference number of  $\sim 1$  and good deformability to compensate for volume changes associated with Li removal and insertion in the electrodes.<sup>8-10</sup> One of the main obstacles hindering the realization of high-performance solid-state batteries is the high internal resistance at the interface between the solid electrolyte and the active electrode materials. Both theoretical predictions and experimental characterizations have shown that the interfacial reactions between the thiophosphate-based electrolyte and the oxide-based cathode produce Li-ion insulating phases, which lead to significant degradation of the rate performance.<sup>11-17</sup> Mitigation of the internal impedance is critical to the commercialization of all-solid-state batteries, and the most efficient method to achieve this is through a protective coating on between the cathode active material and the electrolyte.

To stabilize the cathode-electrolyte interface, thus inhibiting undesirable side reactions, ideal interfacial coating materials should be thermodynamically stable against both the electrode and the electrolyte. Commonly applied coatings are primarily ternary oxides such as  $\text{LiNbO}_3$ ,<sup>18, 19</sup>  $\text{Li}_4\text{Ti}_5\text{O}_{12}$ ,<sup>20</sup>  $\text{Li}_2\text{SiO}_3$ ,<sup>21</sup>  $\text{Li}_4\text{SiO}_4$ ,<sup>22</sup>  $\text{Li}_2\text{ZrO}_3$ ,  $\text{LiTaO}_3$ ,<sup>23</sup> and  $\text{Li}_3\text{PO}_4$ ,<sup>24</sup> along with composites such as  $\text{Li}_2\text{BO}_3\text{-Li}_2\text{CO}_3$ ,<sup>25</sup>  $\text{LiInO}_2\text{-LiI}$ ,<sup>26</sup> and  $\text{Li}_2\text{O-ZrO}_2$ .<sup>27</sup> Based on first-principle studies, these oxide-based coatings effectively extend the electrochemical stability window of the solid electrolyte.<sup>28</sup> Automated computational screening of vast crystalline material databases enables efficient identification of additional coating candidates with high interfacial stability, thus greatly accelerating the discovery of candidate coating materials.<sup>12, 29-32</sup>

In addition to the thermodynamic considerations, high Li-ion conductivity is also a key factor in achieving high rate capability. The identification of a single coating material that has all desirable properties, including high stability and high lithium-ion conductivity, has not yet been achieved. Previous studies have suggested that the scarcity of ideal coating materials may be explained by the inverse correlation between Li-ion conductivity and electrochemical stability.<sup>12, 33</sup> Searching for the few outliers in a broad chemical space poses a challenge for material discovery.

To overcome this limitation, we explore a double-layer coating strategy to ease the stringent requirements on a single material, therefore enhancing the likelihood of finding suitable combinations of properties. In a previous study, we predicted that the interphase components that form stable interfaces with the thiophosphate electrolyte are likely to be either phosphates or sulfides.<sup>17</sup> The formation of phosphates is attributed to reductive decomposition due to  $\text{P}^{+5}/\text{P}^{-3}$  redox reactions at potentials of about 1.7 V versus  $\text{Li}^+/\text{Li}$ . This may be prevented by a protective oxide-free sulfide layer that inhibits the oxidation of phosphorous in the electrolyte.<sup>17</sup> In addition, we found several sulfide-based interphase components such as  $\text{Li}_4\text{GeS}_4$  and  $\text{Li}_3\text{PS}_4$  with relatively high Li-ion conductivity, consistent with the fact that sulfides generally possess higher ionic conductivity than oxides due to larger and more polarizable  $\text{S}^{2-}$  than  $\text{O}^{2-}$ .<sup>34</sup> Given their electrochemical properties and ionic conductivity, we search among the lithium sulfides for coatings that are thermodynamically compatible with the thiophosphate electrolytes  $\text{Li}_6\text{PS}_5\text{Cl}$  and  $\text{Li}_{10}\text{GeP}_2\text{S}_{12}$  while maintaining rate capabilities.

In comparison to thermodynamic compatibility, which can rapidly be evaluated by theoretical phase diagram prediction, diffusion kinetics in candidate coating materials are more difficult to investigate in a systematic, high-throughput way. The rate of Li-ion diffusion through commonly used coating materials is typically several orders of magnitude slower than through solid electrolytes.<sup>35</sup> Widely used approaches such as the climbing image nudged elastic band method<sup>36, 37</sup> or ab-initio molecular dynamics are either difficult to automate or prohibitively expensive for screening coating candidates.<sup>38</sup> To circumvent this problem, we employ a recently developed scheme using machine-learned interatomic potentials<sup>39</sup> to enhance the capacity of ionic conductor screening. To preserve the accuracy of the machine-learned potential and minimize the cost of generating training data, molecular dynamics simulations are coupled with on-the-fly machine

learning (LOTF-MD). Based on the activation energies for diffusion computed using LOTF-MD, we identify 14 sulfides predicted to have high lithium-ion conductivity, of which  $\text{Li}_5\text{B}_7\text{S}_{13}$ ,  $\text{Li}_4\text{TiS}_4$  (with Cmc symmetry),  $\text{Li}_7\text{Y}_7\text{Zr}_9\text{S}_{32}$ ,  $\text{Li}_3\text{BiS}_3$  are recognized for exceptionally high Li-ion conductivity comparable to the leading thiophosphate-based solid electrolytes. To our knowledge, several sulfides such as  $\text{Li}_4\text{TiS}_4$  (Cmc),  $\text{Li}_7\text{Y}_7\text{Zr}_9\text{S}_{32}$ ,  $\text{Li}_3\text{BiS}_3$  and  $\text{LiSbS}_2$  are characterized in terms of viable  $\text{Li}^+$  conduction the first time.

According to previous interphase studies, sulfides are susceptible to oxidation decomposition when exposed to oxide-based cathodes.<sup>14, 17, 40</sup> To optimize the interfacial rate performance, a coating layer that conducts lithium ions at sufficiently high speed and eliminates interfacial reactivity must be placed between the sulfide layer and the cathode. Here we focus on the cathode material  $\text{LiNi}_{1/3}\text{Mn}_{1/3}\text{Co}_{1/3}\text{O}_2$  (NCM), which is an attractive alternative to  $\text{LiCoO}_2$  due to its relatively low cost, high capacity, and better thermal stability.<sup>41</sup> NCM is currently one of the most promising cathode candidates for commercial applications, but sulfides in contact with NCM usually result in high Li-ion impedance attributed to the formation of sulfates.<sup>17, 40</sup> From a first-principle database of over 1500 Li-containing crystalline materials, we find 11 candidate NCM-stable coatings that are predicted to have sufficiently low activation barriers for Li diffusion. Most of the identified NCM coating candidates are halides, followed by a small fraction of polyanionic oxides. Among the lithium halides we find, the composition of  $\text{Li}_3\text{MX}_6$  ( $\text{X}=\text{Cl}, \text{F}$ ) is predicted to exhibit the highest ionic conductivity. These have recently attracted attention as electrolyte candidates for application in all-solid-state lithium batteries owing to high room-temperature ionic conductivity, wide electrochemical windows, and good thermodynamic compatibility with the oxide cathode.<sup>42</sup>

Guided by the interfacial compatibility, we suggest several pairs of sulfide + halide double-layer coatings with low chemical reactivity at the interface and reasonably high ionic conductivity. Notably, the ultrafast ionic conductors  $\text{Li}_5\text{B}_7\text{S}_{13}$  and  $\text{Li}_7\text{Y}_7\text{Zr}_9\text{S}_{32}$  are distinguished for their outstanding chemical stability with the cathode-stable coating layers. We suggest a half-cell architecture comprising of the sulfide + halide coating, thiophosphate electrolyte, and NCM cathode with overall activation energy for diffusion of less than 0.5 eV through all components, and chemical reactivity no less than -30 meV/atom at all interfaces. This battery design may maintain the high rate capability and capacity of the state-of-art electrolytes and cathodes while preventing interfacial degradation. Additionally, with the discovery of several superior ionic conductors comparable to thiophosphate electrolytes, we discuss the possibility of simplifying the battery architecture by replacing the thiophosphate electrolyte with sulfide-based or halide-based superionic conductors.

## 2. METHODS

### 2.1. Thermodynamic calculations of phase stability and interfacial stability

Thermodynamic stability of materials can be estimated by constructing convex hulls of DFT-calculated formation energies with respect to composition. We have assessed the thermodynamic stability of each compound using energies of all compounds in the chemical space from the Materials Project database.<sup>43</sup> The convex hull was constructed using the pymatgen software package<sup>44</sup>. The difference between the energy of the compound and the energy of the convex hull at the same composition corresponds to the driving force of decomposition, which can be used to quantify the thermodynamic stability of a material.<sup>45</sup> Phases that are on the convex hull are predicted to be stable at 0 K. However, many phases with formation energies above the calculated 0 K convex hull exist in nature for extended periods of time, as a result of either DFT error, entropic stabilization, or phase metastability preserved by a high kinetic barrier of decomposition. Thus we used 30 meV/atom above the convex hull as the pre-screening criterion for phase stability.

Two phases can coexist in local thermodynamic equilibrium if they are connected by a tie-line on the phase diagram. Thermodynamically compatible interfaces are identified using the interfacial reaction energy methodology developed by Richards et al.<sup>46</sup>, in which the reactivity is quantified by determining at which fraction of two phases the reaction driving force becomes maximal. An interfacial reaction energy of zero indicates that the two phases can form a thermodynamically stable interface. The pymatgen software package<sup>44</sup> was used to assess interfacial stability.

## 2.2. Ab initio molecular dynamics

Ab initio molecular dynamic (AIMD) simulations were performed using the Perdew-Burk-Ernzerhof generalized gradient approximation exchange-correlation functional<sup>47</sup> and projector augmented wave<sup>48</sup> potentials as listed in the Supplementary Information Table S1-S2. We used a plane wave energy cut-off of 400 eV and a minimal  $\Gamma$ -centered  $1 \times 1 \times 1$   $k$ -point mesh which has been shown to provide a satisfactory balance between computational accuracy and cost.<sup>49, 50</sup> Computations were performed with spin polarization and with magnetic ions initialized in a high-spin ferromagnetic state for materials containing transition metal atoms. Non-spin-polarized calculations were performed otherwise. A time step of 2 fs was adopted. The supercell sizes were constructed to ensure there were at least 9 Å between neighboring images<sup>49</sup> to avoid periodic boundary effects. The lattice parameter was fixed at that of the cell fully relaxed at 0 K.

For each material, 15 ps AIMD was carried out at 1000 K in the NVT ensemble. In each AIMD run, the first 2 ps was used for equilibration and then a 13 ps dynamic trajectory was generated to compute the mean squared displacement of the Li ions. In the case that the maximal mean-square-displacement of lithium has exceeded 9 Å<sup>2</sup> over 13 ps at elevated temperature, the material was considered to be a potential Li-ion conductor at room temperature and additional molecular dynamics simulations were performed with moment tensor potentials trained on the AIMD trajectory as described in the next section. Otherwise, the material was estimated to be ionically insulating at room temperature and no further calculations were performed to determine its ionic conductivity.

## 2.3. Learning on-the-fly molecular dynamics

The ionic conductivity of each material passing the AIMD screening was computed using a scheme we developed recently<sup>39</sup> based on learning-on-the-fly molecular dynamics (LOTF-MD) and moment tensor potentials (MTPs).<sup>51</sup> The use of this scheme enables us to gather orders of magnitude more data than can be achieved using AIMD with little loss of accuracy in the calculated energies, resulting in greatly reduced statistical variance in the estimated ionic conductivity. For each structure the MTP was initially trained on 13 ps (2 fs per time step) of the AIMD simulation at 1000 K. The molecular dynamics simulations used to evaluate ionic conductivity were performed with the trained MTP in LAMMPS.<sup>52</sup> The interface between moment tensor potential molecular dynamics and DFT geometry optimization was carried out by the Machine Learning of Interatomic Potentials (MLIP) software package.<sup>53</sup> The MTP cutoff radius and the maximum level of basis functions,  $lev_{max}$  were chosen to be 5.0 Å and 10, respectively. Details of the active learning criteria, validation error of energies and forces on various crystalline structures as well as experimental benchmark results can be found in our previous report.<sup>39</sup> In this study, the simulation temperature was initialized at 1000 K for each material and decremented by 50 K until the diffusivity was so low that the total mean squared displacement did not reach 2000 Å<sup>2</sup> within 200 ns. Each MD run lasted for at least 4 ns each until the total mean square displacement of Li<sup>+</sup> reached 2000 Å<sup>2</sup>, which has been reported by He, X. F., et al.<sup>38</sup> to be a sufficient sample size for accurate statistical analysis. To identify the onset of possible melting, we calculated the mean squared displacement of species other than lithium. We determined that melting occurred if the mean squared displacement of non-lithium species exceeded 3 Å<sup>2</sup>. If melting was detected, we decreased the simulation temperature until the structure remained intact during the molecular dynamics run. The lowest five qualifying temperatures were used to fit an Arrhenius relationship to determine the activation energy for diffusion and estimate the room-temperature diffusivity. The room-temperature conductivity  $\sigma_{300K}$  was calculated using the Nernst–Einstein relation under the assumption that the Haven ratio is equal to one:<sup>54, 55</sup>

$$\sigma_{300K} = \frac{ne^2z^2}{k_B T} D_{300K} \quad (1)$$

where  $n$  is the volume density of the diffusing species,  $e$  is the unit electron charge,  $z$  is the charge of the ionic conductor (here 1 for Li<sup>+</sup>), and  $D_{300K}$  is the extrapolated room-temperature tracer diffusivity.

## 2.4 Coating candidates screening criteria

Our screening workflow (Figure 1) was designed to efficiently use computational resources by carrying out relatively fast predictions on a vast pool of candidates, followed by expensive Li conductivity calculations limited to those highly promising compounds satisfying the criteria of high Li content, phase stability, and interfacial stability with either the cathode or electrolyte. Considering all known solid electrolytes contain Li, we searched among the compounds with at least 10% Li by composition. We screened for materials with energies no more than 30 meV/atom above the convex hull. Among those, we calculated the interfacial reaction energy with the solid

electrolyte and with the cathode in both discharged state  $\text{LiNi}_{1/3}\text{Co}_{1/3}\text{Mn}_{1/3}\text{O}_2$  and charged state  $\text{Li}_{1/3}\text{Ni}_{1/3}\text{Co}_{1/3}\text{Mn}_{1/3}\text{O}_2$ .<sup>56</sup> By the criterion of interfacial reaction energy no less than -30 meV/atom, we identified coating candidates compatible with NCM cathode (Table S1) and compatible with the solid electrolytes (Table S2). Among compounds meeting the above requirements, we performed diffusion calculations using LOTF-MD and selected those computed with activation energies less than 0.7 eV to be the coating candidates. This value is lower than that of commonly employed crystalline coating materials such as  $\text{Li}_3\text{PO}_4$  (1.1-1.3 eV),<sup>57, 58</sup>  $\text{Al}_2\text{O}_3$  (1.15 eV),<sup>59</sup> and  $\text{Li}_2\text{CO}_3$  (0.8 eV).<sup>60</sup>

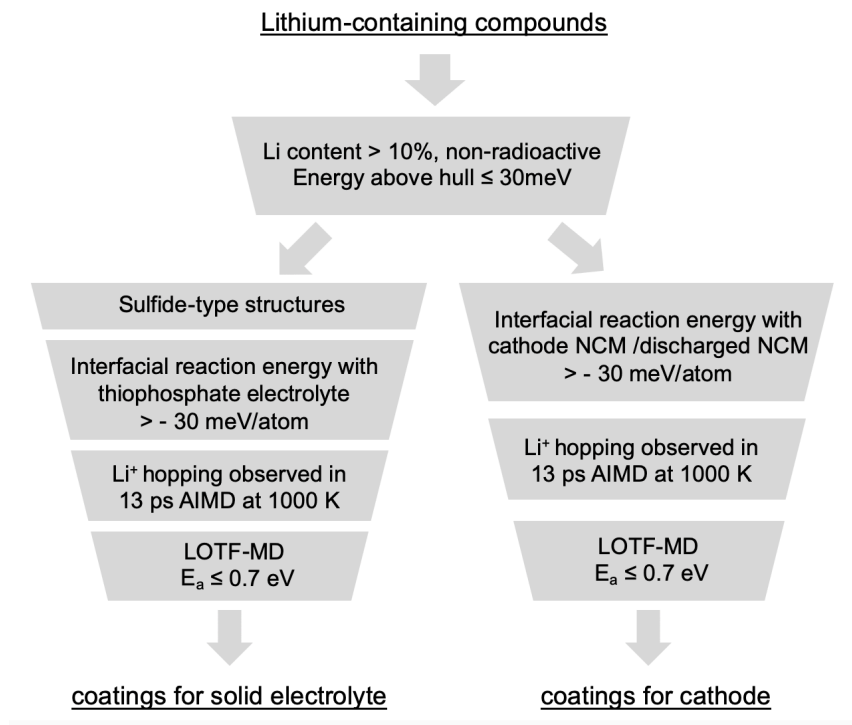


Figure 1. Workflow of high-throughput screening of Li-containing compounds to identify candidate coatings for the lithium thiophosphate electrolyte (left) and NCM cathode (right).

### 3. RESULTS AND DISCUSSION

The initial screening in terms of Li content, phase stability and interfacial stability was conducted rapidly with DFT-calculated energies from the Materials Project database. From 1545 Li-containing crystalline materials, we found 181 materials that are thermodynamically compatible with the cathode in both the charged and discharged states, and 93 materials that are thermodynamically compatible with the solid electrolytes. The same set of 93 materials was stable against both LGPS and  $\text{Li}_6\text{PS}_5\text{Cl}$ . Within the reduced pool of materials, we have performed the two-step molecular dynamics calculation as described in Methods sections 2.2 and 2.3. Based on the Li-ion displacement during 15 ps AIMD simulation at 1000 K, 238 materials were predicted to be ionic insulators at room temperature. We computed the diffusion activation energy of the

remaining 36 materials with LOTF-MD at multiple temperatures and identified 25 materials that meet the criteria of chemical compatibility and sufficiently high Li-ion conductivity.

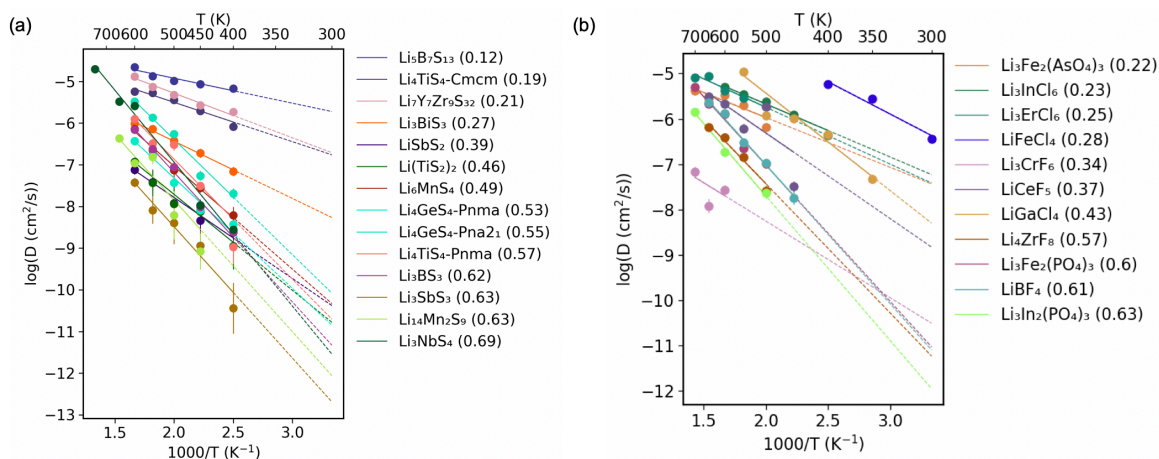


Figure 2. Diffusion coefficients calculated from LOTF-MD for (a) 14 sulfide coating candidates (b) 11 NCM cathode coating candidates. The diffusivity extrapolated to room temperature for each material is shown as a dashed line of the same color. Structures with the same composition are distinguished by space groups in the legend. The activation energies computed from the slope of the fitting line are in the legend (in brackets) in eV.

From 94 pre-selected sulfides, we identified 14 ionic conductors with room-temperature conductivity higher than  $10^{-5}$  mS/cm. Diffusivities computed by LOTF-MD are displayed on the Arrhenius plot in Figure 2a. The activation energy and ionic conductivity extrapolated to room temperature along with the literature reported values for these 14 sulfides are listed in Table 1. Among these candidates are  $\text{Li}_4\text{GeS}_4$ ,  $\text{Li}_5\text{B}_7\text{S}_{13}$ ,  $\text{Li}_3\text{BS}_3$  and  $\text{Li}_3\text{SbS}_3$ , which are known lithium-ion conductors. The first three have been previously suggested in the literature as interface coatings for solid-state batteries.<sup>61, 62</sup> In experiments, improved cycling performance was observed in  $\text{Li}_4\text{SnS}_4$  (isostructural to  $\text{Li}_4\text{GeS}_4$ )-coated  $\text{LiCoO}_2$  compared with pristine  $\text{LiCoO}_2$  in contact with the LGPS electrolyte.<sup>63</sup>

The identified sulfide coating materials effectively extend the electrochemical stability window with respect to phosphorous oxidation (Figure 3a). Despite the passivation of phosphorous, the sulfide coating cannot completely eliminate the reactivity with the oxide cathode. At high electrochemical potentials, the decomposition of thiophosphates involves the oxidation of sulfur, which likewise contributes to the decomposition of sulfide coatings, as shown in Figure 3b. To protect the sulfide layers from oxidation reactions, we searched for a second layer that could stabilize the interface against the oxide-based cathode. Among the 181 compounds stable against the NCM cathode, 11 have diffusion activation energies lower than 0.7 eV as calculated using LOTF-MD. Aside from three polyanionic compounds  $\text{Li}_3\text{X}_2(\text{YO}_4)_3$  ( $\text{X} = \text{Fe}, \text{In}; \text{Y} = \text{P}, \text{As}$ ),



the majority are chlorides and fluorides (Figure 2b and Table 2). Halides have been recently recognized as promising materials for applications in solid-state batteries due to their desirable balance between electrochemical stability and ionic conductivity.<sup>42</sup> We exclude  $\text{Li}_3\text{Fe}_2(\text{AsO}_4)_3$  and  $\text{LiFeCl}_4$  as coating candidates because of the toxicity of the arsenate and the low predicted melting temperature ( $\leq 400$  K) of  $\text{LiFeCl}_4$ .

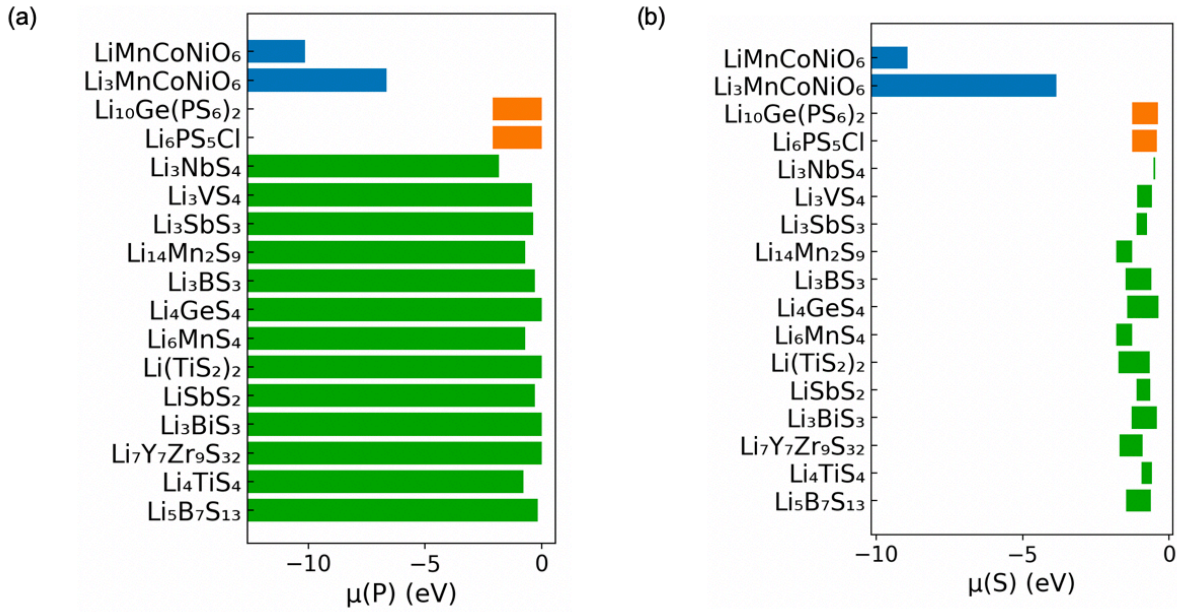


Figure 3. Stability windows with respect to the chemical potential of (a) phosphorous (b) sulfur in NCM (blue), thiophosphate electrolytes (orange), and sulfide-based coating candidates (green).

Table 1. Calculated activation energies for diffusion ( $E_a$ ) in 14 sulfide-based electrolyte coating candidates computed by LOTF-MD. Literature values are listed where available.

Materials Project entry id	Composition	Space group	LOTF-MD-computed $E_a$ (eV)	Extrapolated 300 K conductivity (mS/cm)	Literature reported $\text{Li}^+$ conductivity at 300 K if available
mp-532413	$\text{Li}_5\text{B}_7\text{S}_{13}$	Cc	$0.12 \pm 0.01$	137.924	74 mS/cm <sup>64</sup>
mp-756811	$\text{Li}_4\text{TiS}_4$	Cmcm	$0.19 \pm 0.02$	24.448	
mp-767467	$\text{Li}_7\text{Y}_7\text{Zr}_9\text{S}_{32}$	P1	$0.21 \pm 0.02$	6.504	
mp-753720	$\text{Li}_3\text{BiS}_3$	R-3	$0.27 \pm 0.03$	0.695	
mp-14591	$\text{LiSbS}_2$	R-3	$0.39 \pm 0.07$	0.003	
mp-1045384	$\text{Li}(\text{TiS}_2)_2$	Fd-3m	$0.46 \pm 0.09$	0.0008	
mp-756490	$\text{Li}_6\text{MnS}_4$	$\text{P4}_2/\text{nmc}$	$0.49 \pm 0.04$	0.0092	
mp-30249	$\text{Li}_4\text{GeS}_4$	Pnma	$0.53 \pm 0.03$	0.0020	$2.0 \times 10^{-4}$ mS/cm <sup>65</sup>
mp-1222582	$\text{Li}_4\text{GeS}_4$	Pna2 <sub>1</sub>	$0.55 \pm 0.03$	0.0121	
mp-766540	$\text{Li}_4\text{TiS}_4$	Pnma	$0.57 \pm 0.04$	0.0028	
mp-5614	$\text{Li}_3\text{BS}_3$	Pnma	$0.62 \pm 0.04$	0.0007	0.0031 - 9.7 mS/cm <sup>61</sup>
mp-756198	$\text{Li}_{14}\text{Mn}_2\text{S}_9$	P-3	$0.63 \pm 0.09$	0.0002	
mp-1194339	$\text{Li}_3\text{SbS}_3$	Pna2 <sub>1</sub>	$0.63 \pm 0.1$	$2.73 \times 10^{-5}$	$1.6 \times 10^{-6}$ mS/cm <sup>66</sup>
mp-760375	$\text{Li}_3\text{VS}_4$	P-43m	$0.66 \pm 0.15$	$1.35 \times 10^{-5}$	
mp-755309	$\text{Li}_3\text{NbS}_4$	P-43m	$0.69 \pm 0.03$	0.0002	

Table 2. Calculated activation energies for diffusion ( $E_a$ ) in 11 cathode coating candidates computed by LOTF-MD. Literature values are listed where available.

Materials Project entry id	Composition	Space group	LOTF-MD-computed $E_a$ (eV)	Extrapolated 300 K conductivity (mS/cm)	Literature value of $\text{Li}^+$ conductivity at 300 K or $E_a$ if available
mp-556256	$\text{Li}_3\text{Fe}_2(\text{AsO}_4)_3$	$\text{P2}_1/\text{c}$	$0.22 \pm 0.03$	2.7632	
mp-676109	$\text{Li}_3\text{InCl}_6$	C2	$0.23 \pm 0.02$	4.9162	0.05–0.3 mS/cm <sup>67</sup>
mp-676361	$\text{Li}_3\text{ErCl}_6$	P321	$0.25 \pm 0.02$	3.3038	1.49 mS/cm <sup>68</sup>
mp-1210931	$\text{LiFeCl}_4$	$\text{P2}_1/\text{c}$	$0.28 \pm 0.02$	16.5844	
mp-561396	$\text{Li}_3\text{CrF}_6$	$\text{C2}/\text{c}$	$0.34 \pm 0.13$	0.0053	
mp-1195868	$\text{LiCeF}_5$	$\text{I4}_1/\text{a}$	$0.37 \pm 0.03$	0.0998	
mp-28341	$\text{LiGaCl}_4$	$\text{P2}_1/\text{c}$	$0.43 \pm 0.02$	0.1945	
mp-9308	$\text{Li}_4\text{ZrF}_8$	Pnma	$0.57 \pm 0.06$	0.0011	$E_a = 0.22$ eV <sup>69</sup>
mp-31788	$\text{Li}_3\text{Fe}_2(\text{PO}_4)_3$	$\text{P2}_1/\text{c}$	$0.6 \pm 0.03$	0.00076	$3 \times 10^{-6}$ mS/cm <sup>70</sup>
mp-12403	$\text{LiBF}_4$	$\text{P3}_121$	$0.61 \pm 0.04$	0.00062	
mp-6425	$\text{Li}_3\text{In}_2(\text{PO}_4)_3$	R-3	$0.63 \pm 0.05$	$8.21 \times 10^{-5}$	$E_a = 0.58$ eV <sup>71</sup>

Having identified promising coating candidates for the electrolyte and the cathode, we next evaluated the interfacial stability between pairs of coatings that could be used to create a double-layer coating (Figure 4). The coatings identified by our screen generally reduce the thermodynamic driving force for interfacial reactions compared to direct contact between the cathode material and the electrolyte. The sulfide coatings  $\text{Li}_5\text{B}_7\text{S}_{13}$ ,  $\text{Li}_7\text{Y}_7\text{Zr}_9\text{S}_{32}$  and  $\text{Li}(\text{TiS}_2)_2$  have particularly high compatibility with coating materials for the NCM cathode. We identify eight coating combinations that reduce the maximum magnitude of the interfacial reaction energy from 374 meV with  $\text{Li}_6\text{PS}_5\text{Cl}$  and 360 meV/atom with LGPS to lower than 30 meV/atom. Four of these pairs,  $\text{Li}_5\text{B}_7\text{S}_{13}$ - $\text{Li}_3\text{ErCl}_6$ ,  $\text{Li}(\text{TiS}_2)_2$ - $\text{Li}_3\text{InCl}_6$ ,  $\text{Li}_7\text{Y}_7\text{Zr}_9\text{S}_{32}$ - $\text{Li}_3\text{ErCl}_6$ , and  $\text{Li}(\text{TiS}_2)_2$ - $\text{Li}_3\text{ErCl}_6$  are predicted to have high ionic conductivity, close to those of the solid electrolytes.

Due to the moderate ionic conductivity of commonly employed coatings, deposition thickness at the nanometer scale is often required to achieve competitive rate performance.<sup>35</sup> The coating should also be thick enough to be electronically insulating and protect the electrolyte from the extreme chemical potential of the cathode. Several coating techniques have been investigated to balance these factors and create high-quality coatings in an economically viable way.<sup>24, 72-74</sup> In comparison, some of the double-layer coating designs we propose in this study are predicted to have high ionic conductivity comparable to that of the solid electrolyte. Thus it may be possible to relax the thickness of the coatings to the micrometer scale without significant loss in rate capability, simplifying the production of all-solid-state batteries.

Several candidates from our coating material screen exhibit excellent properties for solid electrolytes. Given the predicted outstanding ionic conductivity and electrochemical stability of  $\text{Li}_5\text{B}_7\text{S}_{13}$  and  $\text{Li}_7\text{Y}_7\text{Zr}_9\text{S}_{32}$ , they are potential alternatives for the best-known thiophosphate-based solid-state electrolyte materials, including  $\text{Li}_{10}\text{GeP}_2\text{S}_{12}$  and  $\text{Li}_6\text{PS}_5\text{Cl}$ . Therefore, along with half cells consisting of thiophosphate SE | sulfide-halide coatings | cathode, we suggest that the phosphorous-free sulfide SE | halide | cathode (sulfide =  $\text{Li}_5\text{B}_7\text{S}_{13}$ ,  $\text{Li}_7\text{Y}_7\text{Zr}_9\text{S}_{32}$ ) may offer comparable or improved performance and simplify the manufacturing process. On the anode side, the sulfide-based electrolyte candidates exhibit interfacial reactivity similar to that with the thiophosphate electrolyte (Table 3), suggesting that if these materials were to be used as electrolytes, anode coatings may be necessary.



## 4. CONCLUSIONS

Computational screening identified specific combinations of sulfides and halides that could yield a double-layer cathode coating with high ionic conductivity and excellent interface compatibility in solid-state batteries. We find the coating pairs  $\text{Li}_5\text{B}_7\text{S}_{13}$ - $\text{Li}_3\text{ErCl}_6$ ,  $\text{Li}(\text{TiS}_2)_2$ - $\text{Li}_3\text{InCl}_6$ ,  $\text{Li}_7\text{Y}_7\text{Zr}_9\text{S}_{32}$ - $\text{Li}_3\text{ErCl}_6$ , and  $\text{Li}(\text{TiS}_2)_2$ - $\text{Li}_3\text{ErCl}_6$  to be particularly promising, as the ionic conductivity could remain close to the value of the original solid electrolyte and there is a low thermodynamic driving force for chemical reactions at all interfaces. Furthermore, the high ionic conductivity and improved electrochemical stability predicted in the two sulfide materials  $\text{Li}_5\text{B}_7\text{S}_{13}$  and  $\text{Li}_7\text{Y}_7\text{Zr}_9\text{S}_{32}$  suggest they may be alternatives to the best-known thiophosphate solid electrolytes.

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# ***ELECTRONIC SUPPLEMENTARY INFORMATION***

## **Computational design of double-layer cathode coatings in all-solid-state batteries**

Chuhong Wang,<sup>1</sup> Koutarou Aoyagi,<sup>1</sup> Tim Mueller<sup>1,\*</sup>

<sup>1</sup> Department of Materials Science and Engineering, Johns Hopkins University, Baltimore,  
Maryland 21218, United States

\* Corresponding Author E-mail: [tmueller@jhu.edu](mailto:tmueller@jhu.edu)

**Table S1. Interfacial reaction energies and ab-initio molecular dynamics (AIMD) parameters of coating candidates compatible with a  $\text{Li}_3\text{MnCoNiO}_6$  (NCM) cathode.**

Materials Project entry id	Composition	Space group	Interfacial reaction energy (meV/atom) with		VASP PAW-PBE potentials
			$\text{Li}_3\text{MnCoNiO}_6$	$\text{LiMnCoNiO}_6$	
mp-1205739	$\text{Li}_2\text{SnF}_6$	P-31m	0	0	Li_sv, Sn_d, F
mp-7791	$\text{Li}_2\text{GeF}_6$	P4_2/mnm	0	0	Li_sv, Ge_d, F
mp-12829	$\text{LiCaGaF}_6$	P-31c	0	0	Li_sv, Ca_sv, Ga_d, F
mp-24199	$\text{LiHF}_2$	R-3m	9	0	Li_sv, H, F
mp-1208619	$\text{SrLiGaF}_6$	P-31c	0	0	Sr_sv, Li_sv, Ga_d, F
mp-6527	$\text{Na}_3\text{Li}_3\text{In}_2\text{F}_{12}$	Ia-3d	0	0	Na_pv, Li_sv, In_d, F
mp-7603	$\text{Li}_2\text{TiF}_6$	P4_2/mnm	0	0	Li_sv, Ti_pv, F
mp-561430	$\text{LiLuF}_4$	C2/c	0	0	Li_sv, Lu_3, F
mp-16577	$\text{Li}_2\text{CaHfF}_8$	I-4	0	0	Li_sv, Ca_sv, Hf_pv, F
mp-1193269	$\text{Li}_2\text{SiF}_6$	P321	0	0	Li_sv, Si, F
mp-29040	$\text{Li}_3\text{Zr}_4\text{F}_{19}$	P-1	0	0	Li_sv, Zr_sv, F
mp-6134	$\text{LiCaAlF}_6$	P-31c	0	0	Li_sv, Ca_sv, Al, F
mp-6591	$\text{SrLiAlF}_6$	P-31c	0	0	Sr_sv, Li_sv, Al, F
mp-14363	$\text{Rb}_2\text{LiAsO}_4$	Cmc2_1	0	0	Rb_sv, Li_sv, As, O
mp-6711	$\text{Na}_3\text{Li}_3\text{Al}_2\text{F}_{12}$	Ia-3d	0	0	Na_pv, Li_sv, Al, F
mp-3700	$\text{LiYF}_4$	I4_1/a	0	0	Li_sv, Y_sv, F
mp-4002	$\text{Li}_2\text{ZrF}_6$	P-31m	0	0	Li_sv, Zr_sv, F
mp-10103	$\text{LiYbAlF}_6$	P-31c	0	0	Li_sv, Yb_2, Al, F
mp-1193222	$\text{LiMgAlF}_6$	P321	0	0	Li_sv, Mg_pv, Al, F
mp-15558	$\text{Li}_3\text{GaF}_6$	C2/c	0	0	Li_sv, Ga_d, F
mp-18704	$\text{CsLiBeF}_4$	P2_1/c	0	0	Cs_sv, Li_sv, Be_sv, F
mp-4622	$\text{Li}_2\text{BeF}_4$	R-3	0	0	Li_sv, Be_sv, F
mp-15254	$\text{Li}_3\text{AlF}_6$	C2/c	0	0	Li_sv, Al, F
mp-1211087	$\text{LiCdFeF}_6$	P-31c	0	0	Li_sv, Cd, Fe_pv, F
mp-14023	$\text{Na}_3\text{Li}_3\text{Sc}_2\text{F}_{12}$	Ia-3d	0	0	Na_pv, Li_sv, Sc_sv, F

mp-543044	BaLiAlF <sub>6</sub>	P2 <sub>1</sub> /c	0	0	Ba_sv, Li_sv, Al, F
mp-557327	Na <sub>3</sub> Li <sub>3</sub> Co <sub>2</sub> F <sub>12</sub>	Ia-3d	-24.4	0	Na_pv, Li_sv, Co, F
mp-1211093	LiCaFeF <sub>6</sub>	P-31c	0	0	Li_sv, Ca_sv, Fe_pv, F
mp-8892	LiInF <sub>4</sub>	Pbcn	0	0	Li_sv, In_d, F
mp-567062	SrLiFeF <sub>6</sub>	P2 <sub>1</sub> /c	0	0	Sr_sv, Li_sv, Fe_pv, F
mp-1195868	LiCeF <sub>5</sub>	I4 <sub>1</sub> /a	0	0	Li_sv, Ce, F
mp-561280	Na <sub>3</sub> Li <sub>3</sub> Fe <sub>2</sub> F <sub>12</sub>	Ia-3d	0	0	Na_pv, Li_sv, Fe_pv, F
mp-1196169	LiThF <sub>5</sub>	I4 <sub>1</sub> /a	0	0	Li_sv, Th, F
mp-7594	CsLiF <sub>2</sub>	C2/c	0	0	Cs_sv, Li_sv, F
mp-9308	Li <sub>4</sub> ZrF <sub>8</sub>	Pnma	0	0	Li_sv, Zr_sv, F
mp-6253	KLiBeF <sub>4</sub>	P6 <sub>3</sub>	0	0	K_sv, Li_sv, Be_sv, F
mp-1138	LiF	Fm-3m	0	0	Li_sv, F
mp-776627	Li <sub>3</sub> FeF <sub>6</sub>	C2/c	0	0	Li_sv, Fe_pv, F
mp-1211435	KLiTbF <sub>5</sub>	P2 <sub>1</sub> /c	0	0	K_sv, Li_sv, Tb_3, F
mp-1211585	KLiGdF <sub>5</sub>	P2 <sub>1</sub> /c	0	0	K_sv, Li_sv, Gd, F
mp-1211494	KLiHoF <sub>5</sub>	P2 <sub>1</sub> /c	0	0	K_sv, Li_sv, Ho_3, F
mp-1211461	KLiDyF <sub>5</sub>	P2 <sub>1</sub> /c	0	0	K_sv, Li_sv, Dy_3, F
mp-1211480	KLiLuF <sub>5</sub>	P2 <sub>1</sub> /c	0	0	K_sv, Li_sv, Lu_3, F
mp-560518	RbLi <sub>2</sub> Be <sub>2</sub> F <sub>7</sub>	P2 <sub>1</sub> /c	0	0	Rb_sv, Li_sv, Be_sv, F
mp-557798	K <sub>5</sub> Li <sub>2</sub> NdF <sub>10</sub>	Pnma	0	0	K_sv, Li_sv, Nd_3, F
mp-1196988	K <sub>5</sub> Li <sub>2</sub> PrF <sub>10</sub>	Pnma	0	0	K_sv, Li_sv, Pr_3, F
mp-1208634	SrLiCrF <sub>6</sub>	P-31c	0	0	Sr_sv, Li_sv, Cr_pv, F
mp-35759	Li <sub>2</sub> NiF <sub>4</sub>	Imma	0	0	Li_sv, Ni_pv, F
mp-1212270	Li <sub>3</sub> Lu(NO <sub>3</sub> ) <sub>6</sub>	P2 <sub>1</sub> /c	0	0	Li_sv, Lu_3, N, O
mp-561330	Na <sub>3</sub> Li <sub>3</sub> Cr <sub>2</sub> F <sub>12</sub>	Ia-3d	0	0	Na_pv, Li_sv, Cr_pv, F
mp-28567	LiBiF <sub>4</sub>	I4 <sub>1</sub> /a	0	0	Li_sv, Bi, F
mp-1211098	LiCrCdF <sub>6</sub>	P-31c	0	0	Li_sv, Cr_pv, Cd, F
mp-565544	BaLiCrF <sub>6</sub>	P2 <sub>1</sub> /c	0	0	Ba_sv, Li_sv, Cr_pv, F
mp-565468	LiCaCrF <sub>6</sub>	P-31c	0	0	Li_sv, Ca_sv, Cr_pv, F

mp-720254	Li <sub>2</sub> H <sub>4</sub> (SO <sub>4</sub> ) <sub>3</sub>	Pccn	-23.2	0	Li_sv, H, S, O
mp-29195	LiPO <sub>3</sub>	P2_1/c	-1	0	Li_sv, P, O
mp-561396	Li <sub>3</sub> CrF <sub>6</sub>	C2/c	0	0	Li_sv, Cr_pv, F
mp-1211213	Li <sub>3</sub> Tb <sub>3</sub> (TeO <sub>6</sub> ) <sub>2</sub>	Ia-3d	0	0	Li_sv, Tb_3, Te, O
mp-558059	LiMnF <sub>4</sub>	P2_1/c	-26.8	0	Li_sv, Mn_pv, F
mp-561011	Li <sub>2</sub> Ta <sub>2</sub> (OF <sub>2</sub> ) <sub>3</sub>	P3_121	0	0	Li_sv, Ta_pv, O, F
mp-1211152	Li <sub>3</sub> Eu <sub>3</sub> (TeO <sub>6</sub> ) <sub>2</sub>	Ia-3d	0	0	Li_sv, Eu, Te, O
mp-559129	Li <sub>3</sub> Er(NO <sub>3</sub> ) <sub>6</sub>	P2_1/c	0	0	Li_sv, Er_3, N, O
mp-1211772	Li <sub>3</sub> Dy(NO <sub>3</sub> ) <sub>6</sub>	P2_1/c	0	0	Li_sv, Dy_3, N, O
mp-1212059	Li <sub>3</sub> Tm(NO <sub>3</sub> ) <sub>6</sub>	P2_1/c	0	0	Li_sv, Tm_3, N, O
mp-1212130	Li <sub>3</sub> Tb(NO <sub>3</sub> ) <sub>6</sub>	P2_1/c	0	0	Li_sv, Tb_3, N, O
mp-1212476	Li <sub>3</sub> Gd(NO <sub>3</sub> ) <sub>6</sub>	P2_1/c	0	0	Li_sv, Gd, N, O
mp-1212171	Li <sub>3</sub> Y(NO <sub>3</sub> ) <sub>6</sub>	P2_1/c	0	0	Li_sv, Y_sv, N, O
mp-6726	CsLiSO <sub>4</sub>	P2_1/c	0	0	Cs_sv, Li_sv, S, O
mp-1212254	Li <sub>3</sub> Yb(NO <sub>3</sub> ) <sub>6</sub>	P2_1/c	0	0	Li_sv, Yb_2, N, O
mp-540946	LiReO <sub>4</sub>	P-1	0	0	Li_sv, Re_pv, O
mp-643458	LiHSO <sub>4</sub>	P2_1/c	-17.6	0	Li_sv, H, S, O
mp-4855	Li <sub>2</sub> SeO <sub>4</sub>	R-3	0	0	Li_sv, Se, O
mp-30301	LiClO <sub>4</sub>	Pnma	0	0	Li_sv, Cl, O
mp-14646	Li <sub>2</sub> Mg <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>	Pbcn	0	0	Li_sv, Mg_pv, S, O
mp-1194281	Li <sub>2</sub> Co(SO <sub>4</sub> ) <sub>2</sub>	P2_1/c	0	0	Li_sv, Co, S, O
mp-8180	LiNO <sub>3</sub>	R-3c	0	0	Li_sv, N, O
mp-4556	Li <sub>2</sub> SO <sub>4</sub>	P2_1/c	0	0	Li_sv, S, O
mp-6412	Li <sub>2</sub> CuP <sub>2</sub> O <sub>7</sub>	C2/c	0	0	Li_sv, Cu_pv, P, O
mp-556229	Li <sub>2</sub> MnF <sub>5</sub>	C2/c	-20.1	0	Li_sv, Mn_pv, F
mp-767932	Li <sub>2</sub> VF <sub>6</sub>	P4_2/mnm	6	0	Li_sv, V_pv, F
mp-24610	LiP(HO <sub>2</sub> ) <sub>2</sub>	Pna2_1	0	0	Li_sv, P, H, O
mp-558902	LiMnGaF <sub>6</sub>	P321	0	0	Li_sv, Mn_pv, Ga_d, F
mp-1194965	Li <sub>2</sub> Th(AsO <sub>4</sub> ) <sub>2</sub>	P2_1/c	0	0	Li_sv, Th, As, O
mp-14484	KNaLi <sub>2</sub> (SO <sub>4</sub> ) <sub>2</sub>	P2_12_12_1	0	0	K_sv, Na_pv, Li_sv, S, O
mp-6800	KLiSO <sub>4</sub>	P31c	0	0	K_sv, Li_sv, S, O
mp-25501	LiCrPO <sub>4</sub> F	P-1	0	0	Li_sv, Cr_pv, P, O, F
mp-9657	LiAsO <sub>3</sub>	R-3	0	0	Li_sv, As, O
mp-6211	RbLiSO <sub>4</sub>	P2_1/c	0	0	Rb_sv, Li_sv, S, O
mp-1195620	LiH <sub>2</sub> ClO <sub>5</sub>	C2/c	0	0	Li_sv, H, Cl, O
mp-560894	Li <sub>4</sub> Be <sub>3</sub> P <sub>3</sub> ClO <sub>12</sub>	P-43n	0	0	Li_sv, Be_sv, P, Cl, O

mp-560072	$\text{Li}_4\text{Be}_3\text{As}_3\text{ClO}_{12}$	P-43n	0	0	Li_sv, Be_sv, As, Cl, O
mp-1193172	$\text{Li}_2\text{Mn}(\text{SO}_4)_2$	P2_1/c	0	0	Li_sv, Mn_pv, S, O
mp-555743	$\text{LiZnPO}_4$	Cc	0	0	Li_sv, Zn, P, O
mp-1196457	$\text{Li}_2\text{B}_3\text{O}_4\text{F}_3$	P2_12_12_1	0	0	Li_sv, B, O, F
mp-25515	$\text{LiFePO}_4\text{F}$	P-1	0	0	Li_sv, Fe_pv, P, O, F
mp-22983	$\text{LiAlCl}_4$	P2_1/c	-21.5	0	Li_sv, Al, Cl
mp-1020705	$\text{Rb}_2\text{Li}_3\text{B}(\text{P}_2\text{O}_7)_2$	Cmce	0	0	Rb_sv, Li_sv, B, P, O
mp-541190	$\text{Cs}_2\text{LiVO}_4$	Cmc2_1	0	0	Cs_sv, Li_sv, V_pv, O
mp-28341	$\text{LiGaCl}_4$	P2_1/c	0	0	Li_sv, Ga_d, Cl
mp-18048	$\text{LiZnAsO}_4$	R3	0	0	Li_sv, Zn, As, O
mp-554560	$\text{Li}_4\text{Be}_3\text{P}_3\text{BrO}_{12}$	P-43n	0	0	Li_sv, Be_sv, P, Br, O
mp-765883	$\text{LiCrPHO}_5$	P-1	0	0	Li_sv, Cr_pv, P, H, O
mp-1019778	$\text{K}_2\text{Li}_3\text{B}(\text{P}_2\text{O}_7)_2$	Cmce	0	0	K_sv, Li_sv, B, P, O
mp-1185263	$\text{LiPaO}_3$	Pm-3m	0	0	Li_sv, Pa, O
mp-555001	$\text{LiMnFeF}_6$	P321	0	0	Li_sv, Mn_pv, Fe_pv, F
mp-25552	$\text{LiMnPO}_4\text{F}$	P-1	-10.8	0	Li_sv, Mn_pv, P, O, F
mp-566629	$\text{Li}_2\text{NiPO}_4\text{F}$	Pnma	0	0	Li_sv, Ni_pv, P, O, F
mp-1222972	$\text{Li}_2\text{TiFe}(\text{PO}_4)_3$	Pna2_1	0	0	Li_sv, Ti_pv, Fe_pv, P, O
mp-565827	$\text{LiMoIO}_6$	P2_1	0	0	Li_sv, Mo_pv, I, O
mp-554577	$\text{Li}_4\text{P}_2\text{O}_7$	P2_1/c	0	0	Li_sv, P, O
mp-556065	$\text{Li}_2\text{MoF}_6$	P4_2/mnm	0	0	Li_sv, Mo_pv, F
mp-1200209	$\text{Li}_2\text{B}_6\text{O}_9\text{F}_2$	Cc	0	0	Li_sv, B, O, F
mp-23626	$\text{RbLi}_2(\text{IO}_3)_3$	P2_1/c	0	0	Rb_sv, Li_sv, I, O
mp-676109	$\text{Li}_3\text{InCl}_6$	C2	0	0	Li_sv, In_d, Cl
mp-1201314	$\text{Li}_8\text{Be}_6\text{P}_7\text{O}_{29}$	P31c	-16.1	0	Li_sv, Be_sv, P, O
mp-23384	$\text{LiIO}_3$	P4_2/n	0	0	Li_sv, I, O
mp-1020646	$\text{NaLi}_2\text{B}(\text{PO}_4)_2$	P-1	0	0	Na_pv, Li_sv, B, P, O
mp-6113	$\text{LiTiAsO}_5$	Pnma	0	0	Li_sv, Ti_pv, As, O
mp-1192681	$\text{LiTcH}_6\text{O}_7$	P6_3mc	0	0	Li_sv, Tc_pv, H, O
mp-25614	$\text{LiNiPO}_4$	Pnma	0	0	Li_sv, Ni_pv, P, O
mp-1222711	$\text{Li}_2\text{GeTeO}_6$	R3	0	0	Li_sv, Ge_d, Te, O

mp-1190687	CsLi <sub>2</sub> Cl <sub>3</sub>	Pbcn	0	0	Cs_sv, Li_sv, Cl
mp-615884	CsLiCrO <sub>4</sub>	P2_1/c	0	0	Cs_sv, Li_sv, Cr_pv, O
mp-771112	Li <sub>2</sub> Mn <sub>3</sub> NiO <sub>8</sub>	P4_332	0	0	Li_sv, Mn_pv, Ni_pv, O
mp-676361	Li <sub>3</sub> ErCl <sub>6</sub>	P321	0	0	Li_sv, Er_3, Cl
mp-1222849	Li <sub>2</sub> ScP <sub>2</sub> HO <sub>8</sub>	P2_1	0	0	Li_sv, Sc_sv, P, H, O
mp-9144	LiAsF <sub>6</sub>	R-3	-17	0	Li_sv, As, F
mp-1080679	LiTaF <sub>6</sub>	R-3	0	0	Li_sv, Ta_pv, F
mp-3980	LiSbF <sub>6</sub>	R-3	0	0	Li_sv, Sb, F
mp-12403	LiBF <sub>4</sub>	P3_121	0	0	Li_sv, B, F
mp-1078799	LiNbF <sub>6</sub>	R-3	0	0	Li_sv, Nb_pv, F
mp-1185319	LiCl	P6_3mc	0	0	Li_sv, Cl
mp-504360	Li <sub>9</sub> Cr <sub>3</sub> P <sub>8</sub> O <sub>29</sub>	P-3c1	0	0	Li_sv, Cr_pv, P, O
mp-1222810	Li <sub>2</sub> InP <sub>2</sub> HO <sub>8</sub>	P2_1	0	0	Li_sv, In_d, P, H, O
mp-24920	Li <sub>2</sub> CrO <sub>4</sub>	R-3	0	0	Li_sv, Cr_pv, O
mp-1223017	Li <sub>2</sub> TiCr(PO <sub>4</sub> ) <sub>3</sub>	Pna2_1	0	0	Li_sv, Ti_pv, Cr_pv, P, O
mp-771864	LiFePHO <sub>5</sub>	P-1	0	0	Li_sv, Fe_pv, P, H, O
mp-1210931	LiFeCl <sub>4</sub>	P2_1/c	0	0	Li_sv, Fe_pv, Cl
mp-22694	LiPPbO <sub>4</sub>	Pna2_1	0	0	Li_sv, P, Pb_d, O
mp-1020015	Li <sub>2</sub> B <sub>3</sub> PO <sub>8</sub>	P-1	0	0	Li_sv, B, P, O
mp-557177	Li <sub>2</sub> Al(BO <sub>2</sub> ) <sub>5</sub>	P2_1/c	0	0	Li_sv, Al, B, O
mp-13725	Li <sub>3</sub> PO <sub>4</sub>	Pmn2_1	0	0	Li_sv, P, O
mp-16828	Li <sub>3</sub> B <sub>7</sub> O <sub>12</sub>	P-1	0	0	Li_sv, B, O
mp-31788	Li <sub>3</sub> Fe <sub>2</sub> (PO <sub>4</sub> ) <sub>3</sub>	P2_1/c	0	0	Li_sv, Fe_pv, P, O
mp-6565	Li <sub>3</sub> Sc <sub>2</sub> (PO <sub>4</sub> ) <sub>3</sub>	P2_1/c	0	0	Li_sv, Sc_sv, P, O
mp-767473	Li <sub>2</sub> FeP <sub>2</sub> HO <sub>8</sub>	P2_1	0	0	Li_sv, Fe_pv, P, H, O
mp-9625	LiMgPO <sub>4</sub>	Pnma	0	0	Li_sv, Mg_pv, P, O
mp-560104	LiTaGeO <sub>5</sub>	P2_1/c	0	0	Li_sv, Ta_pv, Ge_d, O
mp-8673	Li <sub>2</sub> SnTeO <sub>6</sub>	Pnn2	0	0	Li_sv, Sn_d, Te, O
mp-565208	Li <sub>2</sub> U(MoO <sub>5</sub> ) <sub>2</sub>	P-1	0	0	Li_sv, U, Mo_pv, O
mp-863863	LiCoPO <sub>4</sub>	Cc	0	0	Li_sv, Co, P, O
mp-8873	LiGeBO <sub>4</sub>	I-4	0	0	Li_sv, Ge_d, B, O
mp-756117	Li <sub>2</sub> TiTeO <sub>6</sub>	Pnn2	0	0	Li_sv, Ti_pv, Te, O

mp-585492	Li <sub>2</sub> MgMn <sub>3</sub> O <sub>8</sub>	P4_332	0	0	Li_sv, Mg_pv, Mn_pv, O
mp-13843	Li <sub>2</sub> TeO <sub>4</sub>	P4_122	0	0	Li_sv, Te, O
mp-1208625	SrLiVF <sub>6</sub>	P-31c	0	0	Sr_sv, Li_sv, V_pv, F
mp-6425	Li <sub>3</sub> In <sub>2</sub> (PO <sub>4</sub> ) <sub>3</sub>	R-3	0	0	Li_sv, In_d, P, O
mp-6668	LiTiPO <sub>5</sub>	Pnma	0	0	Li_sv, Ti_pv, P, O
mp-1222525	LiAlGeO <sub>4</sub>	R3	0	0	Li_sv, Al, Ge_d, O
mp-772468	Li <sub>2</sub> Cr <sub>3</sub> (CoO <sub>6</sub> ) <sub>2</sub>	Pbcn	-18.2	0	Li_sv, Cr_pv, Co, O
mp-704943	Li <sub>2</sub> Mn <sub>3</sub> ZnO <sub>8</sub>	P2_13	0	0	Li_sv, Mn_pv, Zn, O
mp-770932	LiSbO <sub>3</sub>	C2/m	0	0	Li_sv, Sb, O
mp-761940	Li <sub>2</sub> MnCo <sub>3</sub> O <sub>8</sub>	C2	0	0	Li_sv, Mn_pv, Co, O
mp-556256	Li <sub>3</sub> Fe <sub>2</sub> (AsO <sub>4</sub> ) <sub>3</sub>	P2_1/c	0	0	Li_sv, Fe_pv, As, O
mp-9197	Li <sub>3</sub> AsO <sub>4</sub>	Pmn2_1	0	0	Li_sv, As, O
mp-32316	Li <sub>2</sub> NiGe <sub>3</sub> O <sub>8</sub>	P4_332	0	0	Li_sv, Ni_pv, Ge_d, O
mp-25080	Li <sub>2</sub> MoO <sub>4</sub>	R-3	0	0	Li_sv, Mo_pv, O
mp-23985	LiH <sub>6</sub> BrO <sub>7</sub>	P6_3mc	0	0	Li_sv, H, Br, O
mp-560297	LiTaSiO <sub>5</sub>	P2_1/c	0	0	Li_sv, Ta_pv, Si, O
mp-3054	Li <sub>2</sub> CO <sub>3</sub>	C2/c	0	0	Li_sv, C, O
mp-1222477	LiVZnO <sub>4</sub>	R3	0	0	Li_sv, V_pv, Zn, O
mp-774082	Li(CoO <sub>2</sub> ) <sub>2</sub>	P2_1	0	0	Li_sv, Co, O
mp-19440	LiVO <sub>3</sub>	C2/c	0	0	Li_sv, V_pv, O
mp-18220	LiAlSiO <sub>4</sub>	R3	0	0	Li_sv, Al, Si, O
mp-562137	KLiCO <sub>3</sub>	P2_1/c	0	0	K_sv, Li_sv, C, O
mp-557852	Na <sub>3</sub> Li <sub>3</sub> V <sub>2</sub> F <sub>12</sub>	Ia-3d	0	0	Na_pv, Li_sv, V_pv, F
mp-14364	Cs <sub>2</sub> LiAsO <sub>4</sub>	Cmc2_1	0	0	Cs_sv, Li_sv, As, O
mp-559533	NaLiCO <sub>3</sub>	P-6	0	0	Na_pv, Li_sv, C, O
mp-566323	RbLiMoO <sub>4</sub>	Cc	0	0	Rb_sv, Li_sv, Mo_pv, O
mp-762045	Li <sub>2</sub> Co <sub>3</sub> SbO <sub>8</sub>	P4_332	0	0	Li_sv, Co, Sb, O
mp-561689	CsLiMoO <sub>4</sub>	F-43m	0	0	Cs_sv, Li_sv, Mo_pv, O
mp-18147	LiGaSiO <sub>4</sub>	R3	0	0	Li_sv, Ga_d, Si, O
mp-1210866	LiVCdF <sub>6</sub>	P-31c	0	0	Li_sv, V_pv, Cd, F
mp-1211223	Li <sub>3</sub> Yb <sub>3</sub> (TeO <sub>6</sub> ) <sub>2</sub>	Ia-3d	0	0	Li_sv, Yb_2, Te, O

mp-691115	$\text{Li}_4\text{Mn}_5\text{O}_{12}$	C2/c	0	0	Li_sv, Mn_pv, O
mp-31706	$\text{Li}_3\text{Cr}_2(\text{PO}_4)_3$	R-3	0	0	Li_sv, Cr_pv, P, O
mp-18741	$\text{RbLiCrO}_4$	P31c	0	0	Rb_sv, Li_sv, Cr_pv, O
mp-1178391	$\text{CsLiWO}_4$	I-4	0	0	Cs_sv, Li_sv, W_pv, O
mp-566105	$\text{KLiCrO}_4$	Pna2_1	0	0	K_sv, Li_sv, Cr_pv, O
mp-558045	$\text{NaLi}_2\text{PO}_4$	Pnma	0	0	Na_pv, Li_sv, P, O
mp-9066	$\text{NaLi}_2\text{AsO}_4$	Pmn2_1	0	0	Na_pv, Li_sv, As, O
mp-1199453	$\text{KLiMoO}_4$	P6_3	0	0	K_sv, Li_sv, Mo_pv, O
mp-1176640	$\text{LiMnCrO}_4$	Imma	0	0	Li_sv, Mn_pv, Cr_pv, O
mp-1211142	$\text{Li}_5\text{IO}_6$	P3_112	0	0	Li_sv, I, O



**Table S2. Interfacial reaction energies and AIMD parameters of coating candidates compatible with thiophosphate solid electrolytes.**

Materials Project entry id	Composition	Space group	Interfacial reaction energy (meV/atom) with		VASP PAW-PBE potentials
			Li <sub>6</sub> PS <sub>5</sub> Cl	Li <sub>10</sub> GeP <sub>2</sub> S <sub>12</sub>	
mp-760415	Li <sub>3</sub> SbS <sub>4</sub>	I-42m	0	0	Li_sv, Sb, S
mp-755664	Li <sub>2</sub> (TaS <sub>2</sub> ) <sub>3</sub>	P6 <sub>3</sub> 22	0	0	Li_sv, Ta_pv, S
mp-1045384	Li(TiS <sub>2</sub> ) <sub>2</sub>	Fd-3m	0	0	Li_sv, Ti_pv, S
mp-1192995	Li <sub>4</sub> MnGe <sub>2</sub> S <sub>7</sub>	Cc	-6.2	0	Li_sv, Mn_pv, Ge_d, S
mp-1222404	LiBiS <sub>2</sub>	R-3m	0	0	Li_sv, Bi, S
mp-29410	Li <sub>2</sub> B <sub>2</sub> S <sub>5</sub>	Cmcm	0	0	Li_sv, B, S
mp-19755	Li(TiS <sub>2</sub> ) <sub>3</sub>	P-3m1	-3.9	0	Li_sv, Ti_pv, S
mp-557962	SrLiBS <sub>3</sub>	Pnma	-0.4	0	Sr_sv, Li_sv, B, S
mp-1045435	Li(MnS <sub>2</sub> ) <sub>2</sub>	Imma	0	0	Li_sv, Mn_pv, S
mp-1222722	Li(ZrS <sub>2</sub> ) <sub>2</sub>	P2/m	0	0	Li_sv, Zr_sv, S
mp-853967	LiCuS	Fddd	-3.4	-2.9	Li_sv, Cu_pv, S
mp-1045432	Li(CrS <sub>2</sub> ) <sub>2</sub>	Imma	-16.6	-5.8	Li_sv, Cr_pv, S
mp-774459	Li <sub>8</sub> Ti <sub>16</sub> CuS <sub>32</sub>	P1	-1.5	-1.6	Li_sv, Ti_pv, Cu_pv, S
mp-1222482	Li <sub>6</sub> AsS <sub>5</sub> I	P1	0	0	Li_sv, As, S, I
mp-644419	LiHS	Pmc2 <sub>1</sub>	0	0	Li_sv, H, S
mp-1188784	Li <sub>2</sub> CdSnS <sub>4</sub>	Pmn2 <sub>1</sub>	0	0	Li_sv, Cd, Sn_d, S
mp-753720	Li <sub>3</sub> BiS <sub>3</sub>	R-3	0	0	Li_sv, Bi, S
mp-3647	LiGaS <sub>2</sub>	Pna2 <sub>1</sub>	0	0	Li_sv, Ga_d, S
mp-573030	Cs <sub>2</sub> LiNbS <sub>4</sub>	P-1	0	0	Cs_sv, Li_sv, Nb_pv, S
mp-1222735	Li <sub>2</sub> VCrS <sub>4</sub>	C2/m	0	0	Li_sv, V_pv, Cr_pv, S
mp-767171	Li <sub>5</sub> (NbS <sub>2</sub> ) <sub>7</sub>	C2/c	0	0	Li_sv, Nb_pv, S
mp-755309	Li <sub>3</sub> NbS <sub>4</sub>	P-43m	-0.1	-0.1	Li_sv, Nb_pv, S
mp-754856	Li <sub>7</sub> Y <sub>7</sub> ZrS <sub>16</sub>	P2/m	0	0	Li_sv, Y_sv, Zr_sv, S
mp-1225894	CsLiMnS <sub>2</sub>	I-4m2	-14	-3.5	Cs_sv, Li_sv, Mn_pv, S
mp-28471	Li <sub>3</sub> AsS <sub>3</sub>	Pna2 <sub>1</sub>	0	0	Li_sv, As, S
mp-559814	Li <sub>5</sub> SbS <sub>3</sub> I <sub>2</sub>	Pnnm	0	0	Li_sv, Sb, S, I
mp-1190364	Li <sub>2</sub> SnS <sub>3</sub>	C2/c	-8.4	0	Li_sv, Sn_d, S
mp-767218	Li <sub>9</sub> (NbS <sub>2</sub> ) <sub>14</sub>	P-1	-0.9	0	Li_sv, Nb_pv, S
mp-557892	BaLi(BS <sub>2</sub> ) <sub>3</sub>	Cc	-1.2	0	Ba_sv, Li_sv, B, S
mp-644271	LiHS	Ama2	0	0	Li_sv, H, S
mp-1105291	Li <sub>2</sub> MnSnS <sub>4</sub>	Pc	0	0	Li_sv, Mn_pv, Sn_d, S
mp-766506	Li <sub>3</sub> CuS <sub>2</sub>	Ia-3	0	0	Li_sv, Cu_pv, S
mp-1211176	Li <sub>6</sub> AsS <sub>5</sub> I	Cc	0	0	Li_sv, As, S, I
mp-1079885	LiSbS <sub>2</sub>	C2/c	-0.2	0	Li_sv, Sb, S

mp-753946	LiCoS <sub>2</sub>	P-3m1	0	0	Li_sv, Co, S
mp-554076	BaLiBS <sub>3</sub>	P2_1/c	-1.9	-2	Ba_sv, Li_sv, B, S
mp-767165	K <sub>2</sub> LiVS <sub>4</sub>	Cc	-2.2	0	K_sv, Li_sv, V_pv, S
mp-1210804	Li <sub>2</sub> ZnGeS <sub>4</sub>	Pmn2_1	-3.9	0	Li_sv, Zn, Ge_d, S
mp-1222582	Li <sub>4</sub> GeS <sub>4</sub>	Pna2_1	0	0	Li_sv, Ge_d, S
mp-1191903	Li <sub>2</sub> In <sub>2</sub> SiS <sub>6</sub>	Cc	-12.4	-0.9	Li_sv, In_d, Si, S
mp-1206881	LiTaS <sub>2</sub>	P6_3/mmc	-13.6	-13.7	Li_sv, Ta_pv, S
mp-558815	Rb <sub>2</sub> LiVS <sub>4</sub>	Fddd	0	0	Rb_sv, Li_sv, V_pv, S
mp-1021497	Li <sub>2</sub> SiSnS <sub>4</sub>	I-42m	0	0	Li_sv, Si, Sn_d, S
mp-1195603	Li <sub>2</sub> MnSnS <sub>4</sub>	Pna2_1	0	0	Li_sv, Mn_pv, Sn_d, S
mp-5614	Li <sub>3</sub> BS <sub>3</sub>	Pnma	0	0	Li_sv, B, S
mp-1194339	Li <sub>3</sub> SbS <sub>3</sub>	Pna2_1	0	0	Li_sv, Sb, S
mp-756490	Li <sub>6</sub> MnS <sub>4</sub>	P4_2/nmc	-9.2	-9.2	Li_sv, Mn_pv, S
mp-1176780	LiCoS <sub>2</sub>	P2/c	0	0	Li_sv, Co, S
mp-14591	LiSbS <sub>2</sub>	R-3	-0.2	0	Li_sv, Sb, S
mp-554395	Na <sub>5</sub> Li <sub>3</sub> (TiS <sub>4</sub> ) <sub>2</sub>	C2/c	-5.5	0	Na_pv, Li_sv, Ti_pv, S
mp-33526	LiBiS <sub>2</sub>	I4_1/amd	0	0	Li_sv, Bi, S
mp-19896	Li <sub>2</sub> GePbS <sub>4</sub>	I-42m	-5.6	0	Li_sv, Ge_d, Pb_d, S
mp-558219	SrLi(BS <sub>2</sub> ) <sub>3</sub>	Cc	-2.9	-0.4	Sr_sv, Li_sv, B, S
mp-1188392	LiInS <sub>2</sub>	Pna2_1	0	0	Li_sv, In_d, S
mp-1189383	Li <sub>2</sub> CdGeS <sub>4</sub>	Pmn2_1	0	0	Li_sv, Cd, Ge_d, S
mp-532413	Li <sub>5</sub> B <sub>7</sub> S <sub>13</sub>	Cc	-1.2	0	Li_sv, B, S
mp-510338	Cs <sub>2</sub> LiVS <sub>4</sub>	Fddd	0	0	Cs_sv, Li_sv, V_pv, S
mp-7543	LiVS <sub>2</sub>	P-3m1	0	0	Li_sv, V_pv, S
mp-556085	Rb <sub>2</sub> LiTaS <sub>4</sub>	P-1	0	0	Rb_sv, Li_sv, Ta_pv, S
mp-553962	Rb <sub>2</sub> LiNbS <sub>4</sub>	P-1	0	0	Rb_sv, Li_sv, Nb_pv, S
mp-29829	LiAuS	Fddd	-3.8	0	Li_sv, Au, S
mp-555186	Li <sub>2</sub> ZnSnS <sub>4</sub>	Pc	0	0	Li_sv, Zn, Sn_d, S
mp-558731	Li <sub>2</sub> TeS <sub>3</sub>	P2_1/c	0	0	Li_sv, Te, S
mp-7936	LiNbS <sub>2</sub>	P6_3/mmc	-0.2	-0.2	Li_sv, Nb_pv, S
mp-1001786	LiScS <sub>2</sub>	R-3m	0	0	Li_sv, Sc_sv, S
mp-1001784	LiTiS <sub>2</sub>	R-3m	-5.8	-5.9	Li_sv, Ti_pv, S
mp-756006	LiSbS	P1	0	0	Li_sv, Sb, S
mp-559238	CsLi <sub>2</sub> BS <sub>3</sub>	Pnma	0	0	Cs_sv, Li_sv, B, S
mp-1211182	LiCr <sub>4</sub> InS <sub>8</sub>	F-43m	-18.3	-4.6	Li_sv, Cr_pv, In_d, S
mp-767137	Li <sub>3</sub> Sb <sub>17</sub> S <sub>27</sub>	P1	-5.5	0	Li_sv, Sb, S
mp-15999	Li <sub>3</sub> AuS <sub>2</sub>	Ibam	0	0	Li_sv, Au, S
mp-1153	Li <sub>2</sub> S	Fm-3m	0	0	Li_sv, S

mp-4226	LiCrS <sub>2</sub>	P-3m1	0	0	Li_sv, Cr_pv, S
mp-766540	Li <sub>4</sub> TiS <sub>4</sub>	Pnma	-0.6	0	Li_sv, Ti_pv, S
mp-30959	LiHS	P4_2/mbc	0	0	Li_sv, H, S
mp-769032	Li <sub>3</sub> NbS <sub>4</sub>	I-43m	-0.1	-0.1	Li_sv, Nb_pv, S
mp-1106183	LiAlS <sub>2</sub>	Pna2_1	0	0	Li_sv, Al, S
mp-756316	Li <sub>3</sub> SbS <sub>4</sub>	Pmn2_1	0	0	Li_sv, Sb, S
mp-1191476	Li <sub>2</sub> In <sub>2</sub> GeS <sub>6</sub>	Cc	-15	0	Li_sv, In_d, Ge_d, S
mp-767088	Li <sub>3</sub> Sb <sub>11</sub> S <sub>18</sub>	P1	-5.2	0	Li_sv, Sb, S
mp-555874	LiAsS <sub>2</sub>	Cc	-3.4	0	Li_sv, As, S
mp-861182	Li <sub>4</sub> TiS <sub>4</sub>	I-42m	-0.6	0	Li_sv, Ti_pv, S
mp-995393	LiS <sub>4</sub>	P1	0	0	Li_sv, S
mp-760375	Li <sub>3</sub> VS <sub>4</sub>	P-43m	0	0	Li_sv, V_pv, S
mp-753546	Li <sub>8</sub> TiS <sub>6</sub>	P6_3cm	0	0	Li_sv, Ti_pv, S
mp-1177695	Li <sub>3</sub> CuS <sub>2</sub>	Ibam	0	0	Li_sv, Cu_pv, S
mp-1225887	CsLiZnS <sub>2</sub>	I-4m2	-12.3	-1.5	Cs_sv, Li_sv, Zn, S
mp-756198	Li <sub>14</sub> Mn <sub>2</sub> S <sub>9</sub>	P-3	-8.3	-8.3	Li_sv, Mn_pv, S
mp-1220728	NaLi(AsS <sub>2</sub> ) <sub>2</sub>	P1	-3	0	Na_pv, Li_sv, As, S
mp-756811	Li <sub>4</sub> TiS <sub>4</sub>	Cmcm	-0.6	0	Li_sv, Ti_pv, S
mp-33363	LiYS <sub>2</sub>	I4_1/amd	0	0	Li_sv, Y_sv, S
mp-767467	Li <sub>7</sub> Y <sub>7</sub> Zr <sub>9</sub> S <sub>32</sub>	P1	0	0	Li_sv, Y_sv, Zr_sv, S
mp-15788	LiYS <sub>2</sub>	R-3m	0	0	Li_sv, Y_sv, S