

Pd-Catalyzed C–N Coupling Reactions Facilitated by Organic Bases: Mechanistic Investigation Leads to Enhanced Reactivity in the Arylation of Weakly Binding Amines

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ABSTRACT: The ability to use soluble organic amine bases in Pd-catalyzed C–N cross-coupling reactions has provided a long-awaited solution to the many issues associated with employing traditional, heterogeneous reaction conditions. However, little is known about the precise function of these bases in the catalytic cycle and about the effect of variations in base structure on catalyst reactivity. We used ^{19}F NMR to analyze the kinetic behavior of C–N coupling reactions facilitated by different organic bases. In the case of aniline coupling reactions employing DBU, the resting state was a DBU-bound oxidative addition complex, $\text{LPd}(\text{DBU})(\text{Ar})\text{X}$, and the reaction was found to be inhibited by base. In general, however, depending on the binding properties of the chosen organic base, increased concentration of the base can have a positive or negative influence on the reaction rate. We propose a model in which the turnover-limiting step of the catalytic cycle depends on the relative nucleophilicity of the base compared to that of the amine. This hypothesis guided the discovery of new reaction conditions for the coupling of weakly binding amines, including secondary aryl amines, which were unreactive nucleophiles in our original protocol.

INTRODUCTION

The capability to study and understand organometallic transformations through mechanistic analysis has enabled the development of efficient, practical, and robust synthetic methodologies.^{1–2} In particular, investigation of palladium-catalyzed carbon-nitrogen (C–N) bond forming reactions³ has produced more active catalysts and reliable protocols⁴ that have been widely applied in both academic and industrial settings.⁵ Many of these studies have focused on systems reliant on partially or almost completely insoluble inorganic bases, including metal alkoxide and carbonate bases, to facilitate the deprotonation step (Figure 1). As a result of this heterogeneity, it is difficult to accurately determine the base concentration and to quantify its effect on the reaction rate.^{6–7} In contrast, recently developed homogeneous systems employing soluble organic bases, including phosphazines, amidines, and guanidines,⁸ feature single-phase reaction mixtures that are potentially amenable to precise quantification of dissolved base. Specifically, our laboratory recently reported the use of the soluble organic base 1,8-diazabicyclo[5.4.0]undec-7-ene (DBU) in the coupling of

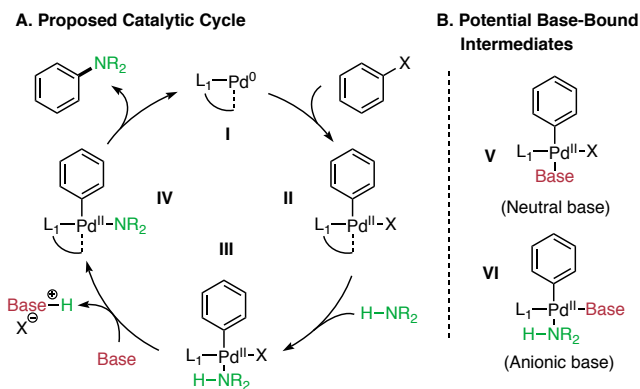


Figure 1. (A) The general catalytic cycle for the Pd-catalyzed coupling of aryl (pseudo)halides with amines: I, monoligated Pd(0) catalyst; II, oxidative addition (OA) complex; III, amine-bound OA complex; IV, amido complex. (B) Potential intermediates involving coordination of a base: V, base-bound OA complex; VI, base-bound/amine-bound OA complex.

aryl amines, amides, and primary aliphatic amines with a variety of aryl (pseudo)halides.⁹ While the soluble organic bases efficiently facilitate the desired C–N

coupling reaction, little is known about their mechanistic role at the molecular level.

The incorporation of nitrogen-containing bases in Pd-catalyzed cross-coupling reactions introduces many potential mechanisms for these Lewis-basic species to influence catalytic intermediates, depending on the inherent nucleophilicity and steric properties of these compounds.^{10, 11} Previously, in alternative catalytic systems, the potential non-innocent nature of bases in C–N coupling has been proposed. For example, in a computational study of the coupling of bromobenzene and morpholine, Norrby and coworkers showed that DBU is able to stabilize catalytic intermediates (Figure 1, VI) by binding to the Pd center.¹² Mechanistic studies by Hartwig have uncovered relationships between the identity of differently substituted alkoxide bases, their concentration, and their effect on overall reaction rate in the coupling of aryl chlorides with *N*-methylbenzylamine.¹³ The ability to understand the effect that amine bases have on C–N cross-coupling would provide valuable fundamental mechanistic information that would also serve as a practical guide for reaction optimizations and the rational determination of which base to employ. Herein, we report the first experimental mechanistic investigations into the effect of organic amine bases in Pd-catalyzed C–N coupling. Based on these findings, we are able to overcome certain limitations of our previously reported protocol by developing modified conditions that permit the coupling of secondary aryl amines with aryl triflates, a transformation that has not yet been facilitated by a weak, organic amine base.¹⁴

RESULTS AND DISCUSSION

To probe the effect of organic bases in coupling reactions facilitated by an AlPhos-supported palladium catalyst, ¹⁹F NMR spectroscopy was used to monitor the rate of cross-coupling between *p*-tolyl triflate (1) with 4-fluoroaniline (2) (Figure 2). The kinetic orders in [Pd], [DBU], [1], and [2] were determined by measuring the rate of formation of 3 over time in THF solutions. As the concentration of 2 was varied between 0.5 and 1.5 M at room temperature, the rate of product formation increased. A plot of ln(Rate) vs ln([2]) fit well with a line with slope +1.02 (Figure 2, A), a result that is consistent with a positive first order dependence on 2. In contrast, when the concentration of DBU was varied, the rate of the reaction decreased. The plot of ln(Rate) vs ln([DBU]) fit a line with slope -1.16, indicating that there was a negative first order dependence on DBU (Figure 2, B). These results suggest that both the amine and the base are involved prior to, or during, the rate-limiting step(s) (RLS) of the reaction. Varying the concentration of 1 did not influence the reaction rate (zeroth order), and

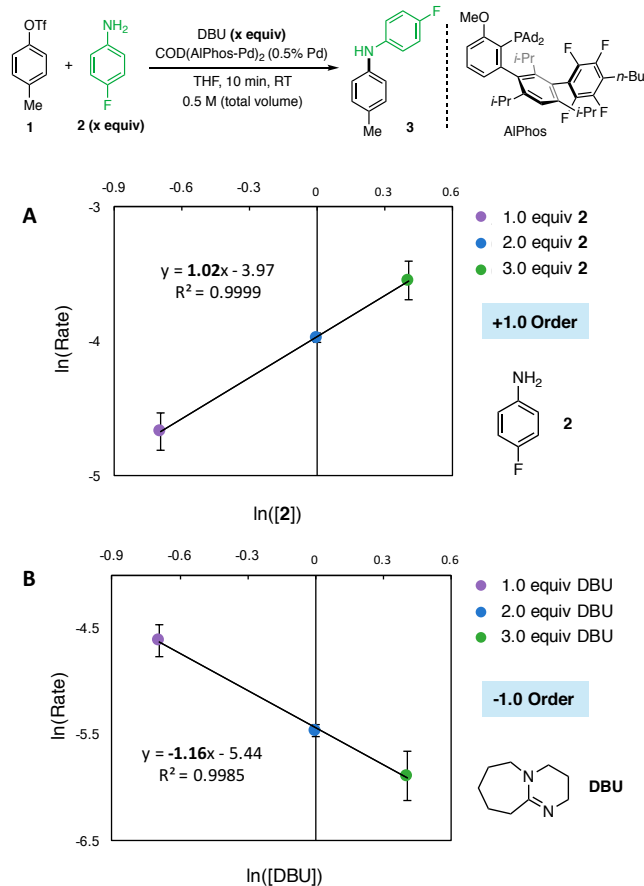


Figure 2. Orders in reactants for the coupling of aryl triflate (1) with aniline (2) in the presence of DBU. Reaction rates were measured in triplicate by ¹⁹F NMR and represent six time points per reaction (18 time points total). The average rate is plotted for reactions performed at 0.5, 1.0, and 1.5 M relative to the varied reagent. Error bars represent the logarithmic standard deviation of the three reaction rates. (A) Logarithmic plot showing a first order dependence on [2]. (B) Logarithmic plot showing a negative first order dependence on [DBU].

positive, linear effect (first order dependence) was observed when increasing the total [Pd] from 0.5 to 1.5 mol% Pd (see the Supporting Information for kinetic data and a description of the studies carried out). As we expected, these results are consistent with a monomeric active Pd species. Moreover, because there is no dependence on [aryl triflate], we believe that the oxidative addition (Figure 1, II) sequence is facile under these reaction conditions. These kinetic data show that the base inhibits (negative order) the coupling reaction of aryl triflates and aryl amines, suggesting that this reagent does more than facilitate the removal of a proton from a palladium-bound amine intermediate or act as a simple acid scavenger.

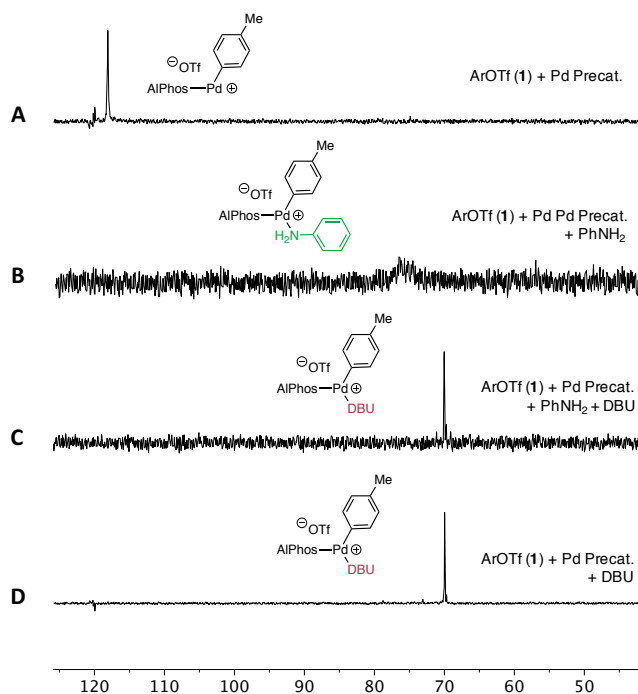


Figure 3. Analysis of reaction intermediates by ^{31}P NMR: (A) OA complex. (B) aniline-bound OA complex. (C) DBU-bound OA complex resting state. (D) DBU-bound OA complex. Pd Precat. denotes $\text{COD}(\text{Pd}(\text{AlPhos})_2)$.

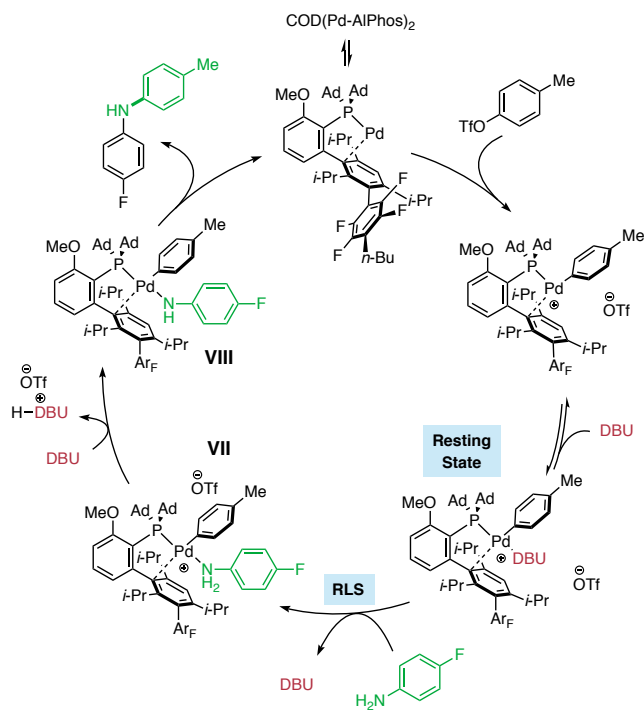
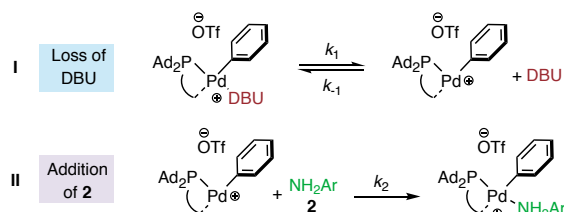


Figure 4. Proposed catalytic cycle of the coupling of aryl triflates and aryl amines in the presence of DBU.

To interpret the effect of [2] and [DBU] on the reaction rate, we used ^{31}P NMR to determine the likely resting state of the catalytic cycle. First, the relative binding affinity of Pd for 2 and DBU was studied via sequential addition of each reaction component to an NMR tube. The addition of a catalytic amount (4 mol% Pd) of precatalyst ($\text{COD}(\text{Pd}(\text{AlPhos})_2)$) to a solution of 1 in THF (Figure 3, A) showed only one signal (118.1 ppm) consistent with the oxidative addition (OA) complex (Figure 1, II), which has been previously isolated and reported.⁹ Next, addition of aniline to this solution resulted in the complete consumption of the OA complex and the formation of a species exhibiting a broad resonance at 77 ppm (Figure 3, B), which we believe to be the aniline-bound OA complex (Figure 1, III) based on previous experiments involving aliphatic amines.⁹ Addition of DBU to this solution initiated the catalytic reaction and resulted in the formation of a new, sharp resonance at 70.1 ppm (Figure 1, C). This spectrum exhibits a chemical shift equivalent to that of a separately prepared DBU-bound OA complex (Figure 3, D), providing evidence that this base-bound Pd species is likely the resting state of the catalyst during the catalytic reaction.¹⁵ While the involvement of the base in the resting state is generally uncommon, Hartwig has shown that Pd-catalyzed cross-coupling of weakly binding fluoroalkylamines facilitated by phenoxide bases proceeds through a $\text{Pd}(\text{Ar})\text{OPh}$ intermediate, demonstrating the ability for the base to bind to Pd.¹⁶ Taking into account the resting state of our reaction, the kinetic orders that we have determined suggest that the RLS of the reaction involves the substitution of DBU by *p*-fluoroaniline on the DBU-bound OA complex. This process would be expected to produce a negative order in DBU and a positive order in 2. Other possible rate-limiting steps, such as deprotonation of the amine-bound OA complex (Figure 4, VII) and reductive elimination of the amido complex (Figure 4, VIII), would likely exhibit zeroth order dependence on base in disagreement with our experimental results.¹⁷

The identity of the resting state and the kinetic orders implicate an amine–DBU exchange process at palladium as the turnover-limiting step. The general mechanism is shown in Figure 4. The rate-limiting exchange can take place through either a dissociation–association (D/A) sequence, in which reversible coordination of an amine takes place before DBU dissociation (Figure 5, A), or through an association–dissociation (A/D) sequence (Figure 5, B), in which reversible detachment of DBU takes place before amine association. Due in part to the observed reluctance of the highly hindered AlPhos-

A. Dissociation/Association (D/A) Mechanism



B. Association/Dissociation (A/D) Mechanism

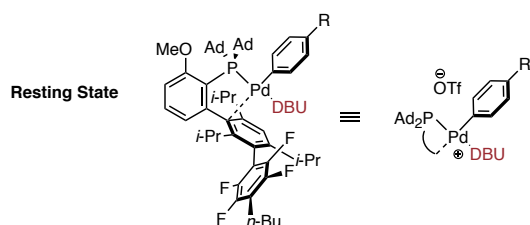
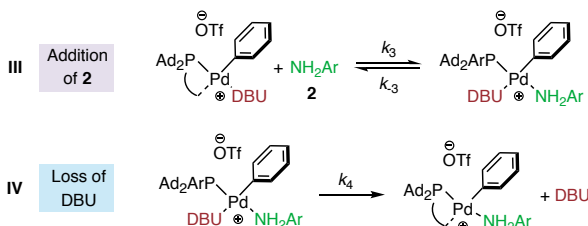


Figure 5. Elementary steps of the two possible rate-limiting steps. (A) Dissociation/Association mechanism; **I**, DBU-dissociation event; **II**, amine binding event. (B) Association/Dissociation mechanism; **III**, amine binding event; **IV**, DBU-dissociation event.

supported arylpalladium(II) to accept two amine ligands, (see Supporting Information for comparisons and further discussion), we assumed the D/A mechanism to be operative. To determine which step within this manifold limits the rate of the reaction, we conducted a Hammett analysis.¹⁸ The rates of catalytic reactions of **2** with aryl triflates bearing various *para*-substituents were measured using ¹⁹F NMR spectroscopy. These rates, normalized to that of phenyl triflate, were plotted against previously determined Hammett constants¹⁹ associated with the substituents. For electron-withdrawing substituents ($\sigma > 0$), such as *p*-chloro, phenylazo, and nitro, a negative Hammett correlation was observed ($\rho = -1.97$), indicating a deceleration effect for electron-poor triflates. In contrast, for electron-releasing substituents ($\sigma < 0$), such as *p*-methyl, methoxy, and *N,N*-dimethylamino, a positive Hammett correlation was observed ($\rho = +1.27$), indicating a deceleration effect for electron-rich triflates. This class of downward curving Hammett plots²⁰ is indicative of a shift in rate-limiting step within this two-step sequence.²¹

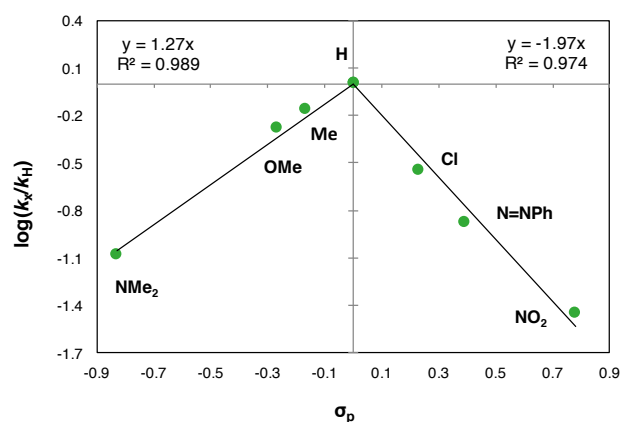
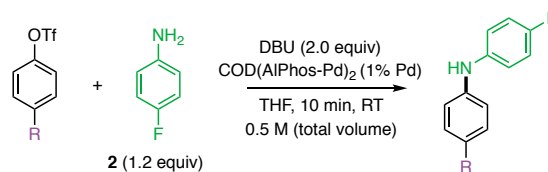


Figure 6. Hammett plot of the coupling of 4-substituted aryl triflates and **2** in the presence of DBU. Data point labels indicate *para* substituent. Reaction rates are plotted as the average of two runs. Positive and negative trend line intercepts are fit to the plot origin.

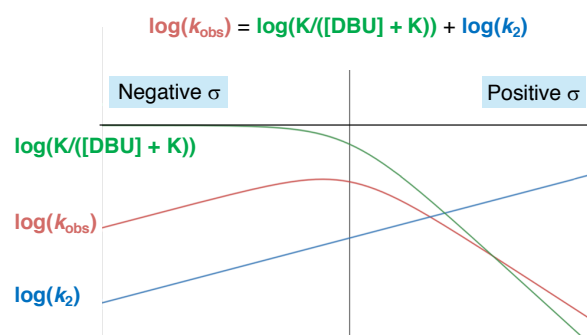


Figure 7. Mathematical model of $\log(K_{\text{obs}})$ as a function of Hammett parameter. The green curve and blue curve show the individual effects of σ on the separate terms of the equation, while the red curve shows the sum of these effects.

$$(1) \quad \text{rate} = \frac{k_2 K [\text{Pd}]_{\text{total}} [\text{2}]}{K + [\text{DBU}]} \quad \left| \quad K = \frac{k_1}{k_{-1}} \right.$$

$$(2) \quad k_{\text{obs}} \propto \frac{k_2 K}{K + [\text{DBU}]}$$

$$(3) \quad \log(k_{\text{obs}}) = \log(k_2) + \log\left(\frac{K}{K + [\text{DBU}]}\right)$$

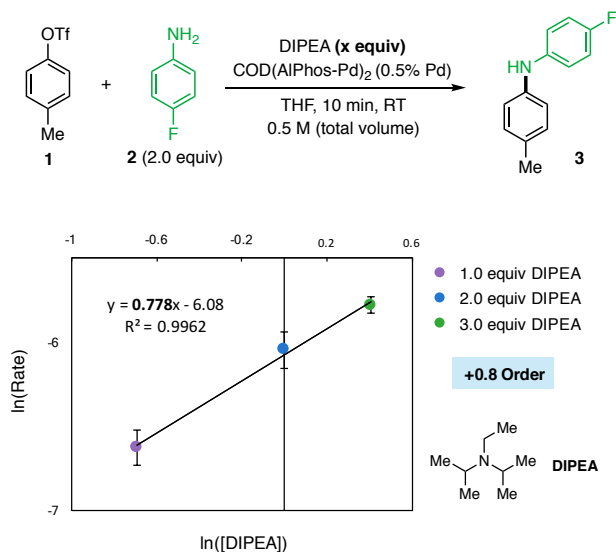


Figure 8. Logarithmic plot showing a positive, partial order dependence on DIPEA in the coupling of **1** and **2**. The average rate is plotted for reactions performed at 0.5, 1.0, and 1.5 M relative to DIPEA. Error bars represent the logarithmic standard deviation of the three reaction rates that were determined from six time points per reaction (18 time points total).

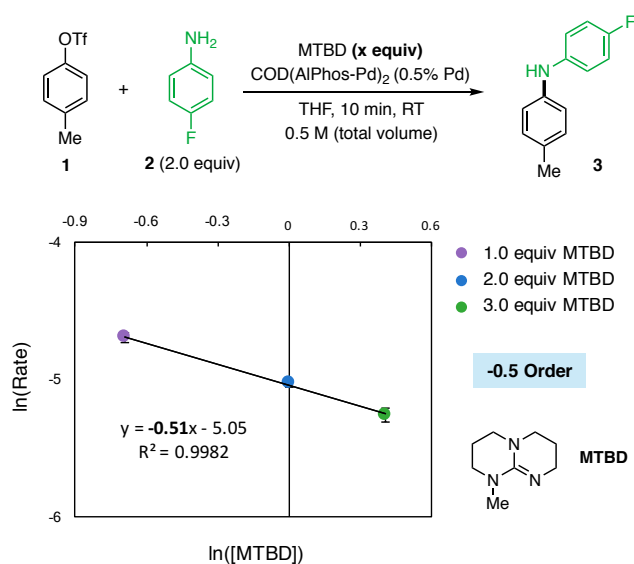


Figure 9. Logarithmic plot showing a -0.5 order dependence on MTBD in the coupling of **1** and **2**. The average rate is plotted for reactions performed at 0.5, 1.0, and 1.5 M relative to MTBD. Error bars represent the logarithmic standard deviation of the three reaction rates that were determined from six time points per reaction (18 time points total).

We can explain this behavior more precisely by examining the steady-state rate law of the D/A mechanism (equation 1, see Supporting Information for derivation).²² We can assume that step II (Figure 5, A) is effectively irreversible because of fast and irreversible subsequent steps: deprotonation by DBU and reductive elimination, respectively.²³ Kinetically, we observed a negative dependence on DBU, which supports our assumption regarding deprotonation. If this step were rate limiting, one would expect to observe zero dependence on this reagent. Thus, the observed rate (k_{obs}) is proportional to a function of two constants (equation 2). Accordingly, the logarithm of the rate constant is the sum of two terms (plus neglected constants, equation 3). The first term, $\log(k_2)$, can be predicted to have a small, positive dependence on the Hammett parameter of the aryl substituent, since the positive charge at palladium decreases in the transition state of step II. The second term can be analyzed by examining its behavior at the extremes. For $K \gg [\text{DBU}]$, this term approaches zero. Meanwhile, for $K \ll [\text{DBU}]$, the term approaches $\log(K/[\text{DBU}])$, which can be predicted to have a large, negative dependence on the Hammett parameter of the aryl substituent, since the positive charge at palladium increases during dissociation of DBU in step I. We note that this negative dependence, which results from an equilibrium (ground state) effect in step I, is likely larger in magnitude than the positive dependence of the first term, which results from a kinetic (transition state) effect in step II. A summary of predicted behavior of each of these terms is shown in Figure 7.

The composite $\log(k_{\text{obs}})$ is shown in red as the sum of the two terms of equation (3), which are separately shown in blue and green. The red curve exhibits the same qualitative behavior as is observed in experiment. Intuitively, the non-linear rate relationship indicates the combination of negative q for step I, positive q for step II, and a shift from step II rate-limiting to step I rate-limiting with increasing q . Finally, while this proposed mechanistic scheme provides good agreement with our kinetically-derived mathematical model and understanding of inductive effects, these experiments do not exclusively rule out other deviations from the D/A mechanism.²⁴ For instance, we note that this analysis holds even if the A/D mechanism is operative rather than the D/A, since the rate laws for these two mechanisms have the same form (see Supporting Information for further discussion).

Based on the intricate dependence of the reaction rate on electronic effects, it is clear that the mechanism of this transformation is sensitive to subtle changes in substrate structure. Accordingly, we considered whether structural variations to the base would have large effects on the kinetic behavior of the reaction. If so, the ability to perturb the system to circumvent inhibition by DBU might lead to a catalytic system manifesting greater reactivity. To study this, we varied both the steric properties and overall concentration of the base and measured the effects of these changes on the reaction rate. We expected that the use of a more hindered base should disfavor the formation of a base-bound Pd complex due to steric interactions. We examined *N,N*-diisopropylethylamine (DIPEA) and 7-methyl-1,5,7-triazabicyclo[4.4.0]dec-5-ene (MTBD), both soluble, organic bases with Lewis-basic nitrogen atoms that are in sterically hindered environments..

The kinetic orders of the reaction of **1** and **2** in the presence of either DIPEA or MTBD were determined by measuring the rate of formation of **3** over time, in an analogous manner to the above-describe kinetic experiments. As the concentration of DIPEA was varied between 0.5 and 1.5 M, a positive linear relationship was observed. Plotting $\ln(\text{Rate})$ vs $\ln([\text{DIPEA}])$ provided a straight line with a slope of +0.79 (Figure 8, A), which indicates a positive, partial first order dependence on [DIPEA]. To investigate the origin of the partial order, we determined the resting state of the catalyst during the reaction catalytic process using ^{31}P NMR. We observed multiple broad signals which we believe correspond to both the DIPEA-bound OA complex and 2-bound OA complex (See the Supporting Information for copies of the spectra). This result demonstrates that there is an equilibrium between the base-bound and amine-bound OA complexes and suggests that more hindered bases bind to the Pd center less efficiently than DBU. It is likely that the presence of multiple resting states of the catalytic reaction results in non-integral order reaction kinetics. For example, if the deprotonation of the amine-bound OA complex, which is predicted to exhibit +1 order in DIPEA, and the dissociation of the DIPEA from the OA complex, which is predicted to exhibit -1 order in DIPEA, are each partially rate-limiting, a non-integral order could result.

For coupling reactions carried out using MTBD as the base (cf. Figure 9), the rate of product formation decreased as the concentration of base increased. The logarithmic plot of Rate vs [DIPEA] showed a slope of -0.51, indicating a negative partial order in MTBD concentration (Figure 9). As in the case of DIPEA, we

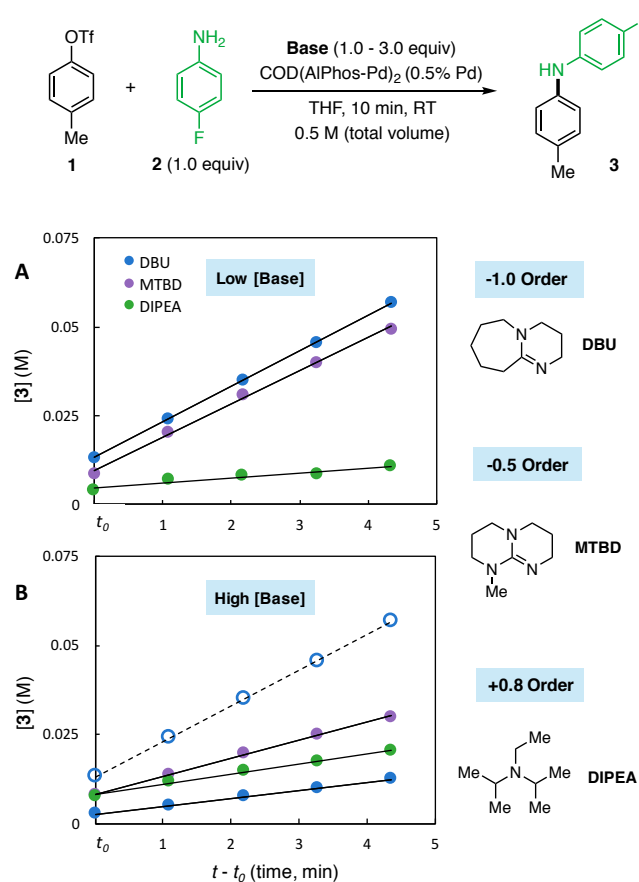
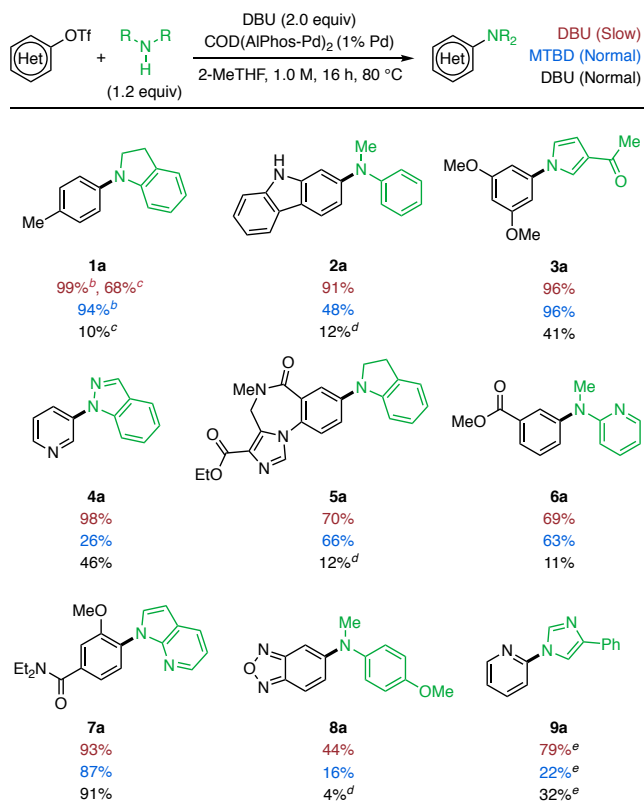


Figure 10. Comparison of the order in [base] and the initial rates of the reactions facilitated by MTBD, DBU, and DIPEA at low and high base concentrations. Data points are the average of three runs. (A) Reactions performed at a low base concentration (1.0 M). (B) reactions performed at a high base concentration (3.0 M). The dotted reference line shows the original rate under low [DBU] concentrations.

believe that the steric environment of the key basic nitrogen in MTBD decrease the efficiency of its binding to the Pd center. In this case the resting state in the catalytic process is likely a mixture of MTBD-bound and 2-bound oxidative addition complex.²⁵ Comparing these results with those employing DBU (*vide supra*) suggests that the resting state of the catalytic reaction depends on the relative binding ability of the base and the aniline substrate. In coupling reactions where the base outcompetes the aniline for binding to the Pd center of the OA complex, both a base-bound resting state and a negative order dependence on base are observed. In contrast, if the amine is more nucleophilic and/or less hindered than the base, then it may outcompete the base in binding to the Pd center. In this case, both an amine-bound OA complex resting state and positive order dependence on amine are observed.

Table 1: Coupling of heterocyclic amines and secondary aryl amines^a



^aRed yields indicate the average isolated yield of two reactions performed with a slow addition of DBU (syringe pump, 12 h, 0.167 mmol/h); blue isolated yields indicate the normal addition of MTBD and are denoted as a single isolated yield. black yields indicate a normal addition of DBU and are denoted as a single isolated or NMR yield. Standard reaction conditions: aryl triflate (0.50 mmol), nucleophile (0.60 mmol), base (1.0 mmol), COD(AIPhos-Pd)₂ (2.5 μmol, 1% Pd) 2-methyltetrahydrofuran (THF, 1.0 M), 80 °C for 16 h. ^bReaction performed at 60 °C. ^c¹H NMR yield of reaction performed at 60 °C with 1.2 equiv of DBU. ^d¹H NMR yield. ^e Reaction performed with 2% Pd.

While the steric properties of the bases had a predictable effect on the reaction order, the overall reaction rates depend on a combination of less predictable factors, including the kinetic and thermodynamic basicity of the chosen base. In the case of reactions performed at low base concentrations (1.0 M), the reactions facilitated by DBU and MTBD were found to be faster than the reaction with DIPEA (Figure 10, A) despite their inhibitory behavior. While DIPEA was found to have a positive dependence on the reaction (+0.8), this base is much weaker than DBU and MTBD, which may affect the overall rate of the reaction due to a slow

deprotonation event. In the case of high base concentrations (3.0 M), the rate of reactions facilitated by DBU and MTBD decreased significantly (Figure 10, B) as expected based on their order dependence. The rate of the reaction facilitated by DIPEA increased under these conditions such that it surpassed the rate of the reaction facilitated by DBU. These results demonstrate that both the base structure *and* concentration affect the reaction rate and should be taken into consideration when employing a soluble base in a Pd-catalyzed C–N cross-coupling. From these observations, we conceived two approaches to improve difficult coupling reactions involving weakly binding amines and potentially nucleophilic bases. First, in reactions that employ bases that slow the rate of the reaction, such as DBU, decreasing the concentration of base should favor formation of an aniline-bound OA complex, increasing the rate of the reaction. We chose to test this hypothesis through slow (syringe-pump) addition of DBU in the coupling reaction of aryl triflates with secondary anilines and aromatic amines.

Under our previously reported reaction conditions using soluble amine bases,⁹ the coupling reactions of secondary amines with aryl halides or triflates resulted in low or no formation of the desired product. For example, the reaction of **1** and indoline resulted in a 10% yield of the desired product (Table 1, **1a**) when heated to 60 °C in 2-methyl tetrahydrofuran (2-MeTHF) in the presence of an AIPhos-supported Pd catalyst and DBU. However, when 2.0 equivalents of DBU were added slowly (0.167 mmol/h) via syringe pump, the desired product was obtained in nearly quantitative isolated yield (Table 1, **1a**). Simply employing lower loadings of base resulted in incomplete conversion of the starting material.²⁶ A number of substituted aryl triflate electrophiles and weakly binding amines were examined, and the yields of the reaction employing the slow addition of base (red yields) were compared to those obtained using the previously reported reaction conditions (black yields). More complex electrophiles, including a Flumazenil derivative (**5a**), could be efficiently transformed to products under these reaction conditions when slightly elevated temperatures were used (80 °C). Notably, a benzoxadiazole electrophile (**8a**) was found to undergo the desired C–N coupling in modest yield despite the possibility for these heteroatom-heteroatom containing substrates to undergo undesired reactions with Pd.^{8c} Other weakly coordinating amines, including acyclic secondary amines such as *N*-methyl aniline derivatives are suitable substrates (**2a**, **6a**, and **8a**). Aromatic amines, including a pyrrole (**3a**), an indazole (**4a**), and an imidazole (**9a**) were found to be reactive coupling partners under the slow addition reaction conditions. In the case of the indazole

and imidazole couplings, *N*1 selectivity was observed and there was no detectable formation of the *N*2-regioisomer.²⁷ In a few cases, limiting the DBU concentration has no effect on the reaction yield, such as with substrate **7a** and indole (result not shown). In these cases, the desired product is still obtained in high yields. Currently, the reactions of sterically demanding secondary aryl amine nucleophiles, including *N*-benzyl aniline, exhibit low or no reactivity under these conditions.

In a second approach to overcome the inhibiting effect of DBU, MTBD, a considerably more expensive base than DBU, was found to be a suitable alternative in the coupling reactions of some weakly binding amines for cases in which syringe pump addition of DBU is considered undesirable. Because MTBD is less inhibiting than DBU (*vide supra*), it could be employed as base without requiring the use of syringe pump addition. The yields of coupling reactions facilitated by MTBD (Table 1, blue yields) were similar to those observed with the slow addition protocol using DBU in coupling reactions of many *N*-alkyl aryl amines (**1a**, **5a**, and **6a**). However, in some cases, the desired product was still observed in modest yields (e.g., **2a,b** and **4a**). The ability to use either DBU with syringe pump addition or MTBD to facilitate these coupling reactions provides two complementary approaches for the synthetic chemist to choose from depending on whether cost-efficiency or operational simplicity is more important to them.

CONCLUSION

In summary, we have demonstrated that organic amine bases play an important but complicated role in the coupling of aryl triflates with aryl amines. We have determined that the nucleophilicity of the base, relative to that of the amine substrate, dictates the resting state of the reaction, as well as the rate limiting step. For example, it was determined that for the Pd-catalyzed coupling of an aryl triflate and an aniline, the reaction displayed a -1 order in DBU when this amidine was used as the base. This and other mechanistic insights were used to develop modified reaction conditions that enabled the coupling of weakly binding heterocyclic amines and secondary aryl amines with aryl triflates. The slow addition of DBU was found to offer a solution for these couplings, while the normal, single-batch addition of MTBD provided an operationally simpler but costlier alternative procedure.

ASSOCIATED CONTENT

Supporting Information

Experimental procedures, raw kinetic data, and spectral data.

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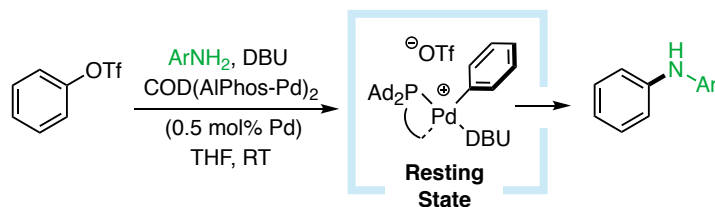
Notes

MIT has filed patents on ligands and precatalysts that are described in this manuscript, from which S.L.B and former coworkers receive royalty payments.

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TOC Image



- **Base nucleophilicity affects kinetic dependence**
- **Slow addition of base enables 2° aryl amine couplings**

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(23) If the base is not strong enough to rapidly deprotonate the amine-bound OA complex, then the amine may dissociate to reform the OA complex. In this case, there is an equilibrium between the two complexes (Figure 5A, **II**). We believe that our assumption holds when DBU or bases of a greater strength are employed.

The thought is that once we reach the amine-bound complex, we are essentially committed to product, rather than going back to the DBU-bound complex. As you mention, this may be true for DBU but not in general. If the base is so weak that deprotonation is slow, it will often go back to DBU-bound complex.

(24) Other theoretical rate limiting steps include deprotonation of the amine-bound OA complex and reductive elimination of the amido complex. Because the kinetics showed a clear inverse dependence on DBU for the coupling *p*-tolyl triflate, a slightly electron-rich substrate, it is unlikely that deprotonation is rate limiting. Moreover, for coupling reactions facilitated by dialkylbiaryl monophosphine ligands, the reductive elimination step is rarely rate-limiting. For particularly large ligands such as AlPhos, this is also thought to be true. In two special cases, the reductive elimination of diaryl amines and fluoroalkyl amines have been rate-limiting. See references 3b and 16 respectively.

(25) Attempts to study the resting state by ^{31}P NMR did not yield a definitive result. Both **2** and MTBD independently bind the OA complex, but the combination of these reagents did not yield any readily identifiable species, which we attribute to rapid exchange of the aniline and MTBD at the Pd(II) center.

(26) Optimizing the addition rate of DBU may enable the cross-coupling with less than 2.0 equivalents of base. However, this may be substrate dependent given the easily influenced reaction rate. For the purpose of this study, 2.0 equivalents added at 0.167 mmol/h provided the product in high yields for substrates with diverse substitution patterns and electronic characteristics.

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